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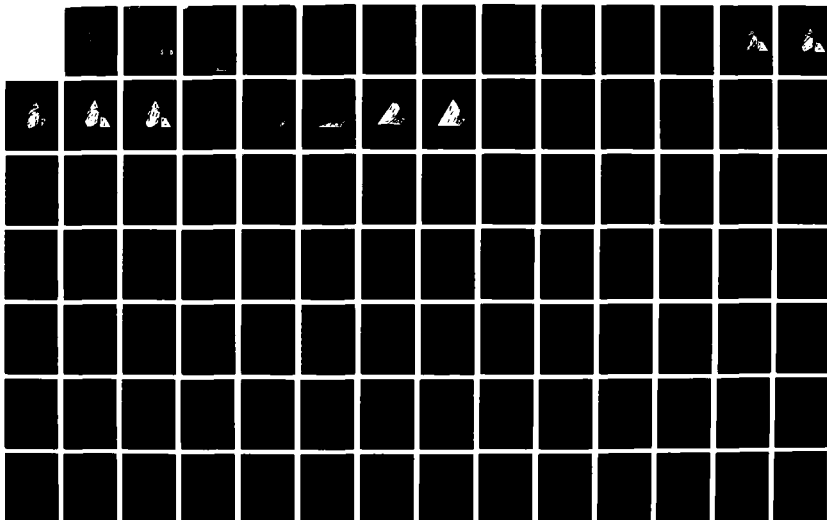
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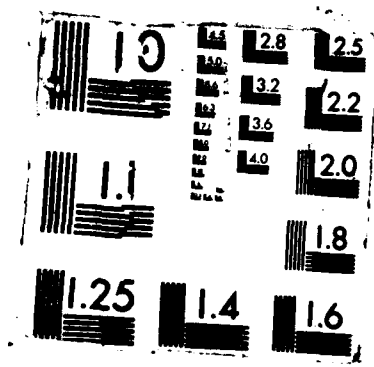
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FINAL REPORT
ON

Contract F-49620-84-C-0078

APPLICATION OF COMPUTER METHODS FOR CALCULATION
OF MULTI COMPONENT PHASE DIAGRAMS OF HIGH TEMPERATURE
STRUCTURAL CERAMICS

1 AUGUST 1984 TO 31 JULY 1987

Submitted to

Air Force Office of Scientific Research (AFSC)
Bolling Air Force Base, D.C. 20332
Att. Major Joseph Hager (202-767-4933)


by
Larry Kaufman

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19. KEY WORDS (Continue on reverse side if necessary and identify by block number) <table border="0"> <tr> <td>Data Base</td> <td>Quasi-binary</td> <td>Thermocalc,</td> </tr> <tr> <td>Phase Diagrams,</td> <td>Quasi-ternary</td> <td>Titanium Carbonitride,</td> </tr> <tr> <td>Structural Ceramics,</td> <td>Germania Ceramics,</td> <td>Yttria Ceramics,</td> </tr> <tr> <td>Thermochemistry,</td> <td>Zirconia Ceramics,</td> <td>Hafnia Ceramics, ←</td> </tr> </table>			Data Base	Quasi-binary	Thermocalc,	Phase Diagrams,	Quasi-ternary	Titanium Carbonitride,	Structural Ceramics,	Germania Ceramics,	Yttria Ceramics,	Thermochemistry,	Zirconia Ceramics,	Hafnia Ceramics, ←
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ABSTRACT (CONCLUDED)

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ABSTRACT

A data base is being developed for calculation of quasi-binary and quasi-ternary phase diagrams of ceramic systems. Previous segments of this base cover combinations of Cr_2O_3 , MgO , Al_2O_3 , SiO_2 , CaO , Si_3N_4 , AlN , BeO , Y_2O_3 and Ce_2O_3 . Sixty-six quasi binary and nineteen quasi ternary systems have been calculated. The current work extends the base to cover GeO_2 , HfO_2 , ZrO_2 and TiO_2 which are of interest in applications requiring toughness and structural performance at high temperatures. This has been effected by employing available sources of thermochemical and phase diagram data. Recently it has been shown that by alloying GeO_2 with SiO_2 a whole range of glasses can be synthesized with tailor-made coefficients of expansion. Utilization of such compositions offers the possibility of enhancing the high temperature oxidation resistance of ceramic composites in which a mixed GeO_2 - SiO_2 phase with a desired CTE would replace the conventional SiO_2 as a filler. One of the major obstacles in the development of complex composite systems is the lack of phase diagram information which can be used to guide the fabrication and processing of a new material and help to predict its performance. The current methods of employing models to predict high temperature behaviour has proven useful when basic data is unavailable or too costly and time consuming to obtain by conventional means. This method consists of developing a data base of thermochemical and phase diagram information in analytical form and employing computer models to extend the description to binary and ternary systems. This technique was also applied in order to provide information relative to the development of oxidation resistant coatings for superalloys through calculations of the Cr-Si-Ni and Al-Ni-Si systems between 700K and 1500K to establish minimum liquidus temperatures relevant to the silicide coating technology as well as descriptions of the CrO_2 - SiO_2 and NiO-SiO_2 system was provided. Calculations of the Fe-Ni-O , Fe-Cr-O , Ti-C-N and Al_2O_3 - Y_2O_3 - ZrO_2 systems were performed on the new "Thermocalc" system developed at the Royal Institute of Technology in Stockholm. These calculations which were carried out at the Massachusetts Institute of Technology in Cambridge, Massachusetts and at the National Bureau of Standards in Gaithersburg, Maryland at locations which were remote from the "Thermocalc" programs which were stored on VAX computers.

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I INTRODUCTION AND SUMMARY

A data base is being developed for calculation of quasi-binary and quasi-ternary phase diagrams of ceramic systems. Previous segments of this base cover combinations of Cr_2O_3 , MgO , Al_2O_3 , SiO_2 , CaO , Si_3N_4 , AlN , BeO , Y_2O_3 and Ce_2O_3 . Lattice Stability, Solution and Compound Phase Parameters have been derived covering the liquid, spinel, corundum, perillase, cristobalite, tridymite, quartz, hexagonal and beta prime phases which appear in the binary systems composed of pairs of these compounds. Compound phases formed from specific binary combinations of these compounds (i.e. $\text{MgO-Cr}_2\text{O}_3$) have been characterized. This description is based on observed thermochemistry and phase diagrams for the binary systems of interest. Selected ternary systems have been computed based on the foregoing data base for comparison with experimental sections in order to illustrate the usefulness of the data base. To date, sixty-six quasi binary and nineteen quasi ternary systems have been calculated. The current work extends the base to cover GeO_2 , HfO_2 , ZrO_2 and TiO_2 . The components are of particular interest in applying ceramic systems in applications requiring toughness and structural performance at high temperatures. This has been effected by employing available sources of thermochemical and phase diagram data. High temperature ceramics have received increased attention during the last few years for structural, thermal protection and engine applications. SIALONS and combinations of zirconia and hafnia with Al_2O_3 , SiC and Si_3N_4 have been shown to develop strength and toughness. This has opened the door to a whole range of new uses for these materials. Recently Slichting and co-workers have shown that by alloying GeO_2 with SiO_2 a whole range of glasses can be synthesized with tailor-made coefficients of expansion. Utilization of such compositions offers the possibility of enhancing the high temperature oxidation resistance of ceramic composites in which a mixed GeO_2 - SiO_2 phase with a desired CTE would replace the conventional SiO_2 as a filler. This kind of compositing would open an entire spectrum of new opportunities for synthesis of high temperature ceramics. One of the major obstacles in the development of complex composite systems is the lack of phase diagram information which can be

used to guide the fabrication and processing of a new material and help to predict its performance. The current methods of employing models to predict high temperature behaviour has proven useful when basic data is unavailable or too costly and time consuming to obtain by conventional means. This method consists of developing a data base of thermochemical and phase diagram information in analytical form and employing computer models to extend the description to binary and ternary systems. Recently J. Lorenz et al. applied this method successfully to SiC-ZrO_2 and $\text{SiC-ZrO}_2\text{-Al}_2\text{O}_3\text{-SiO}_2$ in order to evaluate composition effects and identify fabrication conditions. In the present work, the data base has been expanded by analyzing the following quasi-binary systems: $\text{GeO}_2\text{-HfO}_2$, $\text{GeO}_2\text{-TiO}_2$, $\text{GeO}_2\text{-Al}_2\text{O}_3$, $\text{GeO}_2\text{-MgO}$, $\text{GeO}_2\text{-CaO}$, $\text{GeO}_2\text{-SiO}_2$, $\text{TiO}_2\text{-MgO}$, $\text{HfO}_2\text{-SiO}_2$, $\text{HfO}_2\text{-MgO}$, $\text{HfO}_2\text{-CaO}$, $\text{Al}_2\text{O}_3\text{-HfO}_2$, $\text{HfO}_2\text{-Y}_2\text{O}_3$, $\text{HfO}_2\text{-TiO}_2$, $\text{Ce}_2\text{O}_3\text{-Al}_2\text{O}_3$, $\text{ZrO}_2\text{-HfO}_2$, $\text{ZrO}_2\text{-SiO}_2$, $\text{ZrO}_2\text{-CaO}$, $\text{Y}_2\text{O}_3\text{-CaO}$, $\text{Y}_2\text{O}_3\text{-MgO}$, $\text{TiO}_2\text{-Al}_2\text{O}_3$, $\text{TiO}_2\text{-SiO}_2$, $\text{TiO}_2\text{-CaO}$ and $\text{TiO}_2\text{-Y}_2\text{O}_3$. These results when combined with earlier findings were employed to compute a range of isothermal sections in the following quasi ternary systems sufficient to define their characteristics: $\text{MgO-TiO}_2\text{-SiO}_2$, $\text{MgO-SiO}_2\text{-GeO}_2$, $\text{GeO}_2\text{-MgO-CaO}$, $\text{HfO}_2\text{-CaO-MgO}$, $\text{HfO}_2\text{-SiO}_2\text{-ZrO}_2$, $\text{HfO}_2\text{-CaO-Y}_2\text{O}_3$, $\text{HfO}_2\text{-MgO-Y}_2\text{O}_3$, $\text{HfO}_2\text{-CaO-ZrO}_2$, $\text{SiO}_2\text{-HfO}_2\text{-Y}_2\text{O}_3$, $\text{MgO-SiO}_2\text{-HfO}_2$, $\text{TiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$, $\text{Al}_2\text{O}_3\text{-TiO}_2\text{-SiO}_2$, $\text{TiO}_2\text{-Al}_2\text{O}_3\text{-HfO}_2$ and $\text{MgO-SiO}_2\text{-TiO}_2$.

The coupled thermochemical/phase diagram technique was also applied in order to provide information relative to the development of oxidation resistant coatings for superalloys. Thus, calculations of the Cr-Si-Ni and Al-Ni-Si systems between 700K and 1500K in order to establish minimum liquidus temperatures relevant to the silicide coating technology. In addition a description of the $\text{CrO}_2\text{-SiO}_2$ and NiO-SiO_2 system was provided.

Finally, calculations of the Fe-Ni-O, Fe-Cr-O, Ti-C-N and $\text{Al}_2\text{O}_3\text{-Y}_2\text{O}_3\text{-ZrO}_2$ systems were performed on the new "Thermocalc" system developed at the Royal Institute of Technology in Stockholm. These calculations which were carried out at the Massachusetts Institute of Technology in Cambridge, Massachusetts and at the National Bureau of Standards in Gaithersburg, Maryland at locations which were remote from the "Thermocalc" programs which were stored on VAX computers. This work serves to illustrate how modern computing systems for calculating multicomponent phase diagrams can be used effectively to deal with a wide variety of practical problems.

The following personnel have been active in this program: L. Kaufman, D. Birnie, V. Farber, J. Pershan, E.P. Warekois, P. Neshe, J. Smith, D. Hay, M. Grujicic and W.S. Owen. The technical lectures and papers listed below were presented in connection with work performed under this contract.

1. "Calculation of Quasibinary and Quasiternary Ceramic Systems" CALPHAD XIV, M.I.T. Cambridge, MA June 1985.
2. "Calculation of Ternary Isothermal Sections in Ni-Cr-Al and Ni-Cr-Si Systems" CALPHAD XIV, M.I.T. Cambridge, MA June 1985.
3. "Binary Common Ion Alkali Halide Mixtures-Solid/Liquid Equilibria in Systems Showing Isodimorphism" CALPHAD (1986) vol. 10, No. 2 pp. 163-174.
4. "CALPHAD Generated Multicomponent Phase Diagrams for Elements II Through VI" Annual Meeting TMS-ASM, New Orleans, LA March 1986.
5. "Calculation of Multicomponent Ceramic Phase Diagrams" L. Kaufman CALPHAD XV London, England July 1986.
6. "Calculation of Quasibinary and Quasiternary Ceramic Systems" L. Kaufman, ASM symposium on USER APPLICATIONS OF PHASE DIAGRAMS, Orlando, Florida September 1986, Published July 1987.
7. "Calculation of Multicomponent Ceramic Phase Diagrams" L. Kaufman CALPHAD XVI Irsee, West Germany May 1987.
8. "Calculation of Multicomponent Ceramic Phase Diagrams" L. Kaufman, International Conference on Electronic Structure and Phase Stability of Advanced Ceramic Systems" Argonne National Laboratory Argonne Illinois, August (1987). Proceedings to be published in Physica B (1988)

II CALCULATION OF THE Cr-Si-Ni, Al-Ni-Si, CrO₂-SiO₂, and NiO-SiO₂ PHASE DIAGRAMS

Silicide coatings being considered for application to superalloy systems for special applications. Whenever such coatings are applied and the system heated the interactions between coating and base alloy can produce instabilities which reflect the phase diagram formed along the join between coating and base in specific directions. In order to gain some insight into such reactions the Cr-Si-Ni, Al-Ni-Si, CrO₂-SiO₂ and NiO-SiO₂ phase diagrams were computed along lines described earlier (1,2). the results are shown in Tables 1 and 2 and Figures 1-10 for the metallic systems and Table 3 and Figures 11 and 12 for the oxide systems. The metallic systems are based on the previous descriptions of the metal-silicon systems (3) and the Cr-Ni (4) and Ni-Al (5) systems. Tables 1 and 2 show that all of compound phases were considered to be ideal solutions and that only the ternary sigma phase was added to the Cr-Si-Ni system. Reference to Figures 1-5 shows that an extensive liquid zone exists at 1500K but this zone is virtually eliminated at 1300K. The equilibria calculated below 1300K is in agreement with experimental finding (6-8). The Al-Ni-Cr results in Figures 6-10 show extensive liquid fields at 1100K. The calculated equilibria at low temperature is in agreement with experimental results (9,10).

The calculated CrO₂-SiO₂ and NiO-SiO₂ phase diagrams shown in Figures 11 and 12 show liquids above 1880 and 1900K respectively.

TABLE 1

SUMMARY OF COMPOUND PARAMETERS FOR UNSTABLE COUNTER PHASES AND COMPOUND INTERACTION PARAMETERS IN THE Cr-Si-Ni System.

<u>Compound Chemistry</u>	<u>Name</u>	<u>Base Phase</u>	<u>C_{jk}</u> Joules/ g.at	<u>CAB</u> Joules/ g.at
(Cr,N1).75Si.25	W	B	0	0
(Cr,N1).625Si.375	U	B	0	0
(Cr,N1).5Si.5	Q	B	(Cr.5Si.5 and N1.5Si.5 are stable)	0
(Cr,N1).667Si.333	Z	B	0	0
(N1,Cr).333Si.667	R	A	0	0
(N1,Cr).6Si.4	P	A	0	0
(N1,Cr).667Si.333	H	A	0	0
(N1,Cr).714Si.286	T	A	0	0
(N1,Cr).75Si.25	X	A	0	0
Cr.61Si.11N1.28	Sigma	B	C=13807 + 8.368T	
(Ternary Phase)(Melting Point 1522°K, Entropy of Fusion=13.26 J/g.at°K)				

TABLE 2

SUMMARY OF COMPOUND PARAMETERS FOR UNSTABLE COUNTER PHASES AND COMPOUND INTERACTION PARAMETERS IN THE Al-Ni-Si System

<u>Compound Chemistry</u>	<u>Name</u>	<u>Base Phase</u>	<u>C_{jk}</u> Joules/ g.at	<u>CAB</u> Joules/ g.at
(Al,Si).75Ni.25	V	A	0	0
(Al,Si).6Ni.4	S	A	0	0
Ni.75(Al,Si).25	X	A	Ni.75Al.25 and Ni.75Si.25 are stable	0
Ni.333(Si,Al).667	R	A	0	0
Ni.5(Si,Al).5	Q	B	0	0
Ni.6(Si,Al).4	P	A	0	0
Ni.667(Si,Al).333	H	A	0	0
Ni.714(Si,Al).286	T	A	0	0

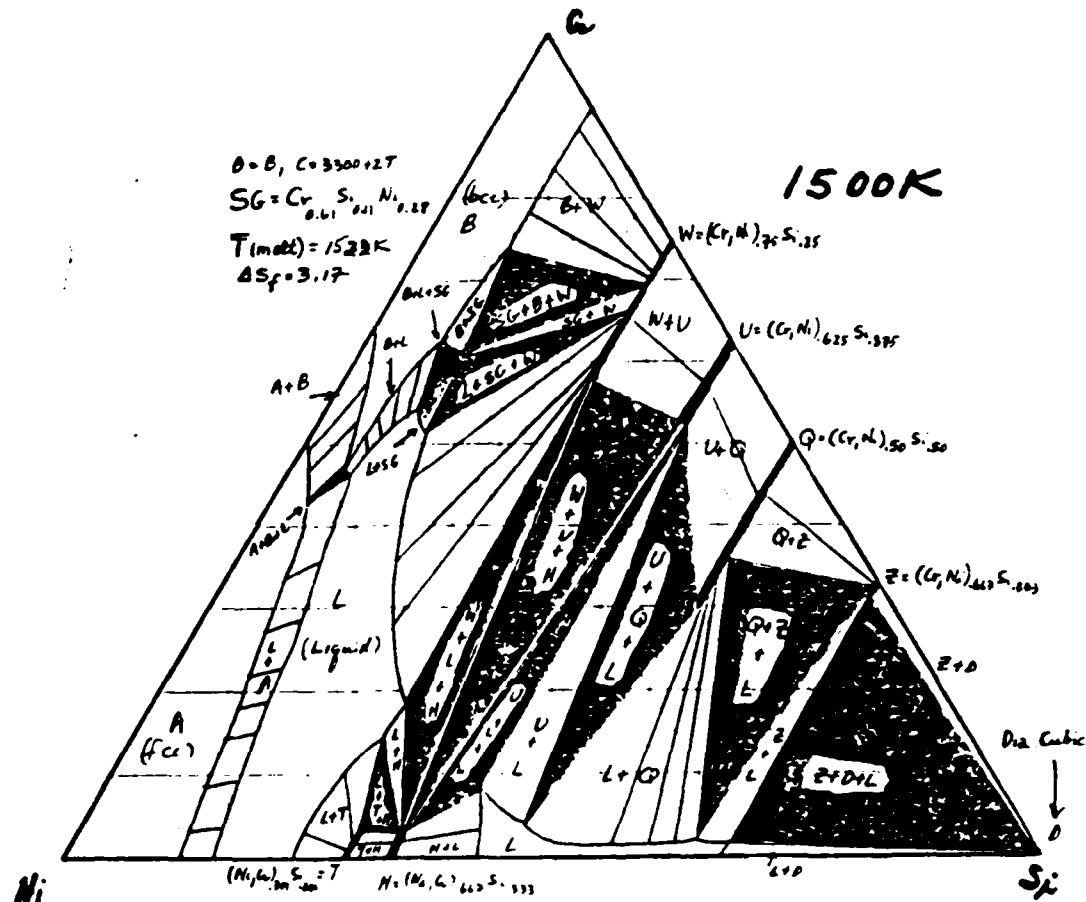


Figure 1 Calculated Isothermal Section in the Cr-Si-Ni System at 1500K

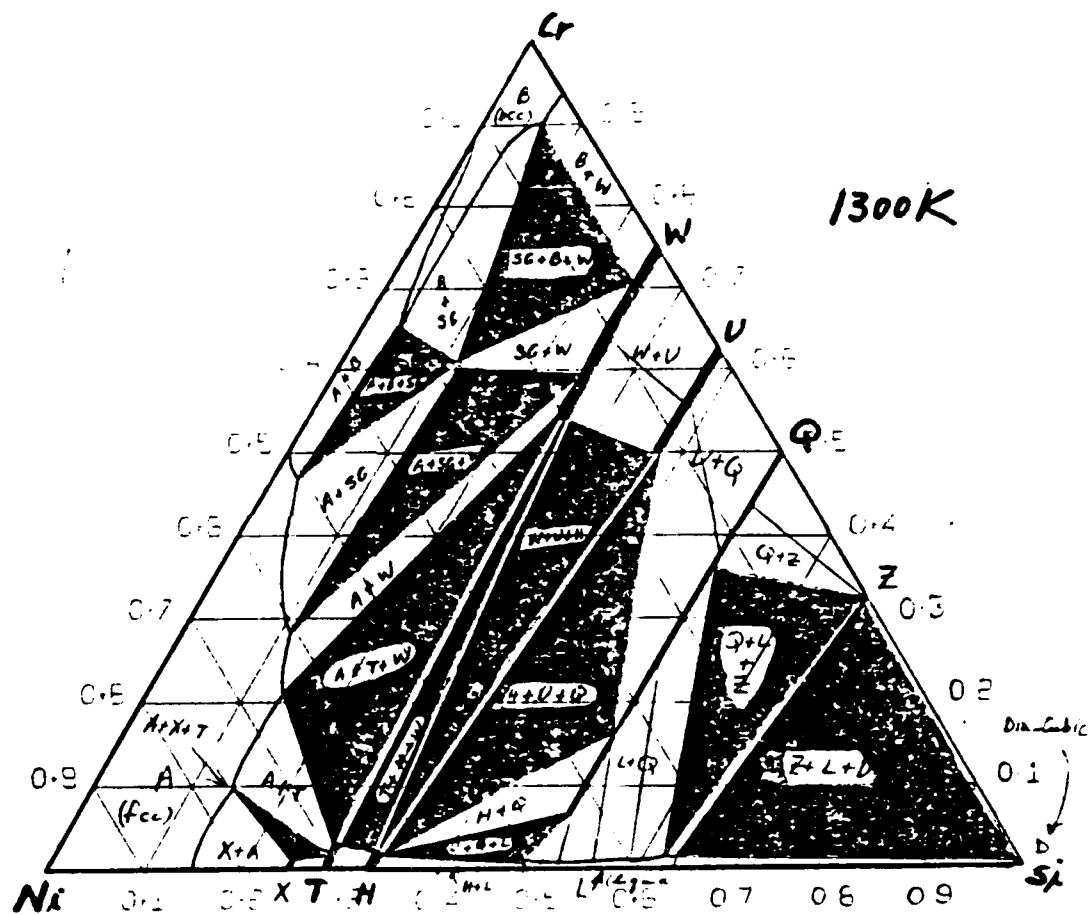
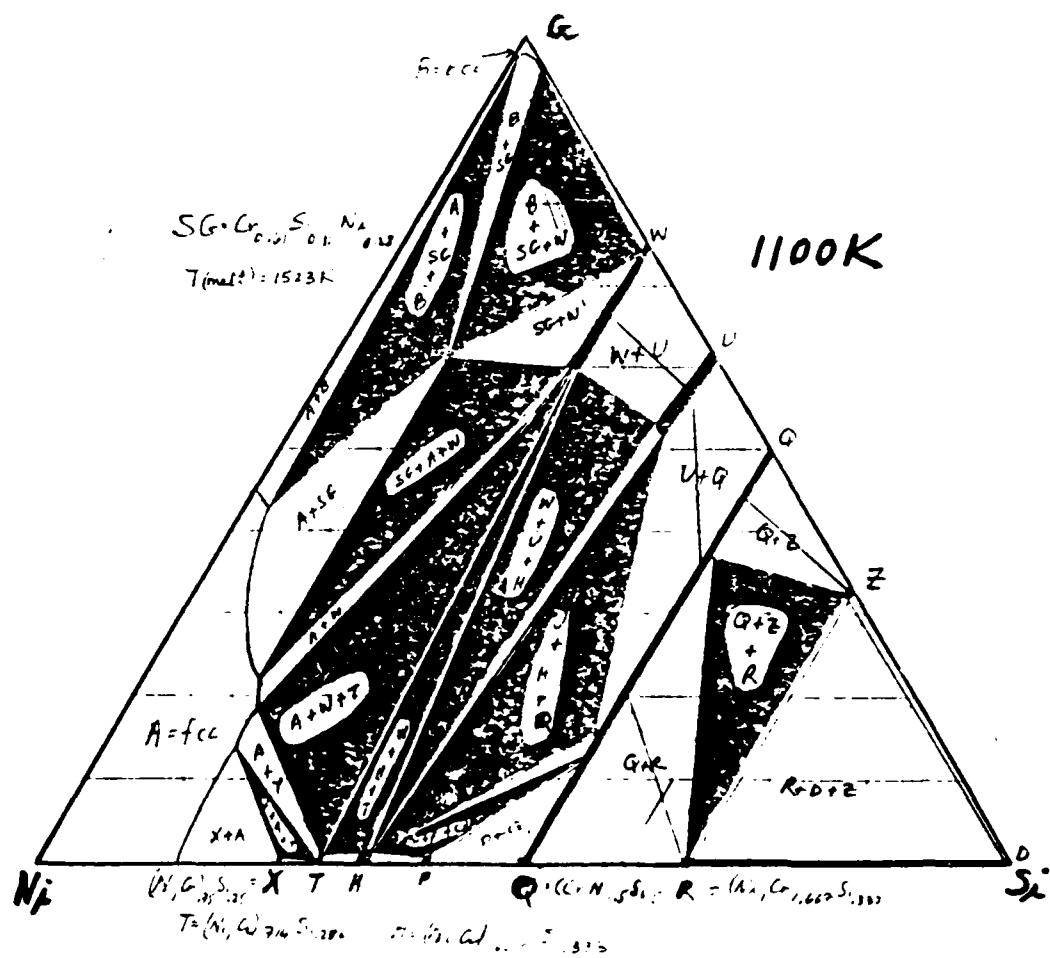


Figure 2 Calculated Isothermal Section in the Cr-Si-Ni System at 1300K.



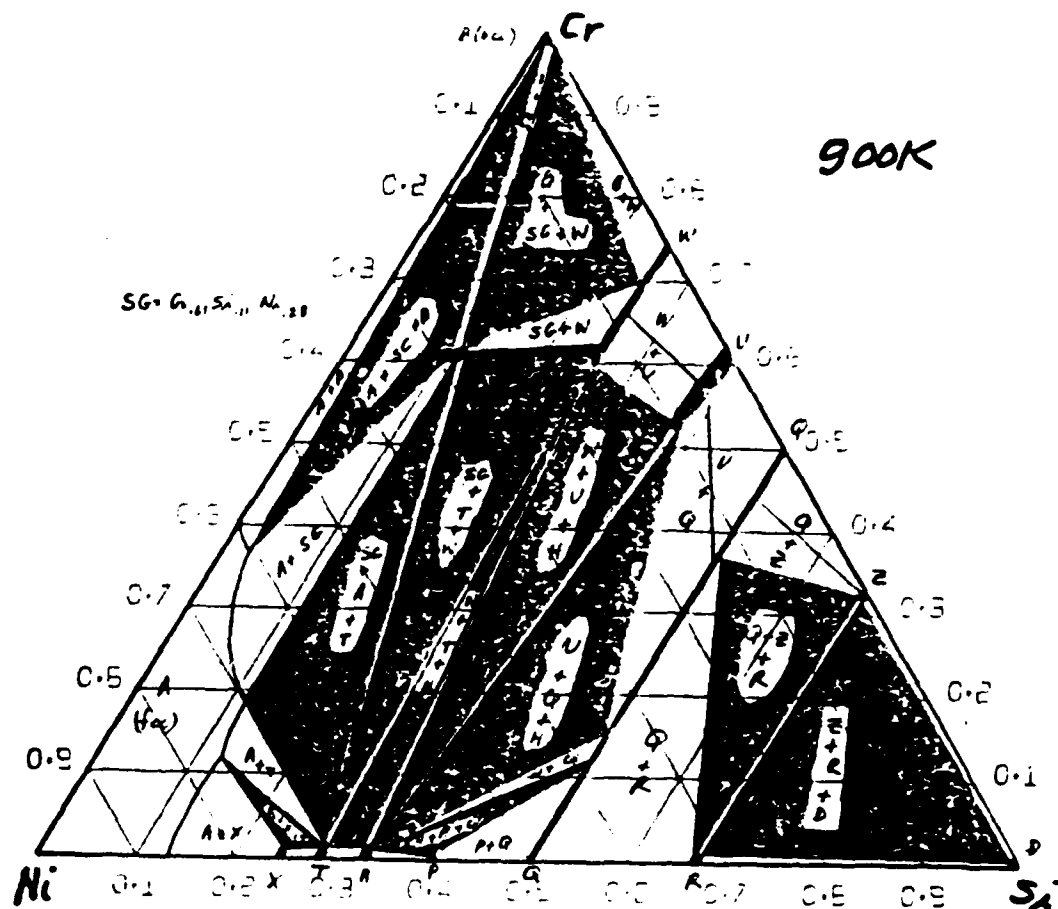


Figure 4 Calculated Isothermal Section in the Cr-Si-Ni System at 900K

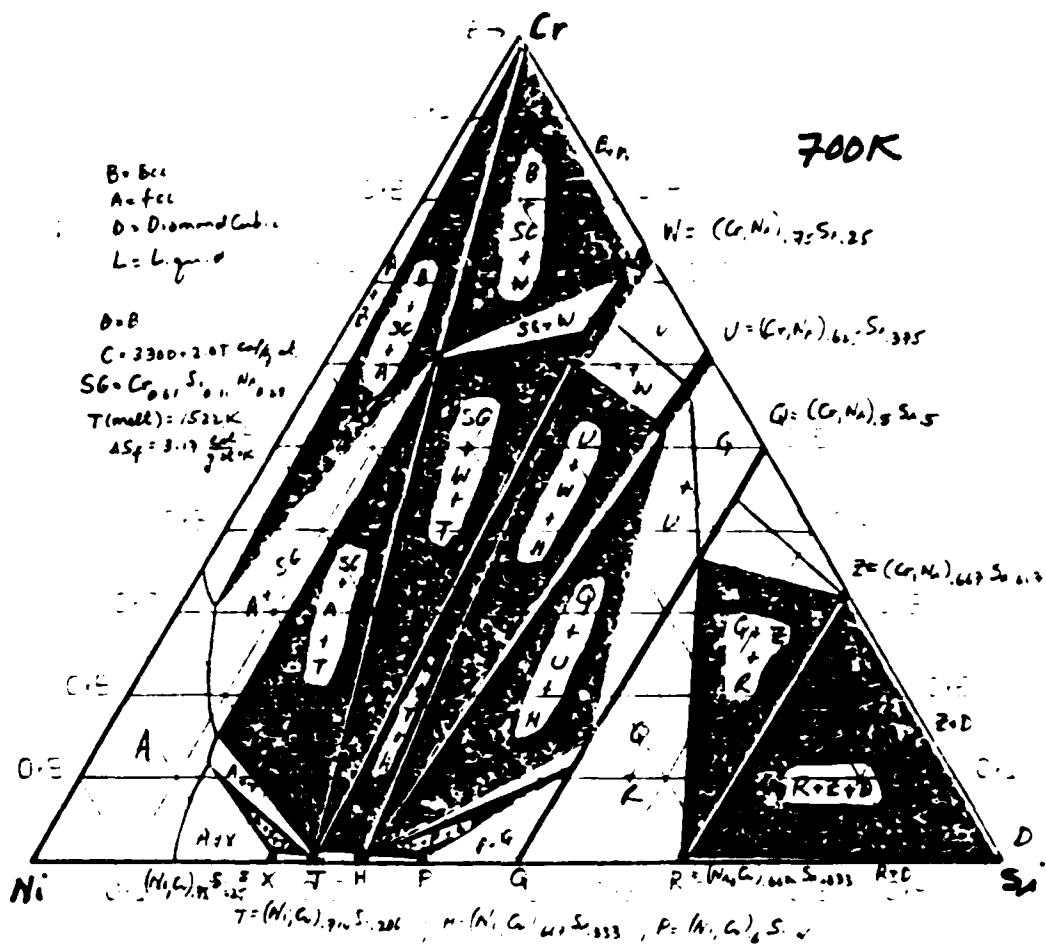


Figure 5 Calculated Isothermal Section in the Cr-Si-Ni System at 700K

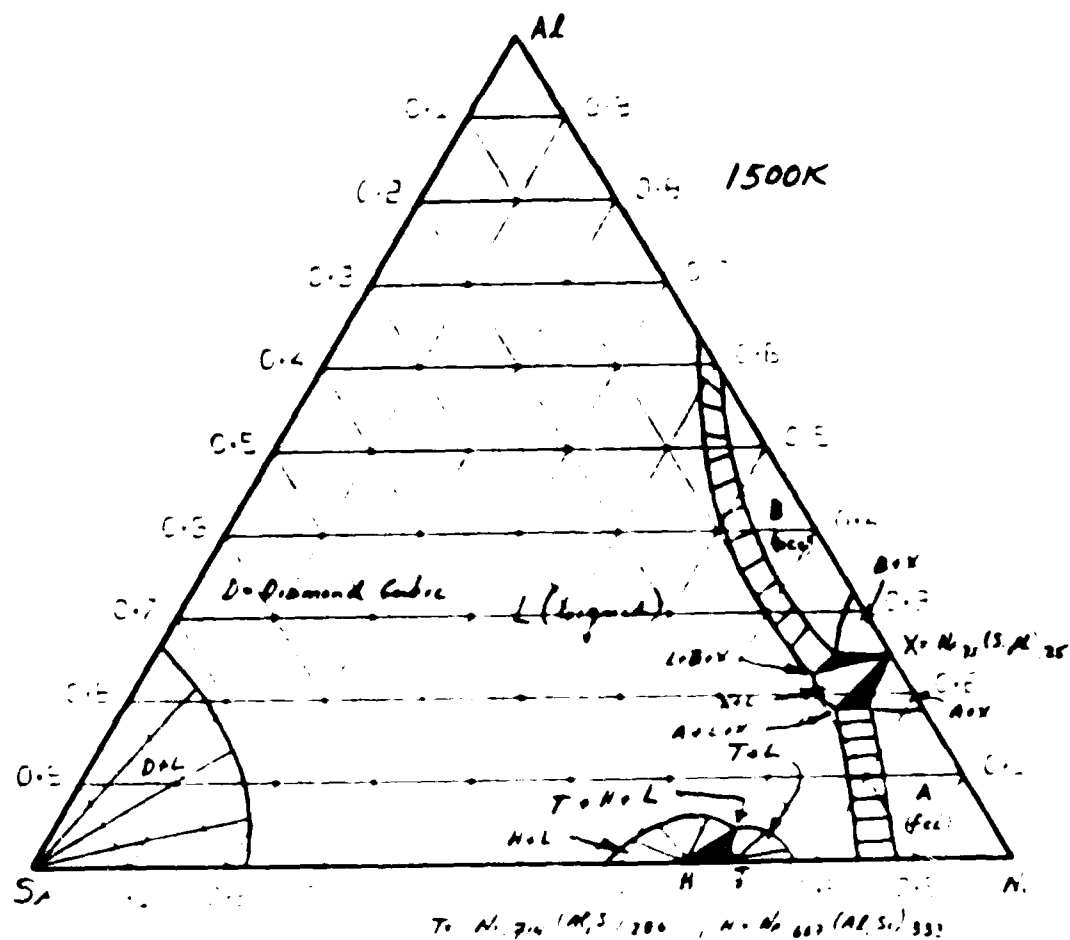
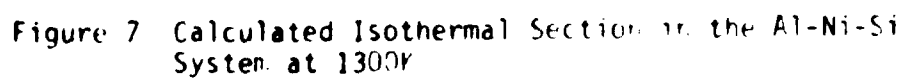


Figure 6 Calculated Isothermal Section in the Al-Ni-Si System at 1500K



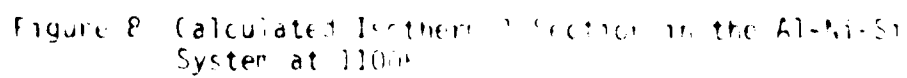


TABLE 3

SUMMARY OF LATTICE STABILITY, SOLUTION AND COMPOUND PARAMETERS FOR THE $\text{CrO}_2\text{-SiO}_2$ AND NiO-SiO_2 SYSTEMS (Units in J/g.at and J/g.at^{°K})

CO = $1/3\text{CrO}_2$, NO= $1/2\text{NiO}$, CO= $1/3\text{CrO}_2$, L=Liquid, C=Corundum,
X=Crystoballite, T=Trydimite, P=Periclase

COCOLC=34351-13.47T	NONOLP=25158-11.30T
COCOLX= -1.67T	NONOLX= -1.67T
SOSOLC= -2.09T	NONOLT= -2.01T
SOSOLP= -2.09T	

SOLUTION PARAMETER

LCOSO=LSOCO=68200		$0 < \text{XNO} < 0.43$
CCOSO=CSOCO=151879	LSONO=33472	LNOSO=379070-225.94T
XCOSO=XSOCO=151879	XSONO=115688-41.8T	XNOSO=400000-225.94T
TCOSO=TSOCO=151879	TSONO=115688-41.8T	TNOSO=400000-225.94T

$0.43 < \text{XNO} < 1.0$

LSONO=147310-97.15T	LNOSO=228166-97.15T
PSONO=189150-97.15T	PNOSO=270006-97.15T

COMPOUND PARAMETER

FA=Payalite= $(\text{SO}_{.429}\text{NO}_{.571})=1/7(\text{SiO}_2 \cdot 2\text{NiO})$
Base=P, Compound Parameter=198711-102.34T

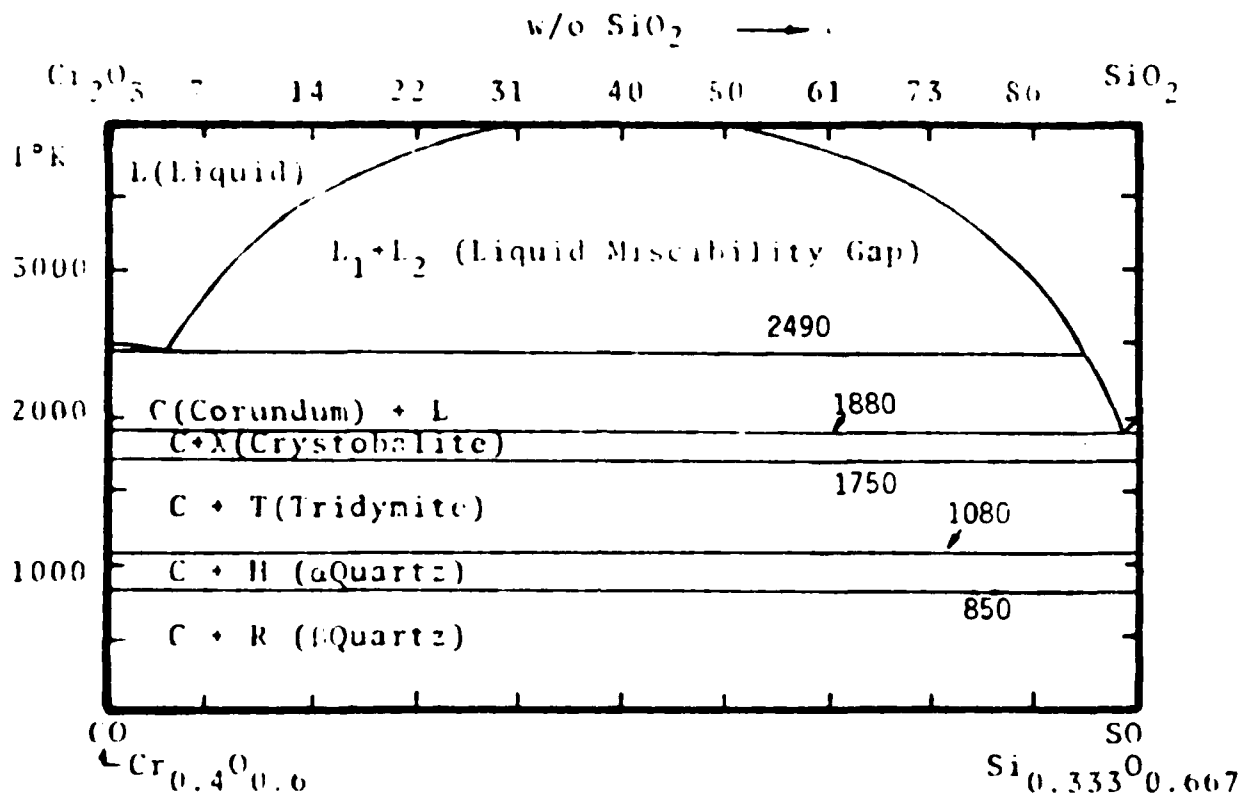


Figure 11. Calculated $\text{Cr}_{0.4}\text{O}_{0.6}$ - $\text{Si}_{0.333}\text{O}_{0.667}$ Phase Diagram

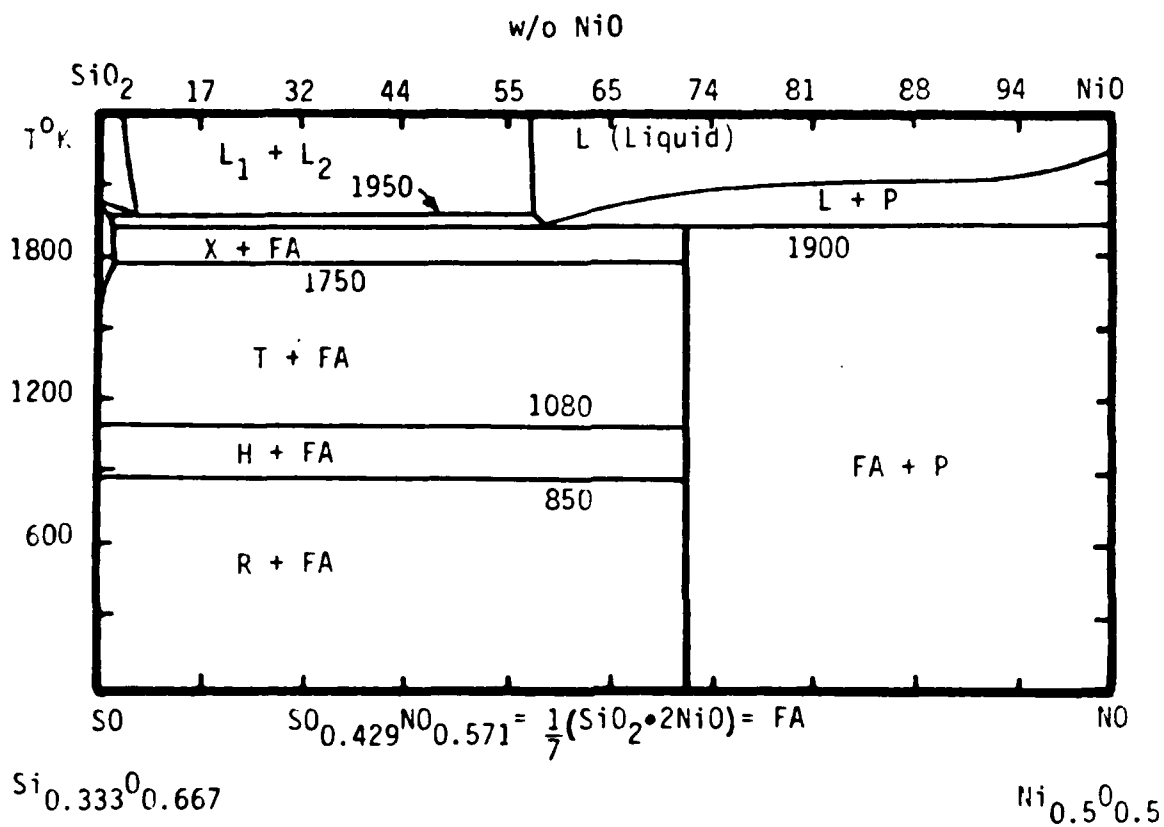


Figure 12 Calculated $\text{Si}_{0.333}\text{O}_{0.667} - \text{Ni}_{0.5}\text{O}_{0.5}$ Phase Diagram

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III. CALCULATION OF QUASIBINARY AND QUASITERNARY CERAMIC SYSTEMS

by

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(Presented October 1986, ASM Symposium on USER APPLICATIONS OF PHASE
DIAGRAMS, Orlando Florida Proceedings to be published in 1987)

ABSTRACT

A data base is being developed for calculation of quasi-binary and quasi-ternary phase diagrams of ceramic systems. Previous segments of this base cover combinations of Cr_2O_3 , MgO , Al_2O_3 , SiO_2 , CaO , Si_3N_4 , AlN , BeO , Y_2O_3 and Ce_2O_3 . Lattice Stability, Solution and Compound Phase Parameters have been derived covering the liquid, spinel, corundum, periclase, cristobalite, tridymite, quartz, hexagonal and beta prime phases which appear in the binary systems composed of pairs of these compounds. Compound phases formed from specific binary combinations of these compounds (i.e. $\text{MgO}\cdot\text{Cr}_2\text{O}_3$) have been characterized. This description is based on observed thermochemistry and phase diagrams for the binary systems of interest. Selected ternary systems have been computed based on the foregoing data base for comparison with experimental sections in order to illustrate the usefulness of the data base. The present paper extends the data base to cover GeO_2 , HfO_2 , ZrO_2 and TiO_2 . The components are of

particular interest in recent developments of structural high temperature ceramics and applications requiring unusual toughness. (6). In order to expand the data base nineteen quasi-binary systems have been analyzed and ten quasi-ternary systems have been calculated over a wide range of temperatures. These samples demonstrate the capability of the data base and computational model for dealing with phase equilibria in multicomponent oxide systems over a wide range of conditions and compositions of practical interest.

INTRODUCTION

Previous papers in the current series (1-5) provide descriptive information for computing condensed phase equilibria in ceramic systems. In view of current interest in applying ceramic systems in applications requiring toughness and structural performance at high temperatures the present data base is being extended to cover GeO_2 , HfO_2 , ZrO_2 and TiO_2 . This has been effected by employing

available sources (7-9) of thermochemical and phase diagram data. High temperature ceramics have received increased attention during the last few years for structural, thermal protection and engine applications. SIALONS and combinations of zirconia and hafnia with Al_2O_3 , SiC and Si_3N_4 have been shown to develop strength and toughness. This has opened the door to a whole range of new uses for these materials. Recently Slichting and co-workers (10-11) have shown that by alloying GeO_2 with SiO_2 a whole range of glasses can be synthesized with tailor-made coefficients of expansion. Utilization of such compositions offers the possibility of enhancing the high temperature oxidation resistance of ceramic composites in which a mixed GeO_2 - SiO_2 phase with a desired CTE could replace the conventional SiO_2 as a filler. This kind of compositing would open an entire spectrum of new opportunities for synthesis of high temperature ceramics. One of the major obstacles in the development of complex composite systems is the lack of phase diagram information which can be used to guide the fabrication and processing of a new material and help to predict its performance. The current methods of employing models to predict high temperature behaviour has proven useful when basic data is unavailable or too costly and time consuming to obtain by conventional means. This method consists of developing a data base of thermochemical and phase diagram information in analytical form and employing computer models to extend

the description to binary and ternary systems. Recently J. Lorenz et al. (12) applied this method successfully to SiC-ZrO_2 and $\text{SiC-ZrO}_2\text{-Al}_2\text{O}_3\text{-SiO}_2$ in order to evaluate composition effects and identify fabrication conditions. In the present paper, the data base has been expanded by analyzing the following quasi-binary systems: $\text{GeO}_2\text{-HfO}_2$, $\text{GeO}_2\text{-TiO}_2$, $\text{GeO}_2\text{-Al}_2\text{O}_3$, $\text{GeO}_2\text{-MgO}$, $\text{GeO}_2\text{-CaO}$, $\text{GeO}_2\text{-SiO}_2$, $\text{TiO}_2\text{-MgO}$, $\text{HfO}_2\text{-SiO}_2$, $\text{HfO}_2\text{-MgO}$, $\text{HfO}_2\text{-CaO}$, $\text{Al}_2\text{O}_3\text{-HfO}_2$, $\text{HfO}_2\text{-Y}_2\text{O}_3$, HfO-TiO_2 , $\text{Ce}_2\text{O}_3\text{-Al}_2\text{O}_3$, $\text{ZrO}_2\text{-HfO}_2$, $\text{ZrO}_2\text{-SiO}_2$, $\text{ZrO}_2\text{-CaO}$, $\text{Y}_2\text{O}_3\text{-CaO}$ and $\text{Y}_2\text{O}_3\text{-MgO}$. These results when combined with earlier findings (1-5) were employed to compute a range of isothermal sections in the following quasi-ternary systems sufficient to define their characteristics: $\text{MgO-TiO}_2\text{-SiO}_2$, $\text{MgO-SiO}_2\text{-GeO}_2$, $\text{GeO}_2\text{-MgO-CaO}$, $\text{HfO}_2\text{-CaO-MgO}$, $\text{HfO}_2\text{-SiO}_2\text{-ZrO}_2$, $\text{HfO}_2\text{-CaO-Y}_2\text{O}_3$, $\text{HfO}_2\text{-MgO-Y}_2\text{O}_3$, $\text{HfO}_2\text{-CaO-ZrO}_2$, $\text{SiO}_2\text{-HfO}_2\text{Y}_2\text{O}_3$ and $\text{MgO-SiO}_2\text{-HfO}_2$.

LATTICE STABILITY VALUES

Table 1 defines the lattice stability values employed in the current study. Data for the stable forms were taken from Kubaschewski and Alcock (7). The remaining values were adopted along with the lines employed previously. As shown in Table 1 the present analysis is based on one gram atom of compound and/or solution phase. Moreover, dissociation of the components or vaporization is not considered!

TABLE 1
(CONCLUDED)

SUMMARY OF LATTICE STABILITY PARAMETERS

(All units in Joules per gram atom (mole of atoms), T in Kelvin).

P = Periclase, C = Corundum, S = Spinel, X = Crystoballite
Tr = Tridymite, H = Hexagonal (w quartz), R(SO) = Triclinic (p quartz)

R = Rutile (TO) A = Cubic (HO and ZO)
T = Tetragonal (HO and ZO) H = Monoclinic (HO and ZO)
Y = High Temperature YO R = Low Temperature YO
L = Liquid

DODOLR =	- 16.23 T	MONOLH =	- 8.37 T
DODOLA =	8732 - 13.73T	MONOLR =	- 10.42 T
DODOPA =	-31016	MONOPA =	-31016
DODOAM =	16736 - 2.134T	MONOLA =	16682 - 15.387T
DODOAT =	-22594	MONOAM =	16736 - 2.134T
DODOLY =	5439 - 8.36RT	MONOAT =	-22594
DODOLY =	2929 - 6.694T	MONOLR =	13389 - 10.042T
DODOLR =	- 10.41RT	MONOLY =	10878 - 8.36RT
TOTOLR =	14539 - 10.54 T	YOTOLY =	22694 - 8.36RT
TOTOLR =	22313 - 10.42 T	YOTOLR =	26878 - 10.042T
TOTOHR =	7774 - 0.12 T	YOTOLA =	22615 - 9.432T
TOTOLP =	- 8.37 T	YOTOAM =	-16736 - 2.134T
TOTOLT =	3527 - 11.757T	YOTOAT =	-8368 - 1.381T
TOTOLA =	1017 - 11.757T	YOTOLP =	- 12.552T
TOTOTM =	-12259	YOTOLP =	- 10.41RT
TOTOLC =	- 10.208T	ZOTOLA =	28008 - 9.832T
TOTOLX =	- 1.674T	ZOTOAT =	1987 - 0.753T
TOTOLY =	- 8.36RT	ZOZOTM =	2008 - 1.381T
TOTOLR =	6862 - 9.205T	ZOTOLP =	- 8.36RT
AOAOLH =	- 10.54 T	SOSOLR =	-4790 - 5.69 T
AOAOLA =	- 9.832T	SOSOLT =	- 2.092T
AOAOLT =	- 10.585T	SOSOLA =	- 2.092T
AOAOLY =	- 8.36RT	SOSOLR =	- 2.092T
AOAOLR =	- 10.41RT		
REFELC =	- 10.208T		

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L = Liquid

GO = 1/3GeO ₂	HO = 1/3HfO ₂	TO = 1/3TiO ₂	AO = 1/3Al ₂ O ₃
MO = 1/2MgO	DO = 1/2CaO	SO = 1/3SiO ₂	YO = 1/3Y ₂ O ₃
CF = 1/3CeO ₂	ZO = 1/3ZrO ₂		
GOOLH = (1/3)GeO ₂ (liquid) - (1/3)GeO ₂ (hexagonal)			
GOOLP = (1/3)GeO ₂ (liquid) - (1/3)GeO ₂ (rutile)			
GOOLR = (1/3)GeO ₂ (hexagonal) - (1/3)GeO ₂ (rutile)			
GOOLH = GOOLH + GOOLR			
GOOLH =	14644 - 10.54T	MONOLA =	34865 - 11.00 T
GOOLP =	22087 - 16.23T	MONOAT =	2239 - 0.753T
GOOLR =	7481 - 5.69T	MONOLT =	37104 - 11.753T
GOOLA =	- 11.00T	MONOTM =	2724 - 1.381T
GOOLT =	- 11.57T	MONOLM =	39828 - 13.134T
GOOLH =	- 13.14T	MONOLX =	- 2.092T
GOOLC =	- 10.21T	MONOLP =	- 8.36RT
GOOLP =	- 8.37T	MONOLC =	- 10.208T
GOOLR =	7108 - 5.40T	MONOLA =	30711 - 10.376T
GOOLA =	- 1.67T	MONOLY =	23849 - 9.196T
GOOLT =	- 2.01T	MONOLR =	21882 - 10.42 T
GOOLH(SO) =	315 - 0.29T		
GOOLTR =	14644 - 8.54T		

* These differences specify the free energy of one phase (i.e. liquid) minus the free energy of the second phase (i.e. hexagonal) for a compound.

BINARY SYSTEMS

Tables 2-4 and Figures 1-19 summarize the results for the binary systems listed above. The solution phases are described as subregular solutions along the lines of Equations (1) and (2) of reference 4. When the subregular parameters are equal (i.e. GO-HO in Table 2) the solutions are regular, degenerating to ideal solutions for the liquid in HO-MO or all of the HO-ZO phases. The compound phases are defined at fixed compositions (i.e. Equations 3-6 of reference 4) in terms of the compound parameter and the base phase. The latter values are listed in Table 3 for the compounds of interest while Table 4 shows the Gibbs energy of formation for these compounds from the component oxides. Figure 1 shows the calculated GO-HO($1/2\text{GeO}_2$ - $1/2\text{HfO}_2$) system derived from the description contained in Tables 1-4. This phase diagram was taken to compare with the analogous (SO-HO) case given by Figure 4443 (8). The system exhibits very little mutual solubility in the solid phase and an equimolar compound $\text{GO}_{.5}\text{HO}_{.5}$. The GO-TO case shown in Figure 2 is based on the experimental phase diagram in Figure 358 (8). The system is characterized by a miscibility gap ($R_1 + R_2$) in the rutile solid solution. The GO-AO system in Figure 3 is based on the experimental diagram given in Figure 4372 (8) and the earlier analog SO-AO analysis (2). This system exhibits little solubility in the solid and a single compound analogous to the SO-AO rutile. The calculated GO-MO phase diagram is shown in Figure 4 and is

based on the experimental phase diagram in Figure 264 (8) and the analogous SO-MO results (2). This system exhibits a liquid miscibility gap and a number of compound phases which have similar stoichiometry to the SO-MO compounds Enstatite (E) and Fosterite (F). The GO-DO ($1/3\text{GeO}_2$ - $1/2\text{CaO}$) case shown in Figure 5 is based on the experimental diagram given in Figure 4309 (8). Figure 6 shows the GO-SO phase diagram ($1/3\text{GeO}_2$ - $1/3\text{SiO}_2$) which is based upon the experimental diagram in Figure 357 (8). The latter is restricted to the low temperature range, however it is sufficient to permit definition of the solution phase parameters listed in Table 2. The system is distinguished by a large range of solid solution within the H phase which is a stable form of GeO_2 and SiO_2 . This continuous solid solution is undoubtedly the basis of Slichting's finding that permitted a continuous variation in the coefficient of thermal expansion between the low values characteristics of SiO_2 to the GeO_2 value, more compatible with the expansion coefficients exhibited by metals. The TO-MO diagram shown in Figure 7 is based on the experimental diagram given in Figure 4336 (8). The SO-HO phase diagram calculated on the basis of the parameters listed in Tables 1-3 and shown in Figure 8 exhibits little solubility in the solid and one equimolar compound in keeping with Figure 4443 (8). The $1/3\text{HfO}_2$ - $1/2\text{MgO}$ (HO-MO) case displayed in Figure 9 shows some solubility in the cubic structure based on HfO_2 and no compounds. This phase diagram was defined on the basis

TABLE 2
(Continued)

QUASIBINARY SOLUTION PARAMETERS FOR OXIDE SYSTEMS (All units in Joules per gram atom (mole of atoms), T in Kelvins)		QUASIBINARY SOLUTION PARAMETERS FOR OXIDE SYSTEMS (All units in Joules per gram atom (mole of atoms), T in Kelvins)	
LGHO = 16736	LHGO = 16736	LSGO = 4184 + 12.552T	LGSO = 4184 + 12.552T
AGHO = 62760	AHGO = 62760	HSOG = 4184 + 12.552T	HGSO = 4184 + 12.552T
TGOH = 62760	THGO = 62760	RSOG = 4184 + 12.552T	RGSO = 4184 + 12.552T
MGHO = 62760	MHGO = 62760	XSGO = 83680	XGSO = 83680
HGOH = 62760	HGOH = 62760	TSOG = 83680	TGSO = 83680
RGHO = 62760	RHGO = 62760	RSOG = 41840	RGSO = 41840
LGTO = 22292	LTGO = - 5908 + 3.347T	LTMO = - 45187 + 26.36T	LMTO = 1883 - 11.39T
MGTO = 18828	MTGO = - 9372 + 3.347T	PTMO = 83680	PMTO = 83680
PGTO = 52300 - 20.92T	RTGO = 19246	RTMO = 41840	RMTO = 41840
LGHO = 21221	LAHO = 21221	LHOS = 46024	LSHO = 46024
HGOH = 41840	HAHO = 41840	XHOS = 62760	XSHO = 62760
RGHO = 41840	RAHO = 41840	AMOS = 125520	ASHO = 125520
CGHO = 83680	CAHO = 83680	THOS = 125520	TSHO = 125520
		MHOS = 125520	MSHO = 125520
LGHO = - 29288 + 18.41T	LMGO = - 216731 + 60.25T	LMHO = 0	LMHO = 0
$0 \leq X_{MO} \leq 0.4$	$0 \leq X_{MO} \leq 0.4$	AMHO = 8368	AMHO = - 29288
LGMO = - 104265 + 35.15T	LMGO = - 104265 + 35.15T	THMO = 8368	TMHO = - 29288
$0.4 \leq X_{MO} \leq 1.0$	$0.4 \leq X_{MO} \leq 1.0$	MHMO = 46024	MMHO = 46024
PGMO = - 20585 + 35.15T	PMGO = - 20585 + 35.15T	PHMO = 83680	PMHO = 83680
HGMO = - 12552 + 18.41T	HMGO = - 12.552 + 60.25T		
LGDO = - 43179 + 20.92T	LDGO = - 226020 + 62.76T	LMDO = - 25104 + 9.20T	LDHO = - 112968 + 41.84T
$0 \leq X_{DO} \leq 0.4$	$0 \leq X_{DO} \leq 0.4$	ADDO = - 12552	ADHO = - 142256 + 41.84T
LGDO = - 116315 + 37.66T	LDGO = - 116315 + 37.66T	THDO = 8368	TDHO = - 29288
$0.4 \leq X_{DO} \leq 1.0$	$0.4 \leq X_{DO} \leq 1.0$	MHDO = 46024	MDHO = 46024
PGDO = 62.760	PDGO = 62.760	PHDO = 83680	PMHO = 83680
HGDO = - 43179 + 20.92T	HDGO = - 226020 + 62.76T	LHDO = 17573	LAHO = 39748
LYDO = 27196	LDYO = 27196	AHDO = 62760	AAHO = 62760
TYDO = 33564	TDYO = 33564	THDO = 62760	TAHO = 62760
RYDO = 35564	RDYO = 35564	MHDO = 62760	MAHO = 62760
PYDO = 81680	PDYO = 81680	CHDO = 62760	CAHO = 62760

TABLE 2
(Continued)

QUANTITATIVE SOLUTION PARAMETERS FOR OXIDE SYSTEMS
(All units in Joules per gram atom (mole of atoms), T in Kelvin)

LYOYO	-3975	-2.51T	LYOMO	-16527	+3.67T
TYOYO	-2510	+7.11T	TYOMO	-11715	+0.618T
MYOYO	-8168		MYOMO	-10042	+8.786T
AYOYO	-1674	+7.11T	AYOMO	-10042	+8.79T
YMOYO	-7113	+8.168T	YMOYO	-7113	+8.168T
BYOYO	-2243	+11.96T	BYOMO	-2243	+11.97T
LYOTO	-8168		LYOTO	-8368	
AYOTO	-12552		AYOTO	-12552	
TYOTO	-12552		TYOTO	-12552	
MYOTO	-12552		MYOTO	-12552	
BYOTO	-8168		BYOTO	-8368	
LYEAO	-27196		LYEAO	-27196	
TYEAO	-81680		TYEAO	-81680	
MYEAO	-81680		MYEAO	-81680	
LYMOO	0		LYMOO	0	
AYMOO	0		AYMOO	0	
TYMOO	0		TYMOO	0	
MYMOO	0		MYMOO	0	
LYMOO	-84040	-15.61T	LYMOO	-84040	-15.61T
AYMOO	-125520		AYMOO	-125520	
TYMOO	-125520		TYMOO	-125520	
MYMOO	-62760		MYMOO	-62760	
BYMOO	-125520		BYMOO	-125520	
LYMOO	-37676	+16.716T	LYMOO	-40468	+16.842T
AYMOO	-81680		AYMOO	-81680	
TYMOO	-54392	+16.716T	TYMOO	-54392	+16.716T
MYMOO	-66107	+16.716T	MYMOO	-66107	+16.716T
BYMOO	-16716	+16.716T	BYMOO	-16716	+16.716T
LYMOO	-27196		LYMOO	-27196	
AYMOO	-35564		AYMOO	-35564	
BYMOO	-35564		BYMOO	-35564	

TABLE 3

SUMMARY OF COMPOUND PARAMETERS FOR BINARY SYSTEMS
(All units in Joules per gram atom (mole of atoms), T in Kelvin)

Compound	Name	Stoichiometry	Stability	Base	Compound/Parameter (Joules/g.at.)
(1/6)(CaO ₂ .MgO ₂)	S	GO ₂ MO ₂ .5	stable	M	66044
(1/21)(2CaO ₂ .3Al ₂ O ₃)	M	GO ₂ MO ₂ .714	stable	C	64634 - 5.21T
(1/5)(MgO.CeO ₂)	E	GO ₂ MO ₂ .4	stable	P	-30719 + 60.75T
(1/7)(CaO ₂ .2MgO)	P	GO ₂ MO ₂ .571	stable	P	-53631 + 46.60T
(1/11)(CaO ₂ .2MgO)	X	GO ₂ MO ₂ .727	stable	P	-61137 + 42.37T
(1/14)(4CaO ₂ .CeO)	M	GO ₂ MO ₂ .143	stable	P	-20108 + 75.00T
(1/8)(2CaO ₂ .CeO)	R	GO ₂ MO ₂ .25	stable	P	-33126 + 76.10T
(1/5)(CaO.CeO ₂)	E	GO ₂ MO ₂ .4	stable	P	-50320 + 62.43T
(1/12)(2CaO ₂ .3CeO)	T	GO ₂ MO ₂ .5	stable	P	-69287 + 61.30T
(1/7)(CaO ₂ .2CeO)	P	GO ₂ MO ₂ .571	stable	P	-74634 + 61.55T
(1/9)(CaO ₂ .3CeO)	V	GO ₂ MO ₂ .667	stable	P	-84701 + 60.00T
(1/8)(2TiO ₂ .MgO)	W	TO ₂ MO ₂ .25	stable	P	54811 + 5.46T
(1/5)(TiO ₂ .MgO)	V	TO ₂ MO ₂ .4	stable	P	48606 + 3.01T
(1/7)(TiO ₂ .2MgO)	U	TO ₂ MO ₂ .571	stable	P	10593 + 0.04T
(1/4)(SiO ₂ .MgO ₂)	S	SO ₂ MO ₂ .5	stable	M	64601 - 4.10T
(1/25)(TMFO ₂ .2CaO)	P	MO ₂ MO ₂ .16	stable	A	35564 + 10.00T
(1/5)(MgO ₂ .CeO)	O	MO ₂ MO ₂ .4	stable	A	9701 + 32.22T
(1/4)(MgO ₂ .TiO ₂)	S	MO ₂ MO ₂ .5	stable	T	30074 + 17.00T
(1/10)(Ce ₂ O ₃ .Al ₂ O ₃)	P	CF ₂ MO ₂ .5	stable	T	61530 + 7.24T
(1/30)(CeAl ₁₁ O ₁₉)	O	CF ₂ MO ₂ .917	stable	C	20610 + 25.00T
(1/4)(TiO ₂ .SiO ₂)	O	TO ₂ MO ₂ .5	stable	M	102600 - 22.51T
(1/14)(4TiO ₂ .CeO)	V	TO ₂ MO ₂ .143	stable	A	30125 + 17.20T
(1/5)(2TiO ₂ .CeO)	O	TO ₂ MO ₂ .4	stable	A	21664 + 20.07T

TABLE 5

SUMMARY OF CONTEMPORARY STOICHIOMETRY AND PARAMETERS EMPLOYED IN TERNARY CALCULATIONS			
Stable Phase (Name)	Base	Counterphase	Counterphase Parameter (Joule/g. at 1)
MC-TO-CO	CO, MC, 6 (1) P	CO, TO, 6	0 (ICAB = -71320)
	CO, 420 MC, 571 (1) P	CO, 420 TO, 571	0 (ICAB = -61000)
	CO, 273 MC, 727 (1) P	CO, 273 TO, 727	0 (ICAB = -72547)
	TO, 34 MC, 25 (1) S	TO, 34 CO, 25	0 (ICAB = -71100)
	TO, 6 MC, 4 (1) P	TO, 6 CO, 4	0 (ICAB = -50556)
	TO, 420 MC, 571 (1) P	TO, 420 CO, 571	0
CO-MC-SO	CO, MC, 4 (1) P	MC, SO, 4	-68735 + 67.207
	CO, 420 MC, 571 (1) P	MC, 571 SO, 420	-66334 + 61.357
	CO, 273 MC, 727 (1) P	MC, 727 SO, 273	-77022 + 61.007
CO-MC-BO	CO, 851 MC, 143 (1) S	MC, 143 CO, 851	0
	CO, 34 MC, 25 (1) P	MC, 34 CO, 25	0
	MC, 6 MC, 4 (1) P	MC, 6 CO, 4	-30710 + 60.757
	CO, 5 MC, 5 (1) P	MC, 5 CO, 5	0
	CO, 420 MC, 571 (1) P	MC, 571 CO, 420	-53631 + 60.607
	CO, 333 MC, 667 (1) P	MC, 667 CO, 333	0
	CO, 273 MC, 727 (1) P	MC, 727 CO, 273	0
SO-TO-MC	SO, 3 MC, 5 (1) S	TO, 5 MC, 3	102600 + 23.517
MC-MC-MC	MC, 80 MC, 16 (1) A	MC, 80 MC, 16	0
	MC, 6 MC, 4 (1) A	MC, 6 MC, 4	0
MC-MC-MC	MC, 80 MC, 16 (1) A	TO, 80 MC, 16	0
	MC, 6 MC, 4 (1) A	TO, 6 MC, 4	0
TO-MC-BO	TO, 851 MC, 143 (1) A	MC, 851 MC, 143	0
	MC, 80 MC, 16 (1) A	TO, 80 MC, 16	0
	MC, 6 MC, 4 (1) A	TO, 6 MC, 4	23664 + 20.077
MC-MC-TO	MC, 5 MC, 5 (1) S	TO, 5 MC, 5	0
	TO, 420 MC, 571 (1) P	TO, 420 MC, 571	0
	TO, 273 MC, 727 (1) P	TO, 273 MC, 727	0
	TO, 651 MC, 343 (1) P	TO, 651 MC, 343	0
MC-MC-SO	MC, 5 MC, 5 (1) S	SO, 5 MC, 5	0
	MC, 6 MC, 4 (1) P	SO, 6 MC, 4	0
	MC, 571 MC, 420 (1) P	SO, 571 MC, 420	0

TABLE 6

CALCULATED FREE ENERGY OF FORMATION OF COMPOUND PHASES (all units in Joules per gram atom (mole of atoms) T in Kelvin)		
Compound	Base Stoichiometry	Free Energy of Formation, ΔG_f° (2000)
From Component Compounds		
		-1506 + 1.557
(1/6)(CaO ₂ , MgO ₂)	CO, 5 MC, 5	-2080 + 0.657
(1/21)(2CaO ₂ , 1Al ₂ O ₃)	MC, 208 MC, 314	-2210 + 0.237
(1/5)(MgO, CaO ₂)	CO, MC, 4	-2020 + 0.207
(1/7)(CaO ₂ , 2MgO)	CO, 420 MC, 571	-2510 + 1.507
(1/11)(CaO ₂ , 2MgO)	CO, 273 MC, 727	-351 + 0.367
(1/10)(4CaO ₂ , CaO)	CO, 851 MC, 143	-510 + 0.007
(1/8)(2CaO ₂ , CaO)	CO, 34 MC, 25	-660 + 10.677
(1/3)(CaO, CaO ₂)	CO, 6 MC, 4	-711 + 0.877
(1/12)(2CaO ₂ , 3CaO)	CO, 5 MC, 5	-741 + 0.257
(1/9)(CaO ₂ , 3CaO)	CO, 420 MC, 571	310 + 7.777
(1/8)(CaO ₂ , 3CaO)	CO, 333 MC, 667	-6045 + 0.077
(1/8)(2TiO ₂ , MgO)	TO, 34 MC, 25	-6602 + 0.547
(1/5)(TiO ₂ , MgO)	TO, MC, 4	-6602 + 0.547
(1/5)(TiO ₂ , MgO)	TO, MC, 4	-6602 + 0.547
(1/21)(TiO ₂ , 2MgO)	TO, 420 MC, 571	-3300 + 0.107
(1/6)(SiO ₂ , MgO ₂)	SO, 5 MC, 5	-2121 + 0.547
(1/25)(2MgO ₂ , 3CaO)	MC, 80 MC, 16	-912 + 1.707
(1/5)(MgO ₂ , CaO)	MC, 6 MC, 4	-1427 + 1.677
(1/6)(MgO ₂ , TiO ₂)	MC, 5 MC, 5	-040 + 0.527
(1/10)(CaO ₂ , Al ₂ O ₃)	CO, 5 MC, 5	3216 + 2.277
(1/10)(CaO, Al ₂ O ₃)	CO, MC, 4	2188 + 1.077
(1/6)(2CaO ₂ , SiO ₂)	CO, 5 MC, 5	-2121 + 1.227
(1/10)(2CaO ₂ , CaO)	CO, 851 MC, 143	-054 + 1.007
(1/5)(2CaO ₂ , CaO)	CO, 34 MC, 25	-2121 + 1.107

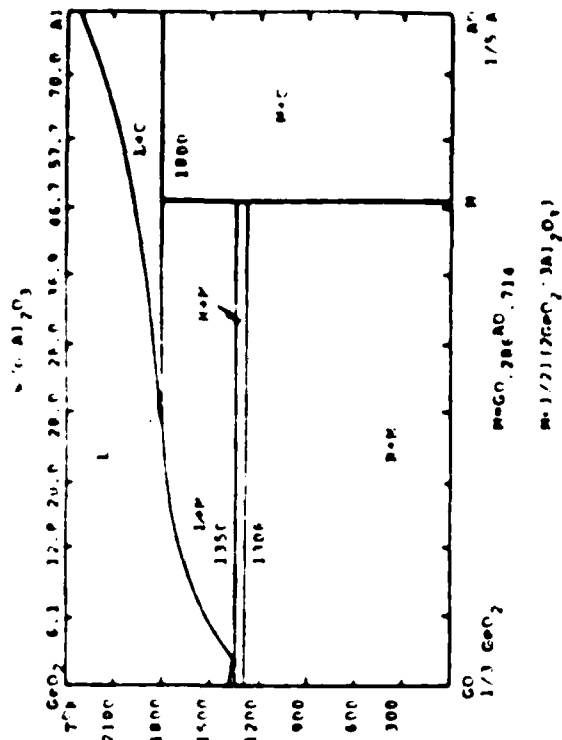


Figure 3. Calculated $\text{GeO}_2\text{-Al}_2\text{O}_3$ Phase Diagram.

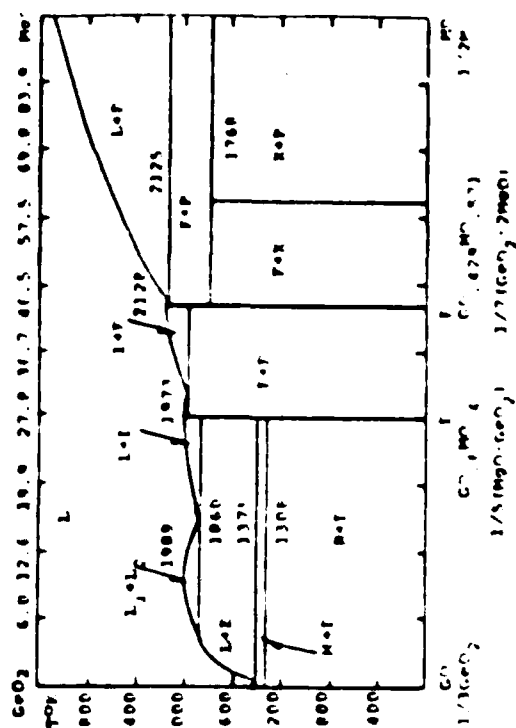


Figure 4. Calculated $\text{GeO}_2\text{-MnO}$ Phase Diagram.

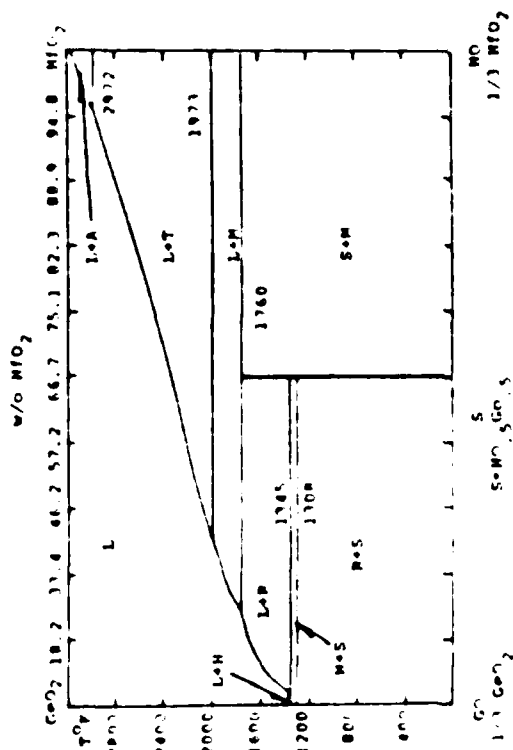


Figure 5. Calculated $\text{GeO}_2\text{-MnO}_2$ Phase Diagram.

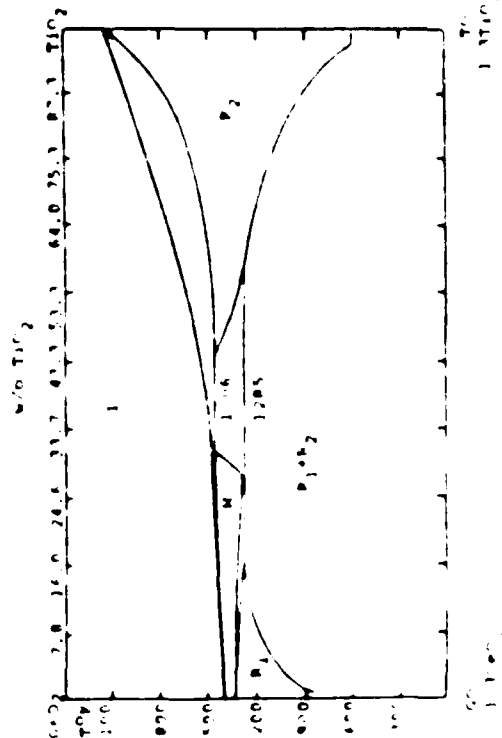
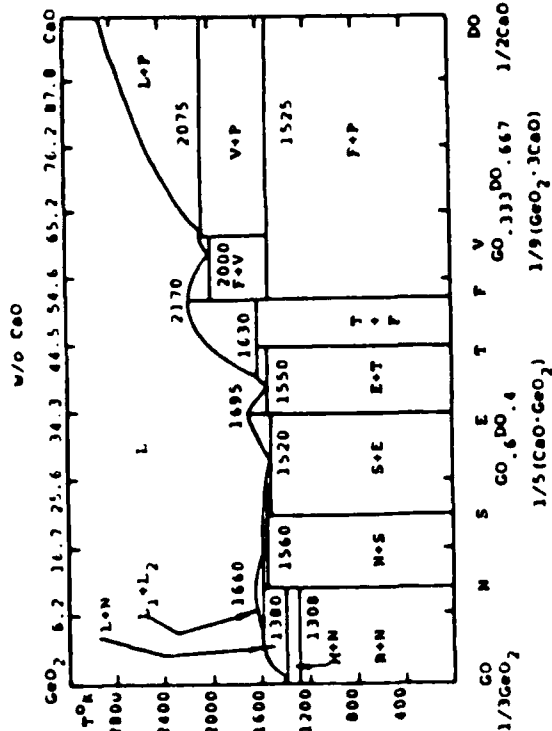


Figure 6. Calculated $\text{GeO}_2\text{-TiO}_2$ Phase Diagram.



$\text{M}=\text{GeO}_{.857}\text{DO}_{.143}=1/14(4\text{GeO}_2\cdot\text{CeO})$ $\text{T}=\text{GeO}_{.5}\text{DO}_{.5}=1/12(2\text{GeO}_2\cdot3\text{CeO})$
 $\text{S}=\text{GeO}_{.750}\text{DO}_{.250}=1/8(2\text{GeO}_2\cdot\text{CeO})$ $\text{F}=\text{GeO}_{.429}\text{DO}_{.571}=1/7(\text{GeO}_2\cdot2\text{CeO})$

Figure 5. Calculated $\text{GeO}_2\text{-CeO}_2$ Phase Diagram.

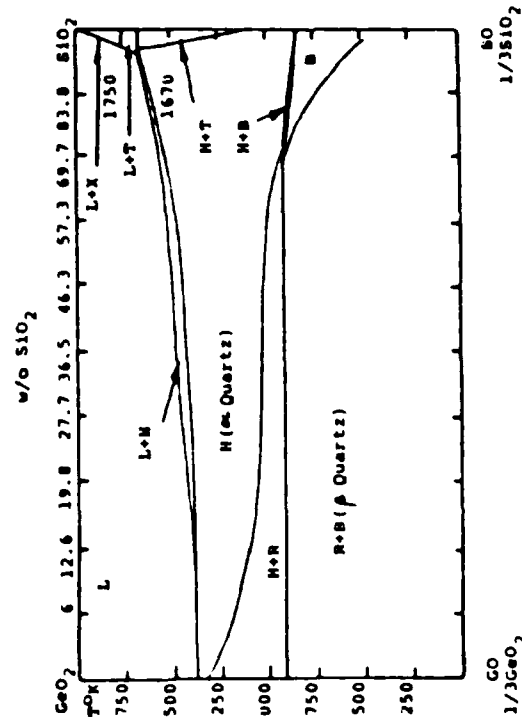


Figure 6. Calculated $\text{GeO}_2\text{-SiO}_2$ Phase Diagram.

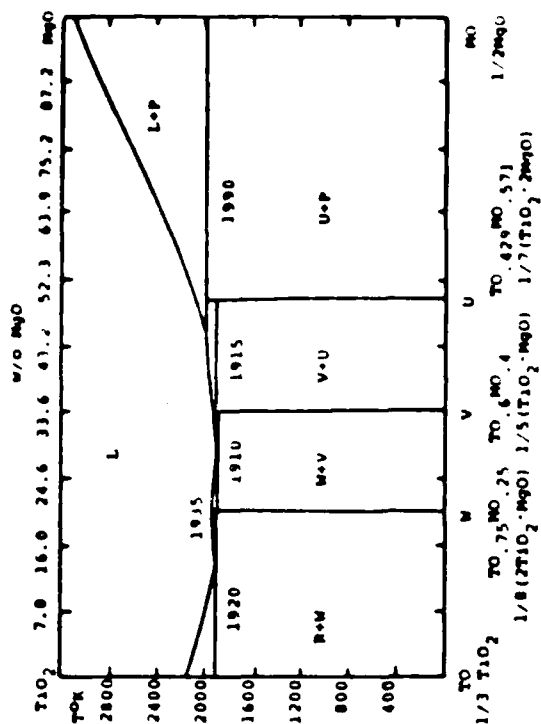


Figure 7. Calculated $\text{TiO}_2\text{-MgO}$ Phase Diagram.

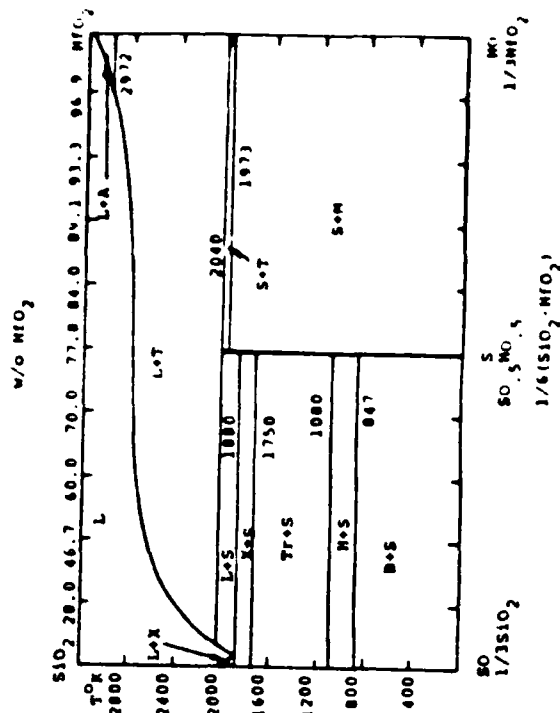


Figure 8. Calculated $\text{SiO}_2\text{-MfO}_2$ Phase Diagram.

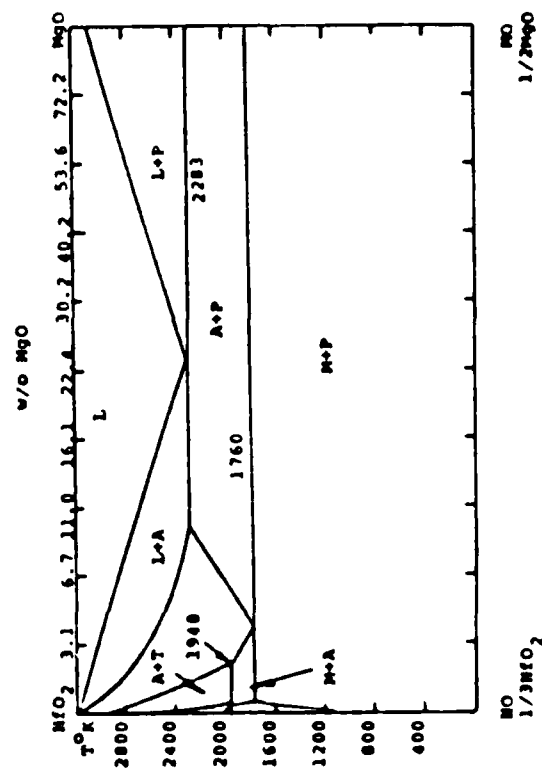


Figure 9. Calculated MnO_2 - MgO Phase Diagram.

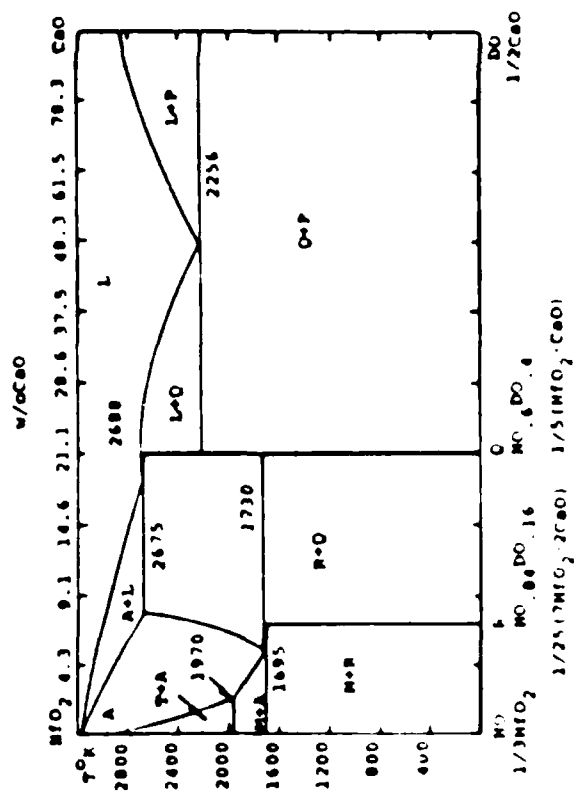


Figure 10. Calculated MnO_2 - CeO Phase Diagram.

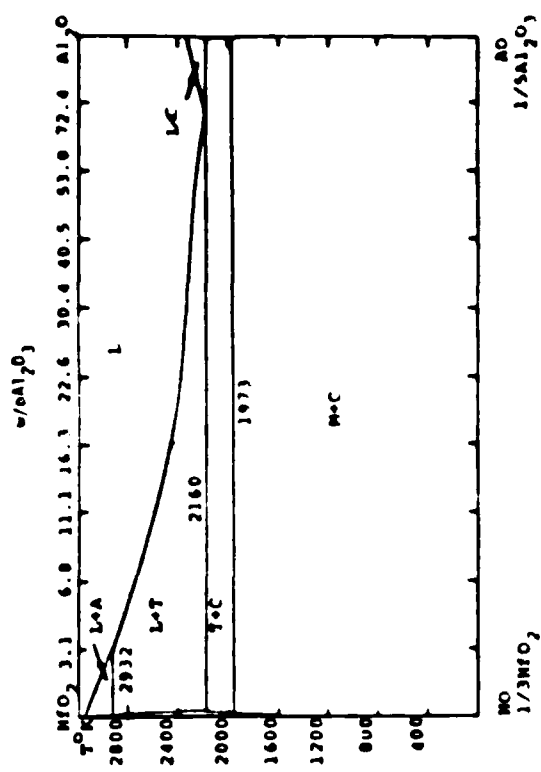


Figure 11. Calculated MnO_2 - Al_2O_3 Phase Diagram.

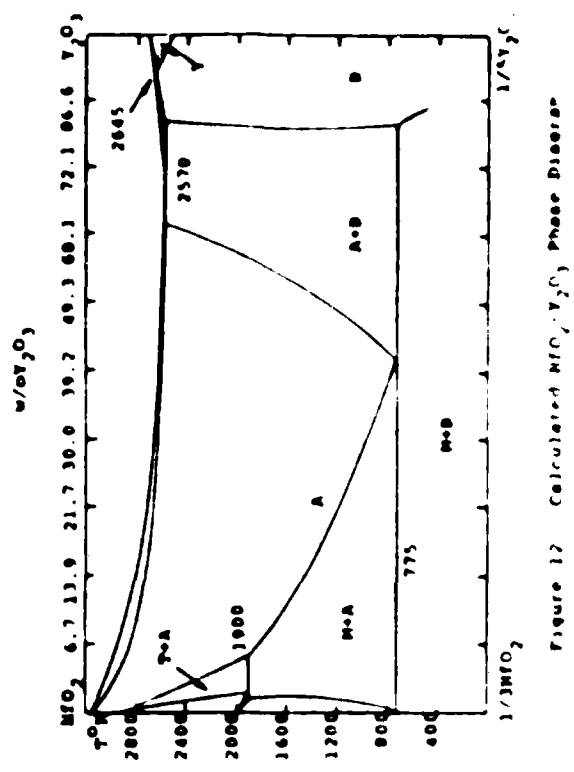


Figure 12. Calculated MnO_2 - V_2O_5 Phase Diagram.

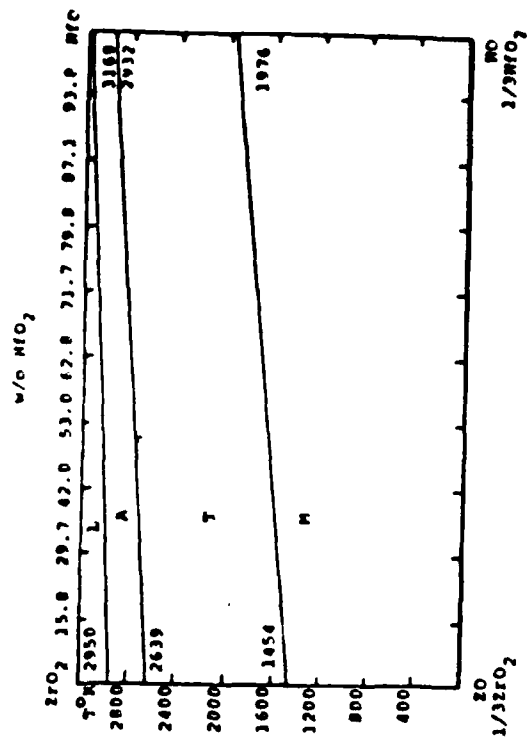


Figure 13. Calculated $\text{MnO}_2\text{-TiO}_2$ Phase Diagram.

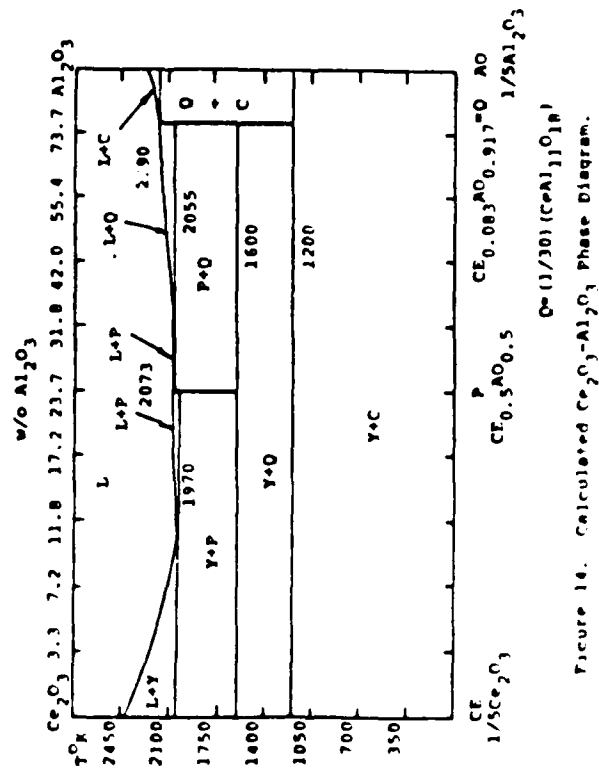


Figure 14. Calculated $\text{Ce}_2\text{O}_3\text{-Al}_2\text{O}_3$ Phase Diagram.

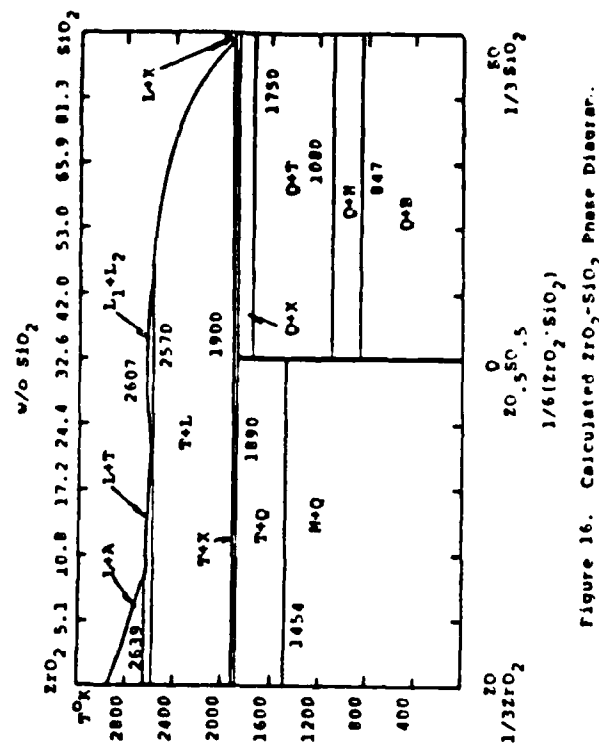


Figure 15. Calculated $\text{ZrO}_2\text{-MfO}_2$ Phase Diagram.

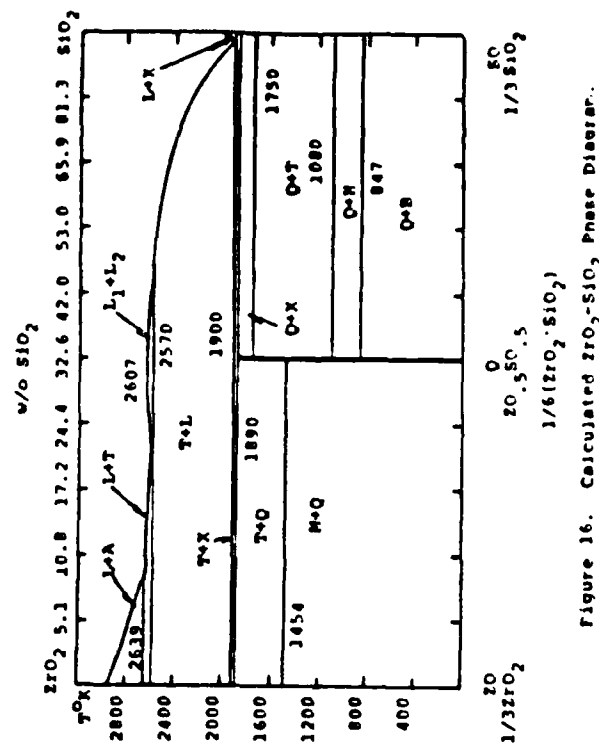


Figure 16. Calculated $\text{ZrO}_2\text{-SiO}_2$ Phase Diagram.

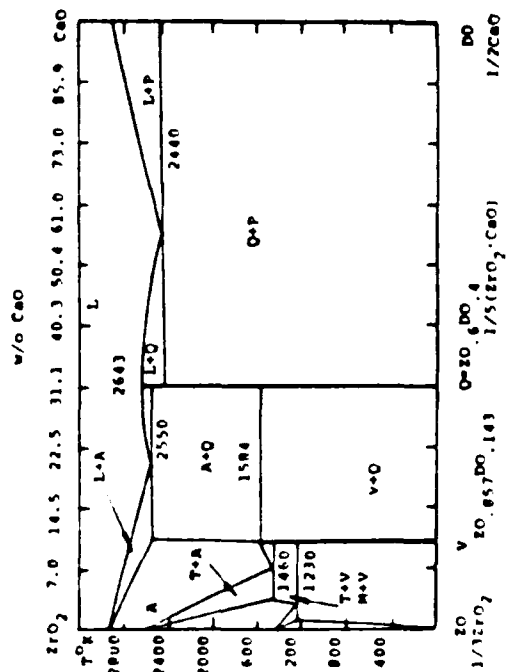


Figure 17. Calculated $2rO_2$ -CaO Phase Diagram.

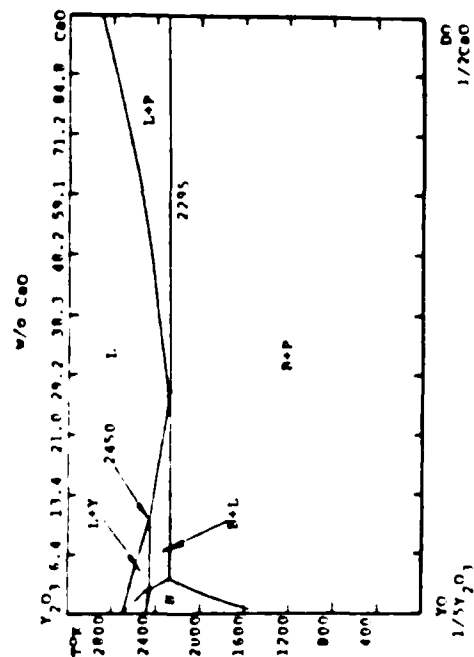


Figure 18. Calculated Y_2O_3 -CaO Phase Diagram.

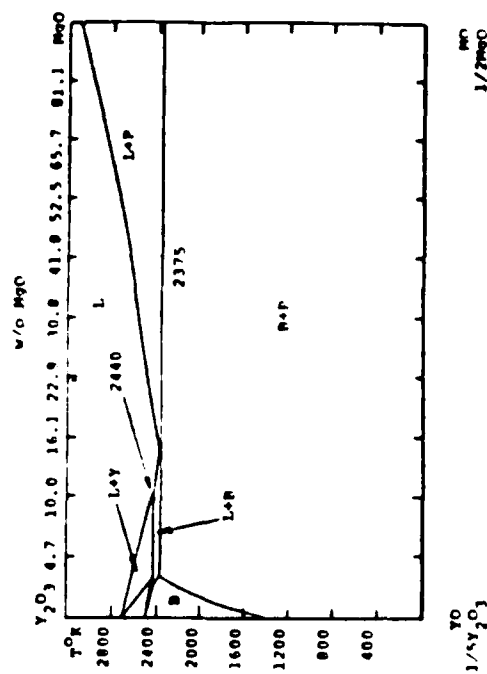


Figure 19. Calculated Y_2O_3 -MgO System.

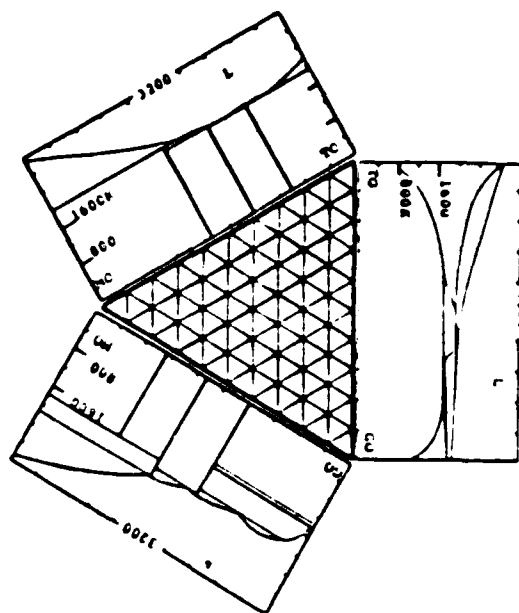


Figure 20. Calculated isothermal sections in the MgO-Y $_2$ O $_3$ -TiO $_2$ -(MgO+TiO $_2$) system.

of Figure 5159 (9). By contrast the $\text{HfO}_2\text{-CaO}$ system in Figure 10, while similar to the $\text{HfO}_2\text{-MgO}$ case exhibits two compounds designated R and Q. This phase diagram is based on the work of Senft and Stubican (13). The HO-AO case in Figure 11 is based on the analogous ZO-AO system, Figure 4378 (8), showing little solubility in the solid phase and no compounds. A very extensive range of solubility in both the cubic structure based on HfO_2 and that based on Y_2O_3 is illustrated in the HO-YO case shown in Figure 12 which is derived from Figure 4436 (8). Inspection of Table 2 shows that this system is defined by solution phase parameters which tend to be positive over most of the temperature range of interest. This is in keeping with the absence of compound phases and the observation that the dominant A solid solution phase decomposes. This will be discussed further in the calculation of the $\text{SiO}_2\text{-HfO}_2\text{-Y}_2\text{O}_3$ system in the next section of this paper. Figure 13 shows the $\text{HfO}_2\text{-TiO}_2$ system based upon the analogous ZO-TO case shown in Figure 4452 (8). This system exhibits solid phase solubility in the cubic phase based on HfO_2 and the rutile phase based on TiO_2 . In addition an equimolar compound in this system is stable and melts congruently. Figure 14 shows the calculated $\text{CE-AO}(1/5\text{Ce}_2\text{O}_3\text{-}1/5\text{Al}_2\text{O}_3)$ system which is based on Figures 356 and 4366 in reference (8). The system contains two compounds designated as F and G in Figure 14 which decomposes into Ce_2O_3 and Al_2O_3 at low temperature. The $\text{ZrO}_2\text{-HfO}_2$ system is shown in Figure 15 which is based on Figure 4444 (8). The calculated phase

diagram is based on assuming ideal solutions for each of the solution phases as can be seen in Table 2. Figure 16 displays the calculated ZO-SO phase diagram which is derived from Figure 2400 (8) and the above noted HO-SO analysis. This system is characterized by a symmetrical miscibility gap, little solubility in the solid phases and an equimolar compound which decomposes prior to melting into tetragonal ZrO_2 and crystoballite. The ZO-DO system shown in Figure 17 is based on Figure 5392 (9). It is quite similar to the HO-DO system in Figure 18 with two compound phases and some solid phase solubility especially in the cubic phase. The YO-DO and YO-MO diagrams shown in Figures 18 and 19 complete the set of systems considered here. The $\text{YO-MO}(1/5\text{Y}_2\text{O}_3\text{-}1/2\text{MgO})$ case in Figure 19 is based on Figure 5156 (9). Neither system contains compounds or substantial solubility in the solid. Examination of Table 2 shows that most of the solution phases exhibit small interaction parameters except for those cases where limited solubility occurs leading to large positive interaction parameters.

TERNARY SYSTEMS

The description of the foregoing systems combined with those presented earlier (1-5) have been employed to calculate isothermal sections in ten quasi-ternary systems over a wide range of temperatures. The results of these calculations are presented in Figures 20-31. In each case the component quasi-binary systems are

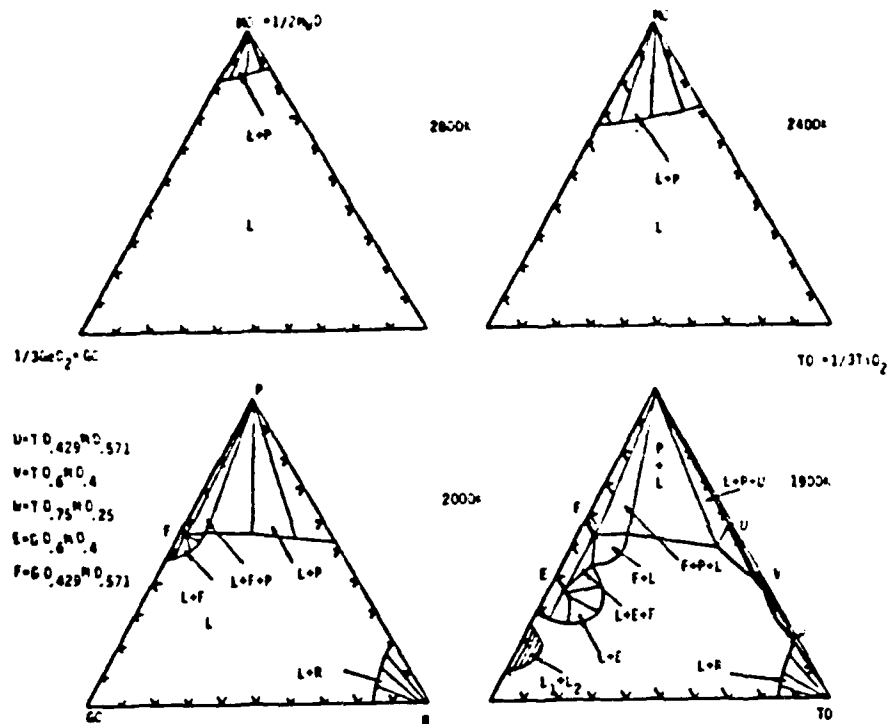


Figure 21 Calculated Isothermal Sections in MD-TD-GC

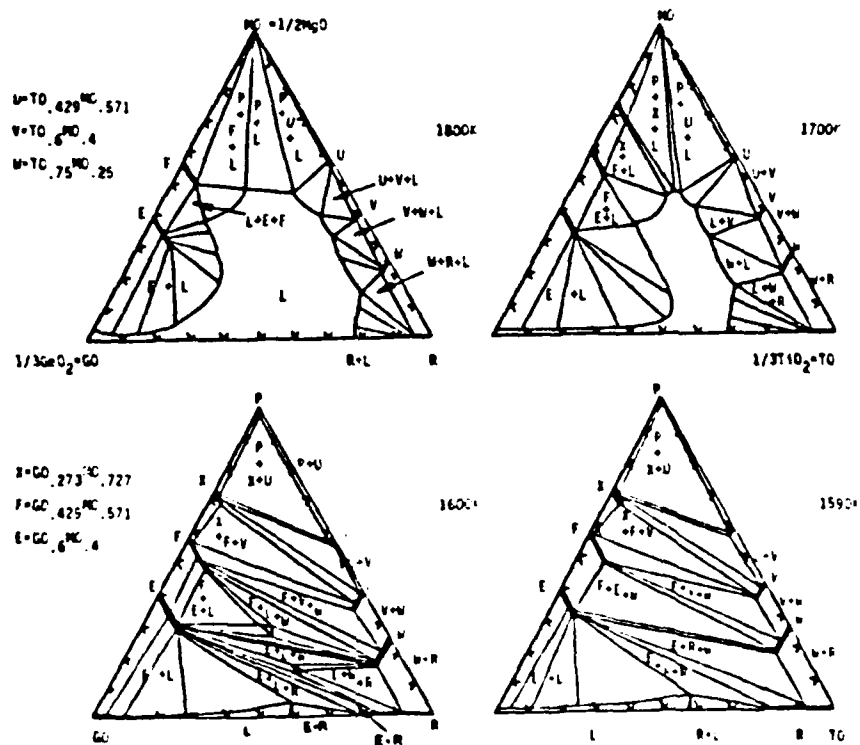


Figure 22 Calculated Isothermal Sections in MC-TD-GC

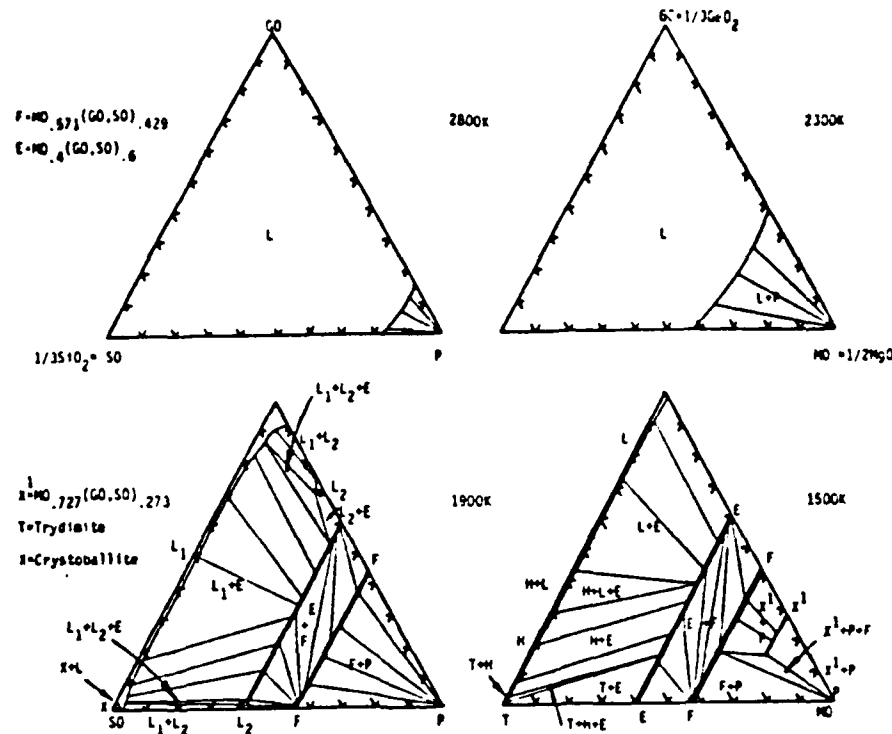


Figure 25. Calculated Isothermal Sections in GeO-MgO-SiO₂

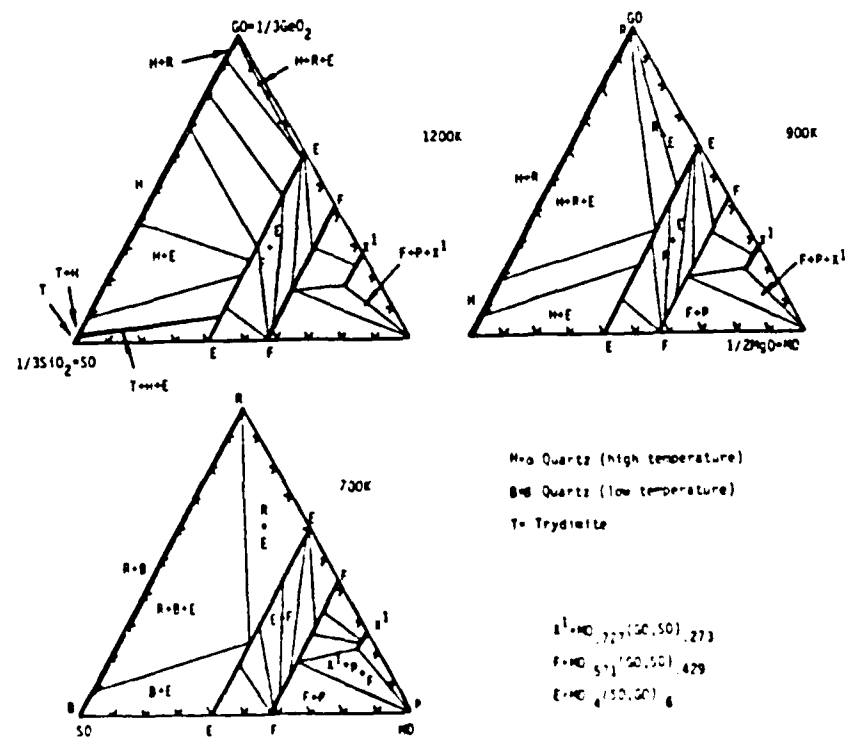


Figure 26. Calculated Isothermal Sections in SiO₂-MgO-SiO₂

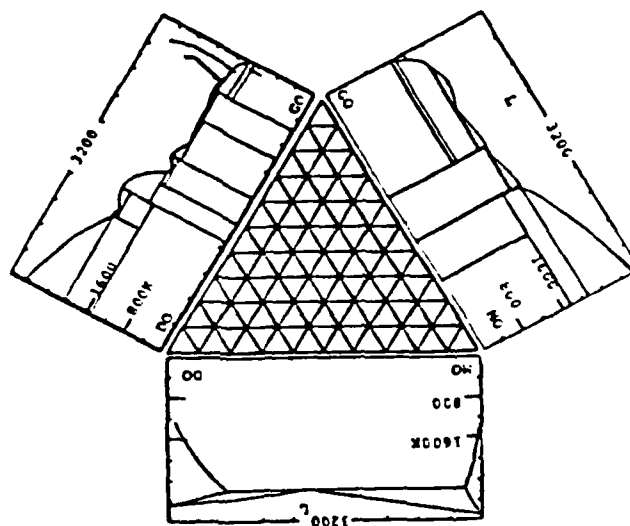


Figure 27 Calculated Isothermal Sections in the GO (1/3 GeO_2) - MO (1/2 MgO) - DO (1/2 CaO) System.

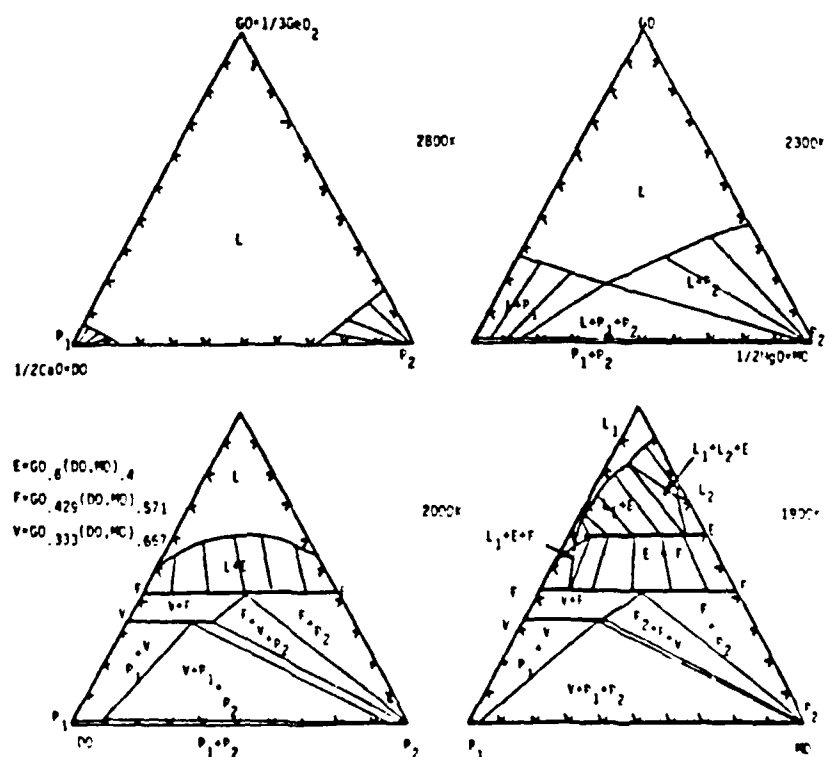


Figure 28 Calculated Isothermal Sections in GO-MO-DO

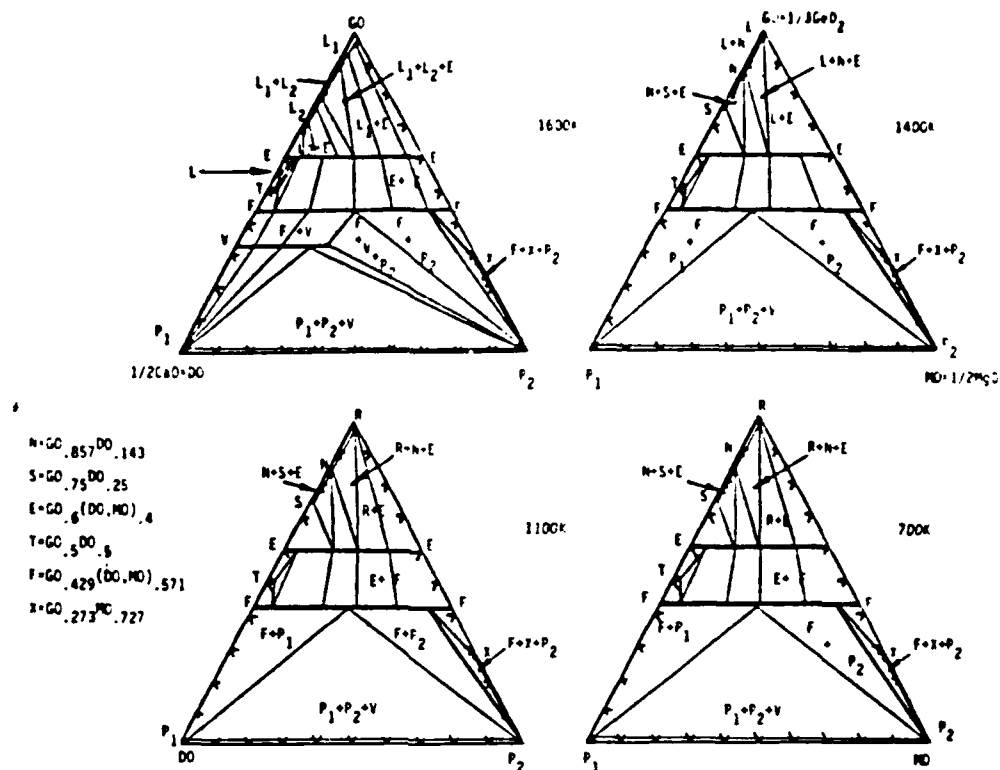


Figure 29. Calculated Isothermal Sections in GO-MO-DO

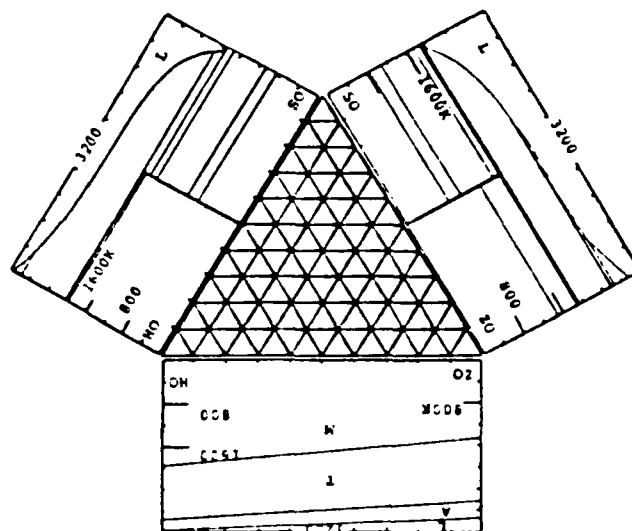


Figure 30. Calculated Isothermal Sections in the $\text{Si}(1-35x\text{O}_2)-20(1-32x\text{O}_2)-\text{MO}(1-3)\text{HfO}_2$ System

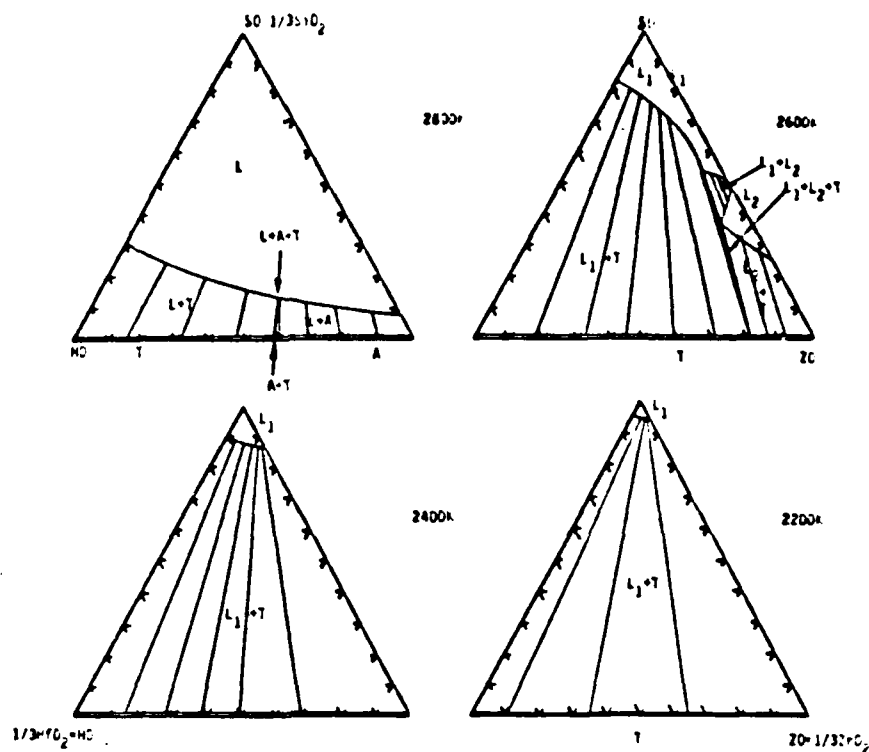


Figure 31. Calculated Isothermal Sections in SO-ZO-MO

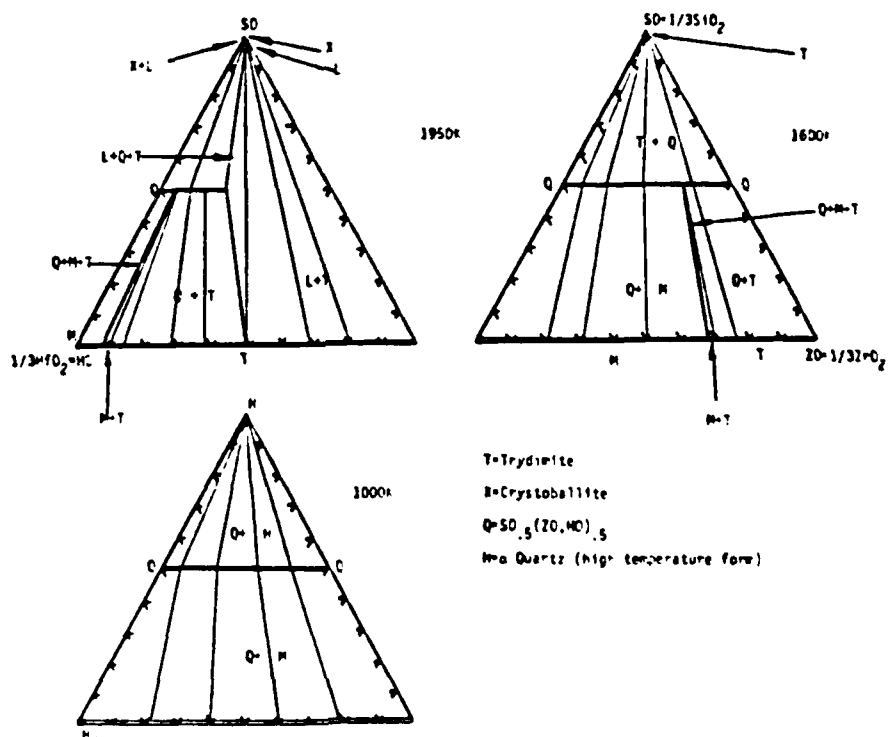


Figure 37. Calculated Isothermal Sections in SO-ZO-MO

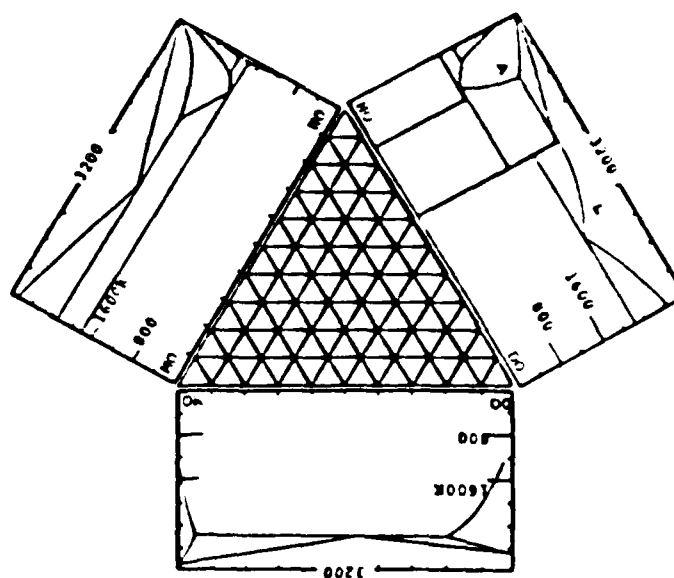


Figure 33. Calculated Isothermal Sections in the $\text{MO}(1/3 \text{MnO}_2) - \text{DC}(1/2 \text{CaO}) - \text{MO}(1/2 \text{MgO})$

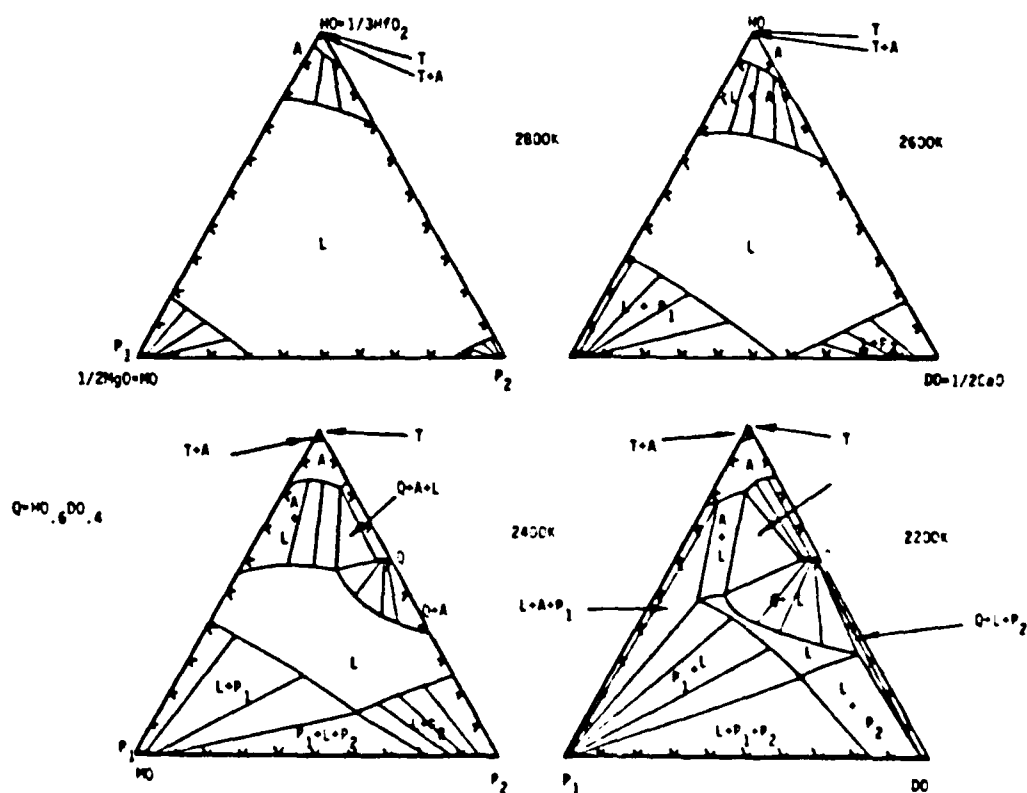


Figure 34. Calculated Isothermal Sections in $\text{MO}-\text{DO}-\text{Mn}$

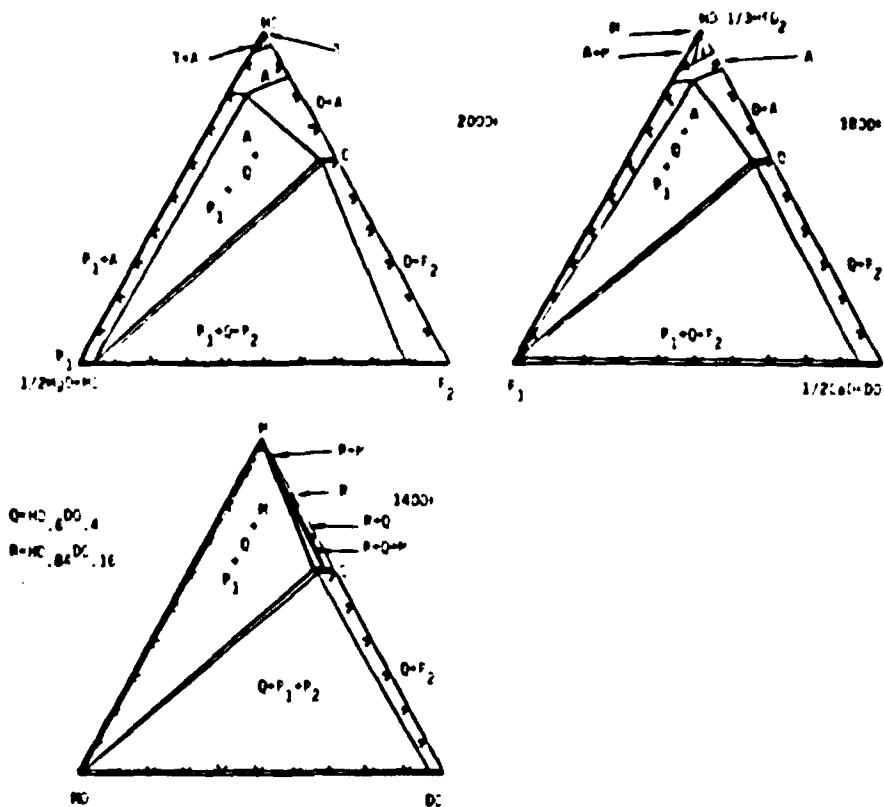


Figure 35. Calculated Isothermal Sections in $\text{MgO}-\text{DC}-\text{Mg}$

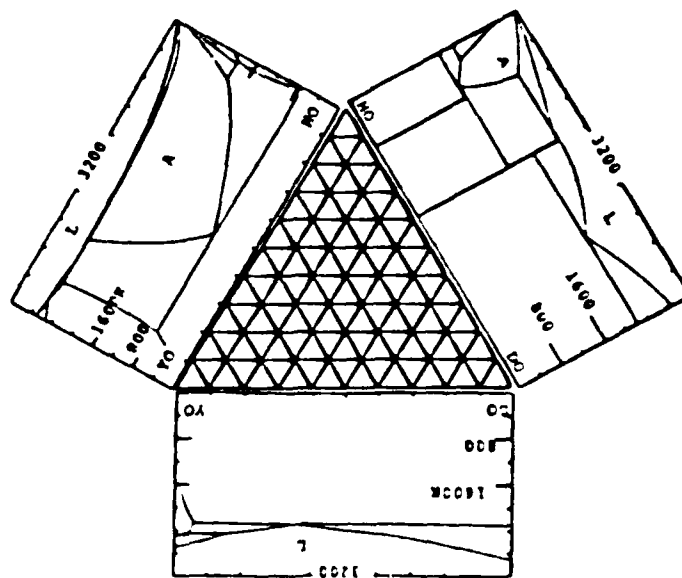


Figure 36. Calculated Isothermal Sections in the $\text{MgO}(1/3\text{MgO}_2)-\text{DC}(1/2\text{CaO})-\text{VC}(1/5\text{Fe}_2\text{O}_3)$ System.

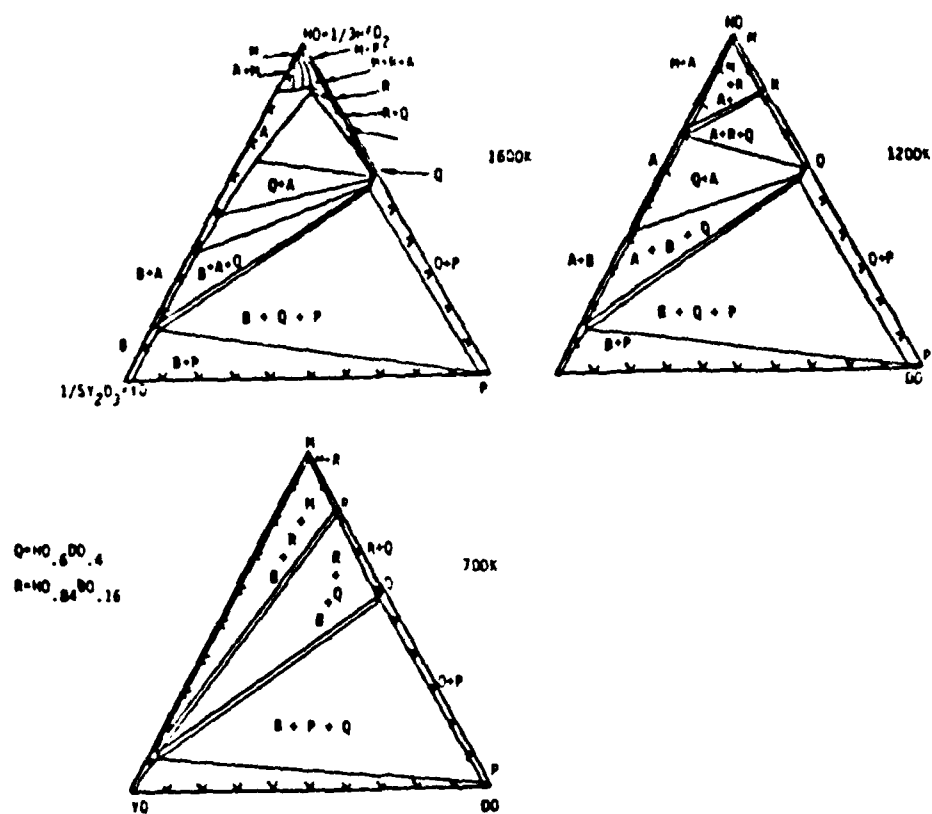


Figure 39. Calculated Isothermal Sections in HQ-DQ-YO

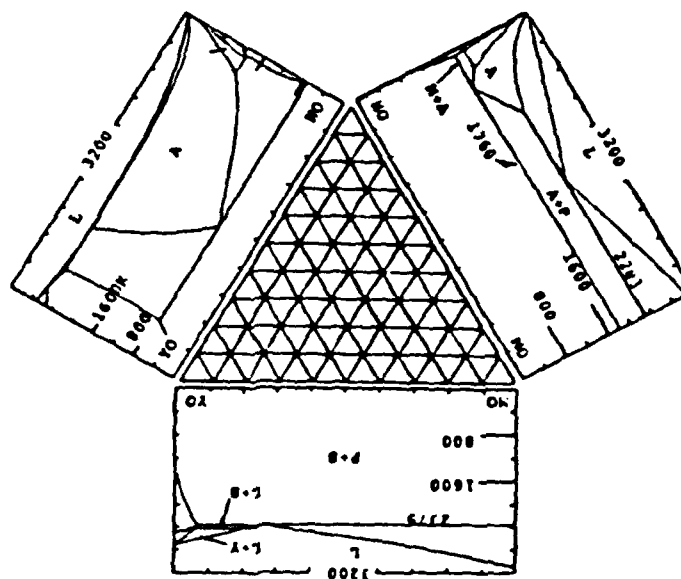


Figure 40. Calculated Isothermal Sections in the HQ(1/3HfO₂)-HQ(1/2MgO)-YO(1/3Y₂O₃) system

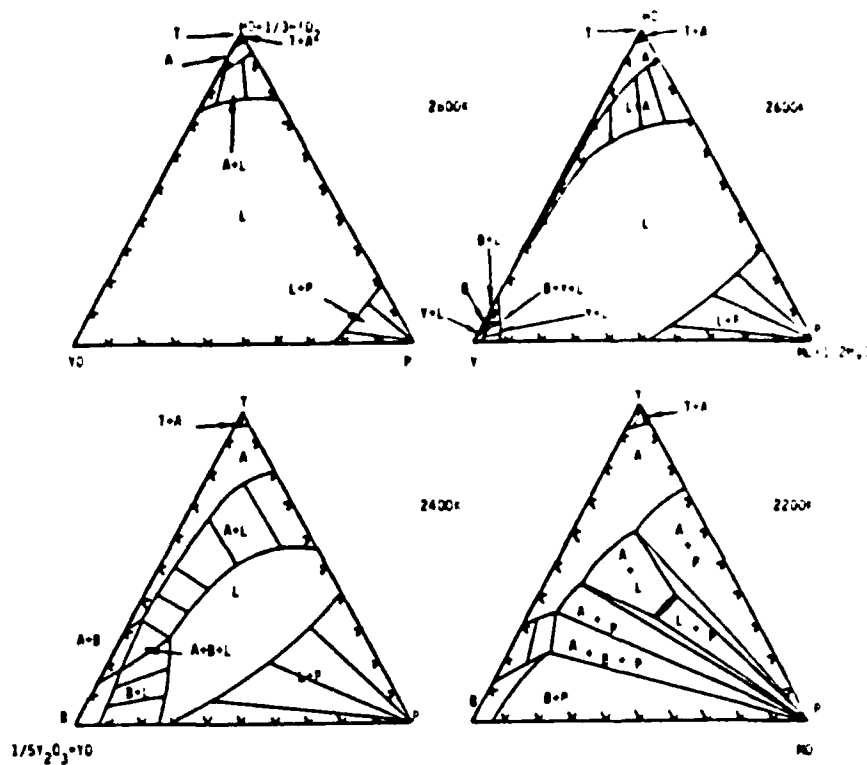


Figure 41. Calculated Isothermal Sections in MO-MO-YO

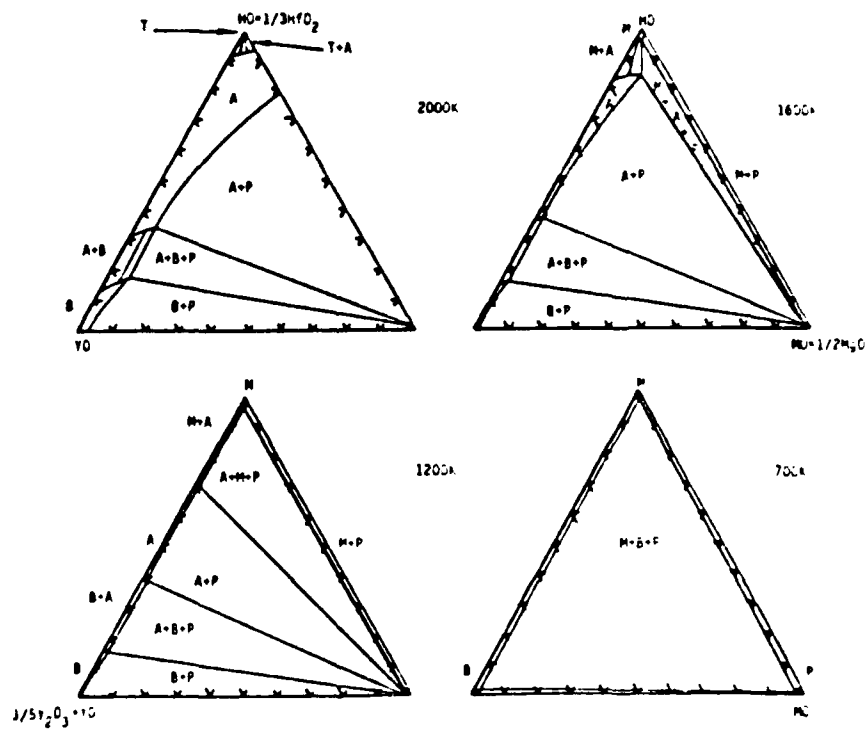


Figure 42. Calculated Isothermal Sections in MO-ML-YO

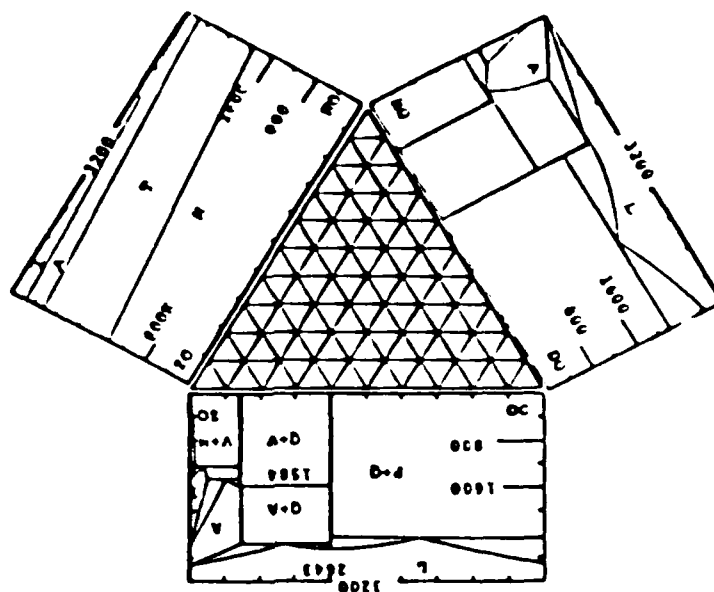


Figure 43. Calculated Isothermal Sections in the $\text{MO}(1/3\text{MnO}_2) - \text{DC}(1/2\text{CaO}) - \text{EC}(1/32\text{rO}_2)$ System

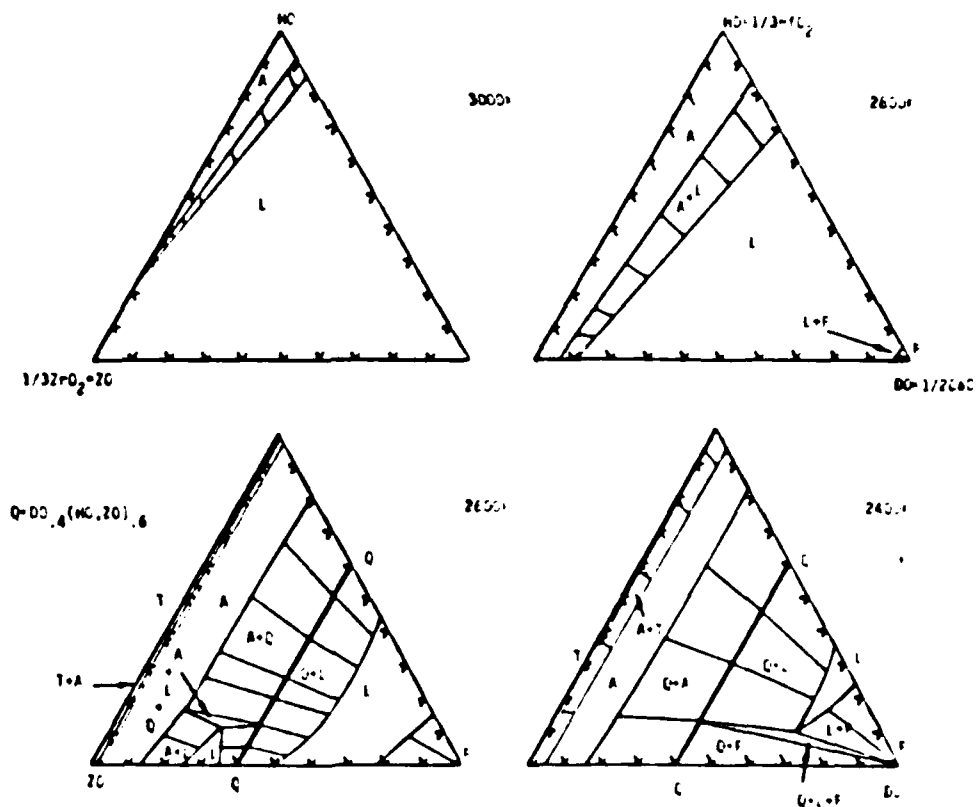


Figure 44. Calculated Isothermal Sections in MO-DC-ZC

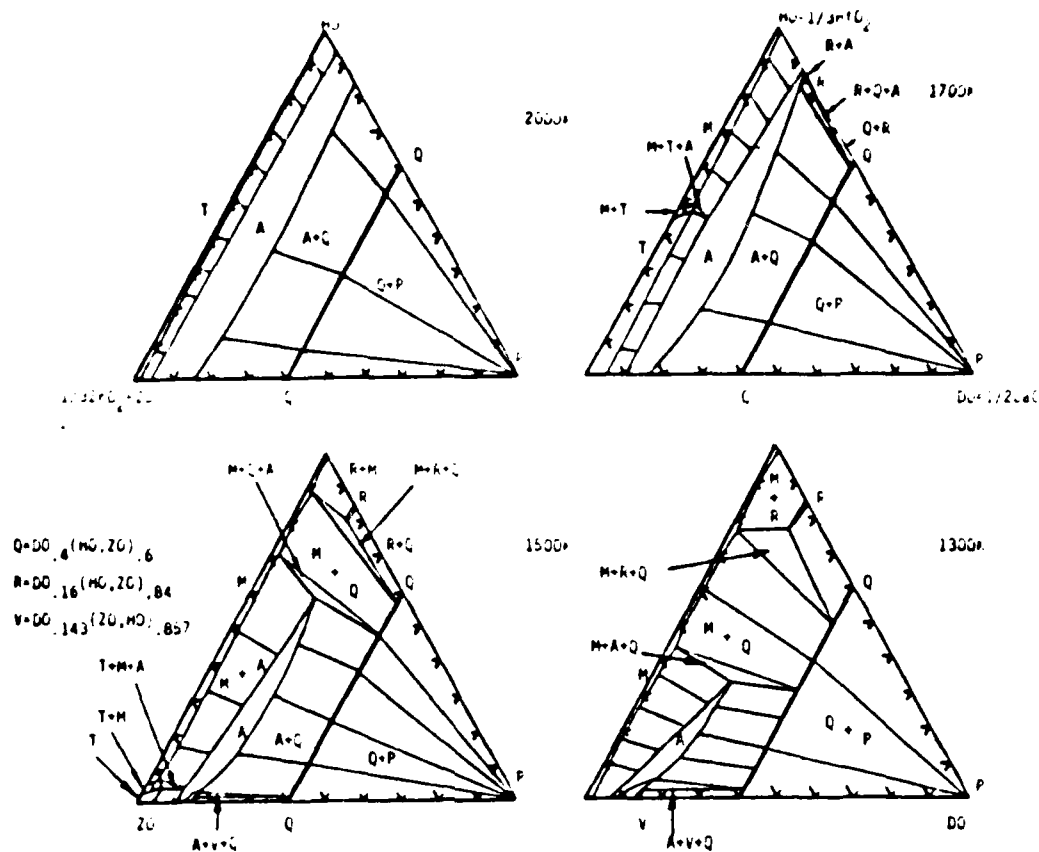


Figure 45. Calculated Isothermal Sections in $\text{H}_2\text{O}-\text{D}_2\text{O}-\text{ZnO}$

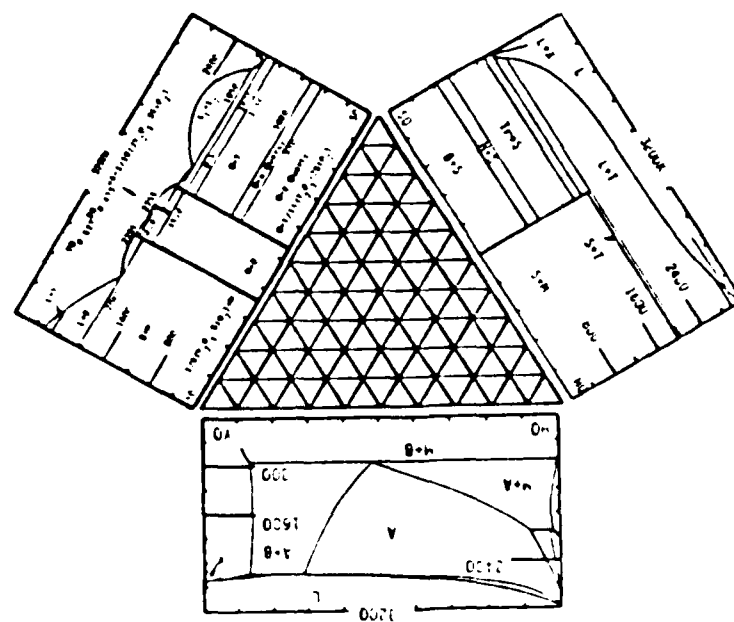


Figure 46. Calculated Isothermal Sections in the $50(1/3510_4)-\text{H}_2\text{O}-1/3\text{H}_2\text{O}_2-1/10(1/51_2\text{O}_3)$ System.

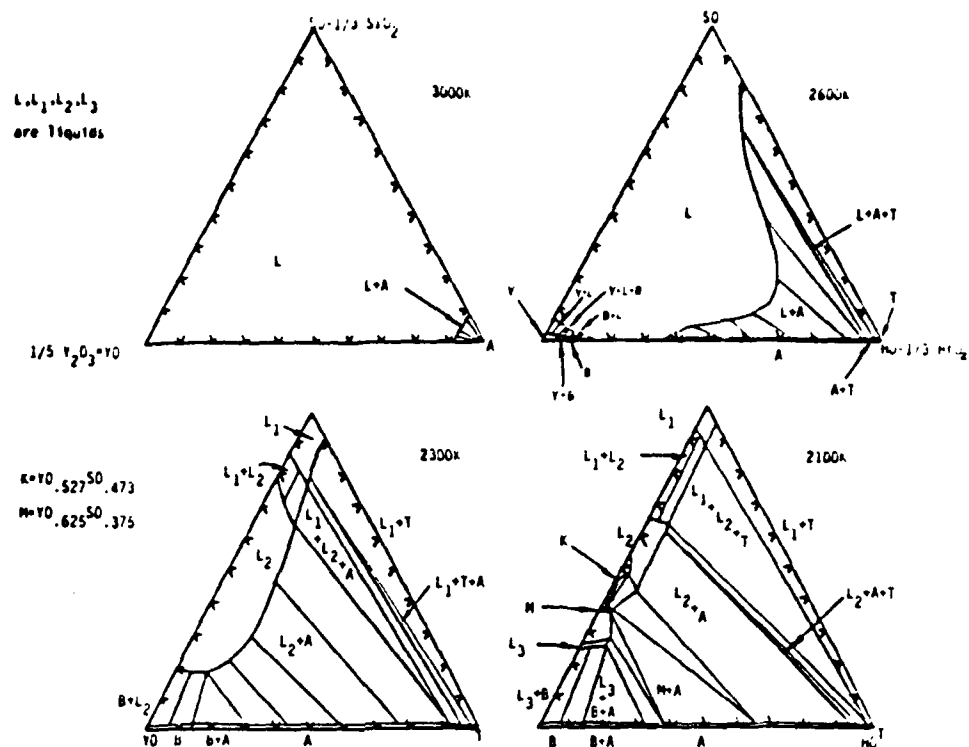


Figure 47. Calculated Isothermal Sections in $SO-HO-YO$

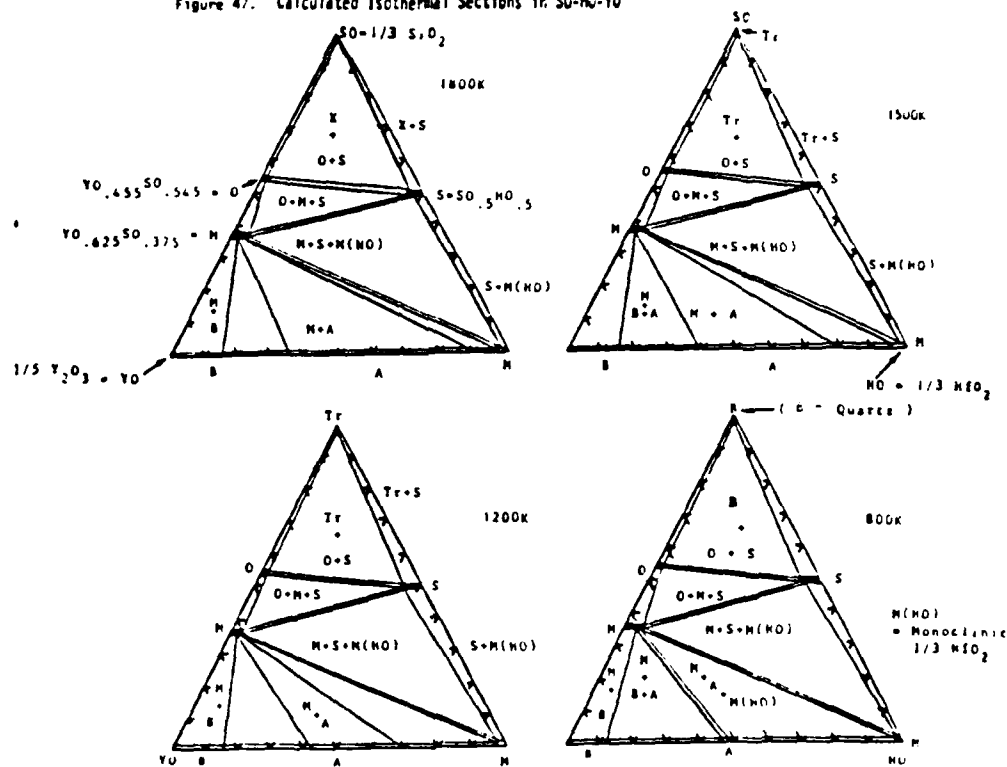
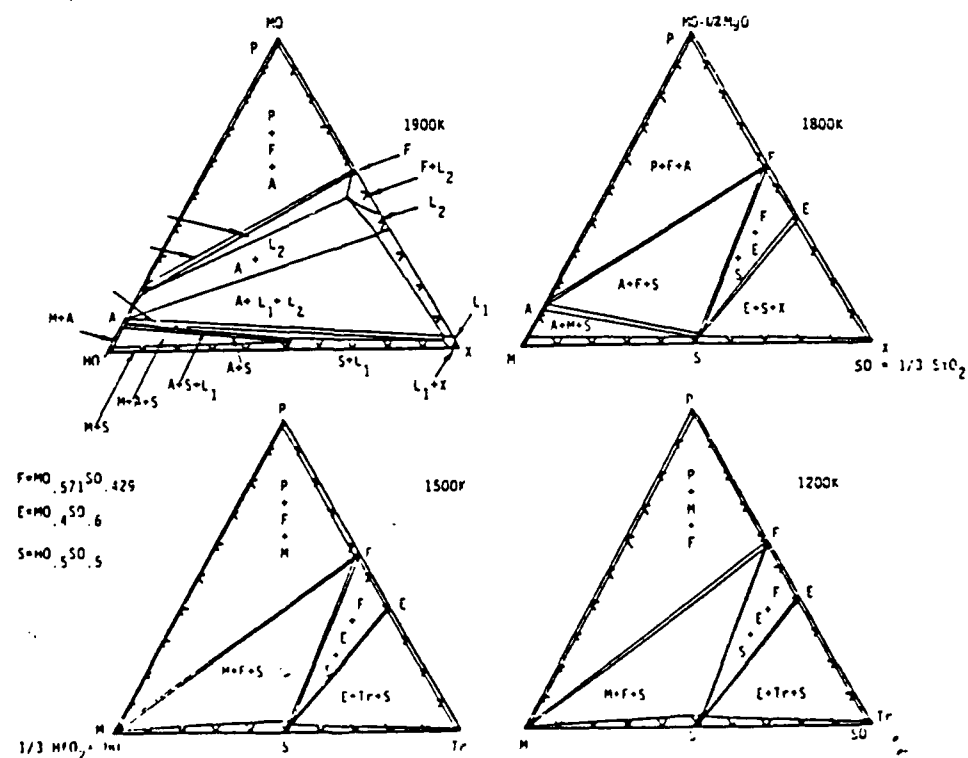
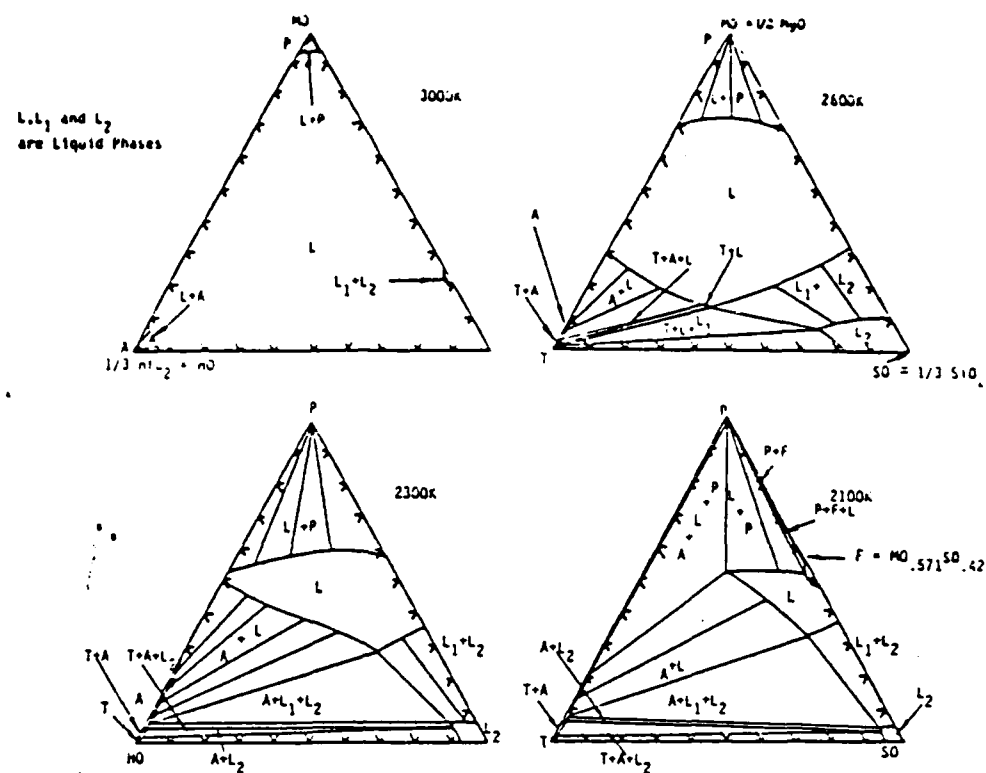


Figure 48. Calculated Isothermal Sections in $SO-HO-YO$



arranged in the appropriate ternary orientation to permit interpretation of the ternary sections. Thus the analysis of the MO-TO-GO case begins with Figure 20 showing the component quasi-binary systems. The calculated isothermal sections for this system between 2800°K and 700°K are shown in Figures 21-23. Figure 720 (8) shows a set of compatibility relations for this system. Examination of Figures 21-23 show how the phase fields covered by the liquid shrinks as the temperature is reduced with the final bit of liquid disappearing between 1600°K and 1590°K in Figure 22. The isothermal sections calculated at 1500°K and below, in Figure 23, agree with the experimental compatibility relations noted above in Figure 720 (8). The extension of the binary description into the ternary is performed along the lines suggested by Equations (7) - (11) in reference (4). Tables 3 and 5 contain all of the required descriptive data. Reference to Table 5 shows that in most cases the counterphase parameter is equal to zero. Two obvious exceptions are the $SO_{.5}(HO,ZO)_{.5}$ and $DO_{.4}(HO,ZO)_{.6}$ cases where the counterphase is stable and can be located in Table 3. In the other five cases where the counterphase is characterized by finite values of the counterphase parameter i.e. in GO-MO-SO and GO-MO-DO, the specific values were chosen to conform to available experimental data. In all cases, the ternary compound parameter, CAB, was set equal to zero except for the MO-TO-GO discussed above. In this case the values chosen were selected to conform to the published

compatibility results given in Figure 720 (8). Figures 24-26 display the calculated isothermal sections for the GO-MO-SO system. The sections at 1200°K and 900°K are in agreement with the compatibility relations shown in Figure 717 (8). The high temperature equilibria is dominated by the liquid miscibility gap entering from the GO-MO edge and the Enstatite ($E = MO_{.4}(GO,SO)_{.6}$) and Fosterite ($F = MO_{.571}(GO,SO)_{.429}$) compounds. The calculated isothermal sections in the GO-MO-DO system shown in Figures 27-29 were derived on the basis of the compatibility diagram in Figure 2470 (8). These are similar to the GO-MO-SO case in that the high temperature equilibria is dominated by the E and F phases. In each of the forgoing systems very little solid solubility is suggested. The SO-ZO-HO ($SiO_2-ZrO_2-HfO_2$) system shown in Figures 30-32 is of specific interest in high temperature thermal protection systems due to low diffusivity for oxygen in silica. However the system is dominated by liquid miscibility gaps and the equimolar compound, Q, which does not appear to possess very high stability (see Tables 3-5). The HO-DO-MO case in Figures 33-35 shows some solid solubility in the cubic A phase at high temperatures in the HO corner of the phase diagram. The calculations suggest that the liquid phase will disappear just below 2200°K. The analog HO-DO-YO and HO-MO-YO systems shown in Figures 36-38 and 39-41 both show extensive ranges of solid solubility in the cubic A and B phases which enter the ternary from the HO-YO edge. In both cases liquid phase is still present at

2200°K. However the liquid is just barely stable in the HO-MO-YO case at 2200°K. In both these systems it is likely that quenching from the high temperature cubic A region could permit retention of this phase at low temperatures. Figures 43-45 show the calculated isothermal sections in the HO-DO-ZO ($1/3\text{HfO}_2$ - $1/2\text{CaO}$ - $1/3\text{ZrO}_2$) system. Experimental sections are displayed in Figures 5392 A and B (9). Although the general features of the calculations and observations are in agreement a number of subtle differences appear which stem from differences in the binary and unary systems. To start with, the experimental diagram in Figure 5392 (9) appears to have combined the M/T transition in HfO_2 . Thus in comparison with Figure 15 or Figure 4444 (8) where an M/T transition is shown at 1976°K and a T/A transition shown at 2932°K, Figure 5392 B (9) shows a M(T)/A transition in HfO_2 near 2273°K. A second difference is that while Figure 15 and Figure 4444 (8) show narrow two-phase fields in HO-ZO Figure 5392 shows a wide two-phase field. Nevertheless the remaining features of the ternary sections are in general agreement. At high temperatures extensive solid solubility is present in cubic A phase leading to opportunities for retention on quenching and the equilibria is dominated by the Q phase $\text{DO}_{.4}(\text{HO},\text{ZO})_{.6}$ with no liquid present below 2300°K. Figures 46-48 show the calculated sections in SO-HO-YO. Vertical sections across the join HO-O and HO-M are shown in Figures 5440 and 5441 (9). Although there is general agreement between the calculated and

experimental results with respect to disappearance of the liquid phase near 2100°K in the HO-M join and 1800°K on the HO-O join, the calculations do not show the wide range of stability in the ternary for the cubic A phase indicated by the experimental diagram in Figures 5440 and 5441 (9). It would appear that the experimental results must reflect "quenching-in" of the high temperature form or difficulty in recognizing the O, M or Q phases experimentally. In view of the forgoing discussion of the heat of formation of the A phase in the HO-YO system presented earlier it is very unlikely that this phase could extend into the ternary to the extent suggested by Figures 5440 and 5441 (9) without precipitation of $\text{Y}_2\text{Si}_2\text{O}_7$, Y_2SiO_5 or HfSiO_2 . Figures 49-51 display the calculated MO-SO-HO sections which is dominated by miscibility gaps in the liquid phase at high temperatures and the E, F and S compound phases at low temperatures. Little solubility is suggested in the solid phases.

SUMMARY

The forgoing set of binary and ternary examples show how model calculations of ceramic phase diagrams can be performed in order to guide development of new structural systems and to evaluate existing experimental data. It also provides a means for planning future experimental studies.

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IV. CALCULATION OF MULTICOMPONENT CERAMIC PHASE DIAGRAMS

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A data base is being developed for calculation of quasi-binary and quasi-ternary phase diagrams of ceramic systems. Previous segments of this base cover combinations of Cr_2O_3 , MgO , Al_2O_3 , SiO_2 , CaO , Si_3N_4 , AlN , BeO , Y_2O_3 and Ce_2O_3 . Lattice Stability, Solution and Compound Phase Parameters have been derived covering the liquid, spinel, corundum, periclase, cristobalite, tridymite, quartz, hexagonal and beta prime phases which appear in the binary systems composed of pairs of these compounds. Compound phases formed from specific binary combinations of these compounds (i.e. $\text{MgO} \cdot \text{Cr}_2\text{O}_3$) have been characterized. This description is based on observed thermochemistry and phase diagrams for the binary systems of interest. Selected ternary systems have been computed based on the foregoing data base for comparison with experimental sections in order to illustrate the usefulness of the data base. To date, sixty six quasi binary and nineteen quasi ternary systems have been calculated. The most recent set of papers extend the base to cover GeO_2 , HfO_2 , ZrO_2 and TiO_2 . The components are of particular

interest in recent developments of structural high temperature ceramics and applications requiring unusual toughness. The current work deals with four quasi binary systems and four quasi ternary systems which have been calculated over a wide range of temperatures. These samples demonstrate the capability of the data base and computational model for dealing with phase equilibria in multicomponent oxide systems over a wide range of conditions and compositions of practical interest.

INTRODUCTION

Previous papers in the current series (1-6) provide descriptive information for computing condensed phase equilibria in ceramic systems. In view of current interest in applying ceramic systems in applications requiring toughness (7) and structural performance at high temperatures the present data base is being extended to cover GeO_2 , HfO_2 , ZrO_2 and TiO_2 . This has been effected by employing available sources (8-10) of thermochemical and phase diagram data. High temperature ceramics have received increased attention during the last few years for structural, thermal protection and engine applications. SIALONS and combinations of zirconia and hafnia with Al_2O_3 , SiC and Si_3N_4 have been shown to develop strength and toughness. This has opened the door to a whole range of new uses for these materials. Recently Slichting and co-workers (11,12) have shown that by alloying GeO_2 with SiO_2 a whole range of glasses can be synthesized with tailor-made coefficients of expansion. Utilization of such compositions offers the possibility of enhancing the high temperature oxidation resistance of ceramic composites in which a mixed GeO_2 - SiO_2 phase with a desired CTE could replace the conventional SiO_2 as a filler. This kind of compositing would open an entire spectrum of new opportunities for synthesis of high

temperature ceramics. One of the major obstacles in the development of complex composite systems is the lack of phase diagram information which can be used to guide the fabrication and processing of a new material and help to predict its performance. The current methods of employing models to predict high temperature behaviour has proven useful when basic data is unavailable or too costly and time consuming to obtain by conventional means. This method consists of developing a data base of thermochemical and phase diagram information in analytical form and employing computer models to extend the description to binary and ternary systems. Recently J. Lorenz et al. (13) applied this method successfully to SiC-ZrO_2 and $\text{SiC-ZrO}_2\text{-Al}_2\text{O}_3\text{-SiO}_2$ in order to evaluate composition effects and identify fabrication conditions. In the present paper, the data base has been expanded by analyzing the following quasi-binary systems: $\text{TiO}_2\text{-Al}_2\text{O}_3$, $\text{TiO}_2\text{-SiO}_2$, $\text{TiO}_2\text{-CaO}$, and $\text{TiO}_2\text{-Y}_2\text{O}_3$. These results when combined with earlier findings (1-6) were employed to compute a range of isothermal sections in the following quasi ternary systems sufficient to define their characteristics: $\text{TiO}_2\text{-Al}_2\text{O}_3\text{-MgO}$, $\text{Al}_2\text{O}_3\text{-TiO}_2\text{-SiO}_2$, $\text{TiO}_2\text{-Al}_2\text{O}_3\text{-HfO}_2$ and $\text{MgO-SiO}_2\text{-TiO}_2$.

LATTICE STABILITY VALUES

Table 1 defines the lattice stability values employed in the current study. Data for the stable forms were taken from Kubaschewski and Alcock (8). The remaining values were adopted along with the lines employed previously. As shown in Table 1 the present analysis is based on one gram atom of compound and/or solution phase. Moreover, dissociation of the components or vaporization is not considered!

TABLE 1

SUMMARY OF LATTICE STABILITY PARAMETERS

(All units in Joules per gram atom (mole of atoms), T in Kelvins).

P = Periclase, C = Corundum, S = Spinel, X = Cristobalite

Tr = Tridymite, H = Hexagonal(α quartz), B(SO) = Trigonal(β quartz)

F = Rutile (TO)

A = Cubic (HO and ZO)

T = Tetragonal (HO and ZO)

M = Monoclinic (HO and ZO)

Y = High Temperature YO

B = Low Temperature YO

L = Liquid

GO = $1/3\text{GeO}_2$ HO = $1/3\text{HfO}_2$ TO = $1/3\text{TiO}_2$ AO = $1/5\text{Al}_2\text{O}_3$ MO = $1/2\text{MgO}$ DO = $1/2\text{CaO}$ SO = $1/3\text{SiO}_2$ YO = $1/5\text{Y}_2\text{O}_3$ CE = $1/3\text{CeO}_2$ ZO = $1/3\text{ZrO}_2$ GOGOLH* = $(1/3)\text{GeO}_2$ (liquid) - $(1/3)\text{GeO}_2$ (hexagonal)GOGOLR = $(1/3)\text{GeO}_2$ (liquid) - $(1/3)\text{GeO}_2$ (rutile)GOGOHR = $(1/3)\text{GeO}_2$ (hexagonal) - $(1/3)\text{GeO}_2$ (rutile)

GOGOLR = GOGOLH + GOGOHR

GOGOLH = 14644 - 10.54T

HOHOLA = 34865 - 11.00 T

GOGOLR = 22087 - 16.23T

HOHOAT = 2239 - 0.753T

GOGOHR = 7443 - 5.69T

HOHOLT = 37104 - 11.753T

GOGOLA = - 11.00T

HOHOTM = 2724 - 1.381T

GOGOLT = - 11.57T

HOHOLM = 39828 - 13.134T

GOGOLM = - 13.14T

HOHOLX = - 2.092T

GOGOLC = - 10.21T

HOHOLP = - 8.368T

GOGOLP = - 8.37T

HOHOLC = - 10.209T

GGOBR = 7109 - 5.40T

HOHOLB = 30711 - 10.376T

GOGOLX = - 1.67T

HOHOLY = 23849 - 9.396T

GOGOLTr = - 2.01T

HOHOLR = 21882 - 10.42 T

GGOHB(SO) = 335 - 0.29T

GGOTrH = 14644 - 8.54T

* These differences specify the free energy of one phase (i.e. liquid) minus the free energy of the second phase (i.e. hexagonal) for a compound.

TABLE 1 (CONCLUDED)

SUMMARY OF LATTICE STABILITY PARAMETERS

(All units in Joules per gram atom (mole of atoms), T in Kelvins).

P = Periclase, C = Corundum, S = Spinel, X = Cristobalite

Tr = Tridymite, H = Hexagonal(α quartz), B(SO) = Trigonal(β quartz)

R = Rutile (TO)

A = Cubic (HO and ZO)

T = Tetragonal (HO and ZO)

M = Monoclinic (HO and ZO)

Y = High Temperature YO

E = Low Temperature YO

L = Liquid

DODOLR = - 16.23 T

MOMOLH = - 8.37 T

DODOLA = 8732 - 13.723T

MOMOLR = - 10.42 T

DODOFA = -31016

MOMOPA = -31016

DODOAM = 16736 - 2.134T

MOMOLA = 16682 - 15.397T

DODOAT = -22594

MOMOAM = 16736 - 2.134T

DODOLB = 5439 - 8.368T

MOMOAT = -22594

DODOLY = 2929 - 6.694T

MOMOLB = 13389 - 10.042T

DODOLR = - 10.418T

MCMOLY = 10878 - 8.368T

TOTOLH = 14539 - 10.54 T

YOYOLY = 22694 - 8.368T

TOTOLR = 22313 - 10.42 T

YOYOLB = 26878 - 10.042T

TOTOHR = 7774 - 0.12 T

YOYOLA = 22615 - 9.832T

TOTOLP = - 8.37 T

YOYOAM = -16736 - 2.134T

TOTOLT = 3527 - 11.757T

YOYOTM = -8368 - 1.381T

TOTOLA = 1017 - 11.757T

YOYOLP = - 12.552T

TOTOTM = -12259

YOYOLR = - 10.418T

TOTOLC = - 10.209T

TOTOLX = - 1.674T

ZOZOLA = 29008 - 9.832T

TOTOLY = - 8.368T

ZOZOAT = 1987 - 0.753T

TOTOLF = 6862 - 9.205T

ZOZOTM = 2008 - 1.381T

ZOZOLF = - 8.368T

AOAOLH = - 10.54 T

AOAOLA = - 9.832T

SOSOHF = -4790 - 5.69 T

AOAOLT = - 10.585T

SOSOLT = - 2.092T

AOAOLY = - 8.368T

SOSOLA = - 2.092T

AOAOLF = - 10.418T

SOSOLF = - 2.092T

CECELC = - 10.209T

BINARY SYSTEMS

Tables 2-4 and Figures 1-4 summarize the results for the binary systems listed above. The solution phases are described as subregular solutions along the lines of Equations (1) and (2) of reference 4. When the subregular parameters are equal (i.e. TO-YO, R, Y and B) the solution is regular. The compound phases are defined at fixed compositions (i.e. Equations 3-6 of reference 4) in terms of the compound parameter and the base phase. The latter values are listed in Table 3 for the compounds of interest while Table 4 shows the Gibbs energy of formation for these compounds from the component oxides. Figure 1 shows the calculated TO-AO ($1/3 \text{ TiO}_2$ - $1/5 \text{ Al}_2\text{O}_3$) system derived from the description contained in Tables 1-4 in conformity with Figure 316 of reference 9. There is one quasi binary compound phase, designated as D, which transforms on heating above 2000K to a different structure, E. Figure 2 shows the calculated TO-SO ($1/3 \text{ TiO}_2$ - $1/3 \text{ SiO}_2$) system based on Tables 1-4 which agrees with the experimental diagram in Figure 113 (9). Figure 3 displays the calculated TO-DO phase diagram ($1/3 \text{ TiO}_2$ - $1/2 \text{ CaO}$) based on Tables 1-4 in agreement with the experimental results shown in Figures 239, 4312 and 4553 (9). The final quasi binary in this group shown in Figure 4 is the calculated TO-YO ($1/3 \text{ TiO}_2$ - $1/5 \text{ Y}_2\text{O}_3$) system based on the parameters listed in Tables 1-4. These parameters were derived from the values given earlier for the HO-YO system (6).

TERNARY SYSTEMS

The description of the forgoing systems combined with those presented earlier (1-6) have been employed to calculate isothermal sections in four quasi-ternary systems over a wide range of temperatures. The results of these calculations are presented in

TABLE 2

QUASIBINARY SOLUTION PARAMETERS FOR OXIDE SYSTEMS

(All units in Joules per gram atom (mole of atoms), T in Kelvins)

LTOAO =	5335 + 13.60T	LAOTO =	7008 + 13.60T
RTOAO =	62760	RAOTO =	62760
CTOAO =	62760	CAOTO =	62760
LTOSO =	51463	LSOTO =	43095
FTOSO =	62760	RSOTO =	62760
XTOSO =	62760	XSOTO =	62760
TTOSO =	62760	TSOTO =	62760
LTODD =	-127194 + 50.208T	LDOTO =	-127194 + 50.208T
RTODD =	41840	RDOTO =	41840
PTODD =	83680	PDOTO =	83680
LTOYO =	14016 + 4.184T	LYOTO =	14016 + 4.184T
RTOYO =	20920	RYOTO =	20920
YTOYO =	20920	YYOTO =	20920
BTOYO =	20920	BYOTO =	20920

TABLE 3

SUMMARY OF COMPOUND PARAMETERS FOR BINARY SYSTEMS

(All units in Joules per gram atom (mole of atoms) T in Kelvins)

<u>Compound</u>	<u>Name</u>	<u>Stoichiometry</u>	<u>Stability</u>	<u>Base</u>	<u>Compound/Parameter</u> (Joules/g.at.)
(1/8)(Al ₂ O ₃ .TiO ₂)	D	TO _{.375} AO _{.625}	stable (below 2000K)	C	81755 + 4.15T
(1/8)(Al ₂ O ₃ .TiO ₂)	E	TO _{.375} AO _{.625}	stable (above 2000K)	C	-16004 + 44.00T
(1/5)(TiO ₂ .CaO)	U	TO _{.6} DO _{.4}	stable	F	8786 + 37.03T
(1/17)(3TiO ₂ .4CaO)	V	TO _{.529} DO _{.471}	stable	F	-5021 + 33.43T
(1/12)(TiO ₂ .3CaO)	V	TO _{.4} DO _{.6}	stable	F	-10042 + 33.83T

TABLE 4

CALCULATED FREE ENERGY OF FORMATION OF COMPOUND PHASES
(All units in Joules per gram atom (mole of atoms) T in Kelvins)

Compound	Name	Stoichiometry	Free Energy of Formation, ΔG_f (298K) from Component Compounds
(1/6)(Al ₂ O ₃ ·7TiO ₂)	D	TO _{.375} AO _{.625} stable below 2000K	-9272 + 4.0E1T
(1/6)(Al ₂ O ₃ ·7TiO ₂)	E	TO _{.375} AO _{.625} stable above 2000K	13606 - 7.373T
(1/5)(TiO ₂ ·CaO)	U	TO _{.6} DO _{.4}	-16175 - 3.95T
(1/5)(TiO ₂ ·CaO)	U	TO _{.6} DO _{.4}	ΔH_f (experimental) = -16150 ± 516
(1/17)(3TiO ₂ ·4CaO)	V	TO _{.529} DO _{.471}	-17451 + 3.03T
(1/17)(3TiO ₂ ·4CaO)	V	TO _{.529} DO _{.471}	ΔH_f (experimental) = -17472
(1/12)(2TiO ₂ ·3CaO)	W	TO _{.5} DO _{.5}	-17397 + 3.0ET
(1/12)(2TiO ₂ ·3CaO)	W	TO _{.5} DO _{.5}	ΔH_f (experimental) = -17431

TABLE 5

SUMMARY OF COUNTERPHASE STOICHIOMETRY AND PARAMETERS
EMPLOYED IN TERNARY CALCULATIONS

System	Stable Phase (Name)	Base	Counterphase	Base	Counterphase Parameter (Joules/g. At.)
TO-AO-MO	TO _{.75} MO _{.75} (W)	F	AO _{.75} MO _{.25}	F	0
	TO _{.6} MO _{.4} (V)	F	AO _{.6} GO _{.4}	F	0
	TO _{.429} MO _{.571} (U)	F	AO _{.429} MO _{.571}	F	0
	TO _{.375} AO _{.625} (D)	C	MO _{.375} AO _{.625}	C	0
	TO _{.375} AO _{.625} (D)	C	MO _{.375} AO _{.625}	C	0
	MO _{.266} AO _{.714} (SP)	S	MO _{.266} TO _{.714}	S	0
AO-TO-SO	TO _{.375} AO _{.625} (D)	C	SO _{.375} AO _{.625}	C	0
	TO _{.375} AO _{.625} (E)	C	SO _{.375} AO _{.625}	C	0
	SO _{.266} AO _{.714} (M)	C	TO _{.266} AO _{.714}	C	0
TO-AO-TI	MO _{.5} TO _{.5} (S)	T	AO _{.5} TO _{.5}	T	0
	MO _{.375} AO _{.625} (D)	C	TI _{.375} AO _{.625}	C	0
	MO _{.375} AO _{.625} (E)	C	TI _{.375} AO _{.625}	C	0
MO-TI-SO	TO _{.75} MO _{.25} (W)	F	SO _{.75} MO _{.25}	F	(ICAF = -25104)
	TO _{.6} MO _{.4} (V)	F	TO _{.6} SO _{.4}	F	(ICAF = -58576)
	TO _{.429} MO _{.571} (U)	F	TO _{.429} SO _{.571}	F	0
	MO _{.375} AO _{.625} (D)	C	TI _{.375} SO _{.625}	C	0
	MO _{.375} AO _{.625} (E)	C	TI _{.375} SO _{.625}	C	0

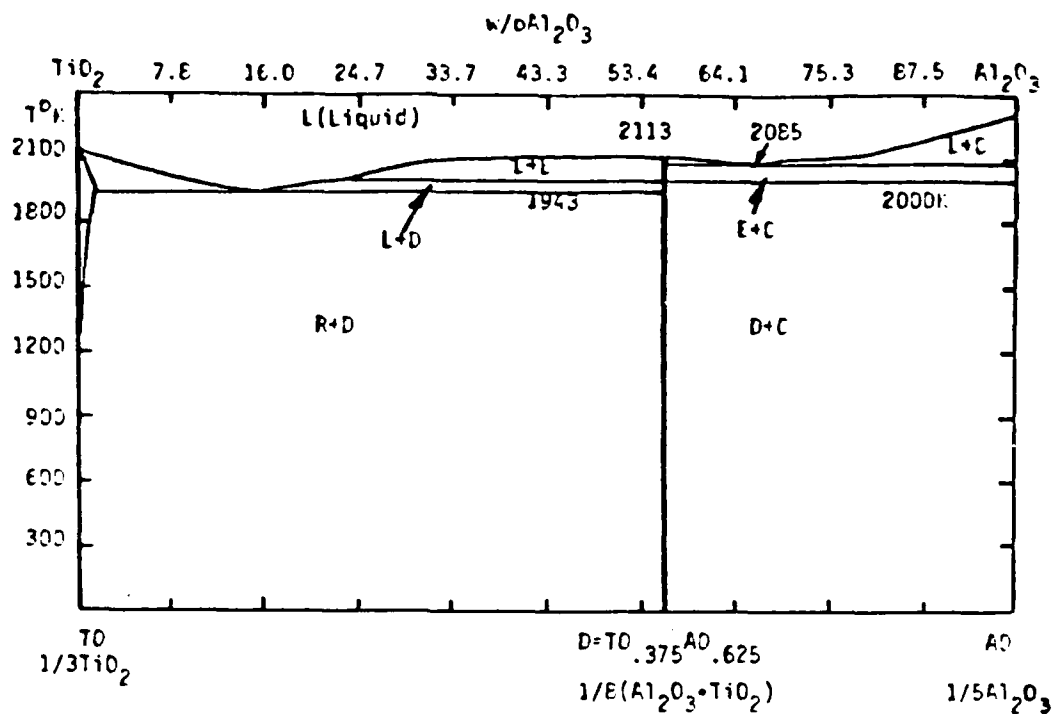


Figure 1. Calculated TiO_2 - Al_2O_3 Phase Diagram.

E=High Temperature form of D

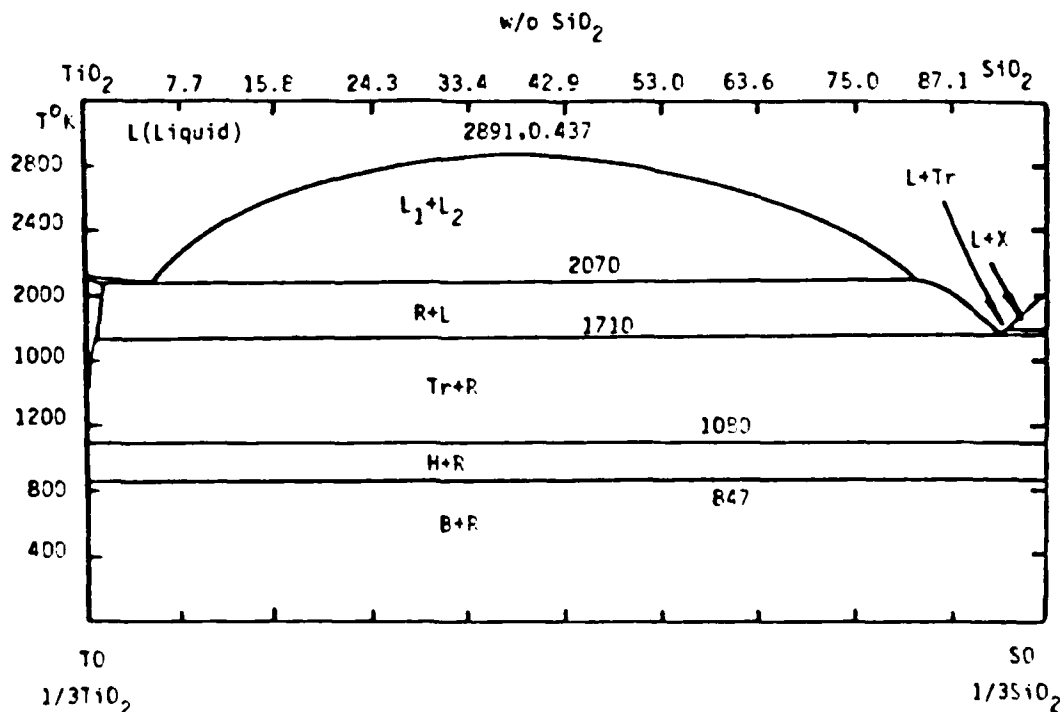


Figure 2. Calculated TiO_2 - SiO_2 Phase Diagram

X=Crystoballite
Tr=Tridymite
H= Quartz
B= Quartz

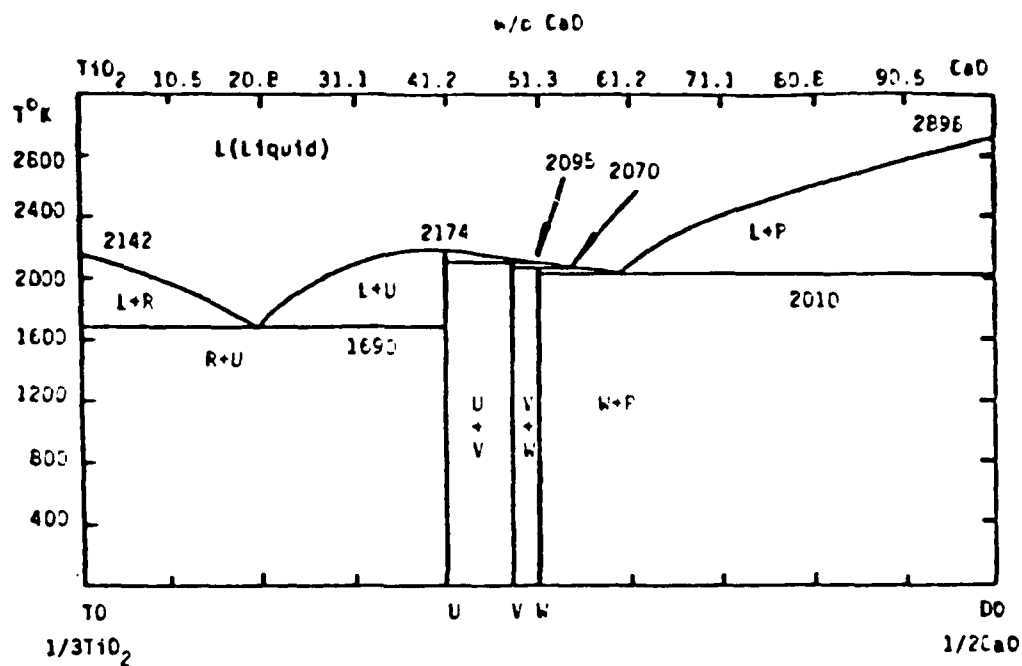


Figure 3. Calculated TiO_2 -CaO Phase Diagram.

$$U = \text{TO}_{.6004} = 1/5(\text{TiO}_2 \cdot \text{CaO})$$

$$V = \text{TO}_{.5290} \text{DO}_{.471} = 1/17(3\text{TiO}_2 \cdot 4\text{CaO})$$

$$W = \text{TO}_{.5005} = 1/12(2\text{TiO}_2 \cdot 3\text{CaO})$$

$$\text{w/o } \text{Y}_2\text{O}_3$$

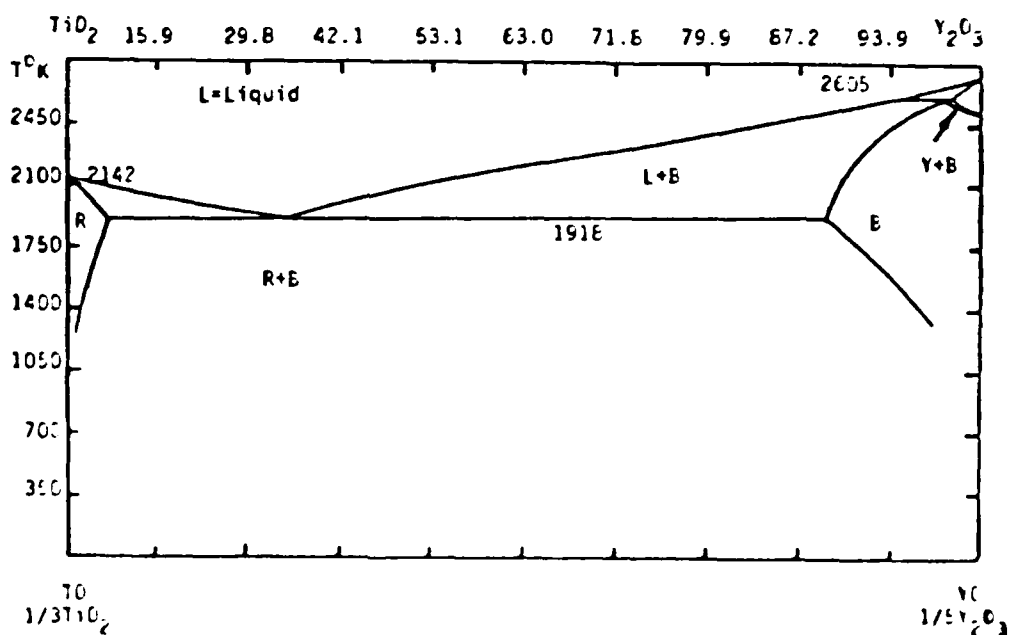


Figure 4. Calculated TiO_2 - Y_2O_3 Phase Diagram.

Figures 5-20. In each case the component quasi-binary systems are arranged in the appropriate ternary orientation to permit interpretation of the ternary sections. Thus the analysis of the TO-AO-MO case begins with Figure 5 showing the component quasi-binary systems. The calculated isothermal sections for this system between 3000°K and 900°K are shown in Figures 6-8. Figures 713 and 714 (8) show experimental results which are in good agreement with the calculations. In particular the minimum melting point is shown at 1843K in the experimental diagram in the TO rich part of the TO-MO edge in excellent agreement with the calculated results shown in Figure 7. In addition, Figure 6 shows the calculated Gibbs energy changes defining the subsolidus stability of the various compounds in the system. The extension of the binary description into the ternary is performed along the lines suggested by Equations (7) - (11) in reference (4). Tables 3 and 5 contain all of the required descriptive data. Reference to Table 5 shows that in most cases the counterphase parameter is equal to zero. This is true in the TO-AO-MO, AO-TO-SO AND AO-TO-HO cases. The values for F, V and W in MO-SO-TO were chosen to agree with experimental results. Figures 9-12 show the AO-TO-SO calculations which are dominated by the miscibility gap in the liquid at high temperatures and the compound interactions at low temperatures. The Gibbs energy change calculated for the interactions between the M+R and D+T pairs is shown in Figure 12 defining 1707K as the critical temperature. The calculated results are in good agreement with the experimental findings shown in Figures 771-775 (9). The calculated sections for TO-AO-HO are shown in Figures 13-16 which display the lowest temperature for liquid stability near the TO corner on the TO-AO edge just below 1900K. Experimental description of TO-AO-ZO in Figures 773-774 (9) shows similar behavior at 1853K. The final set of ternary calculations are presented in Figures 17-20 for the MO-SO-TO system. Extensive experimental data have been reported for this system in Figures

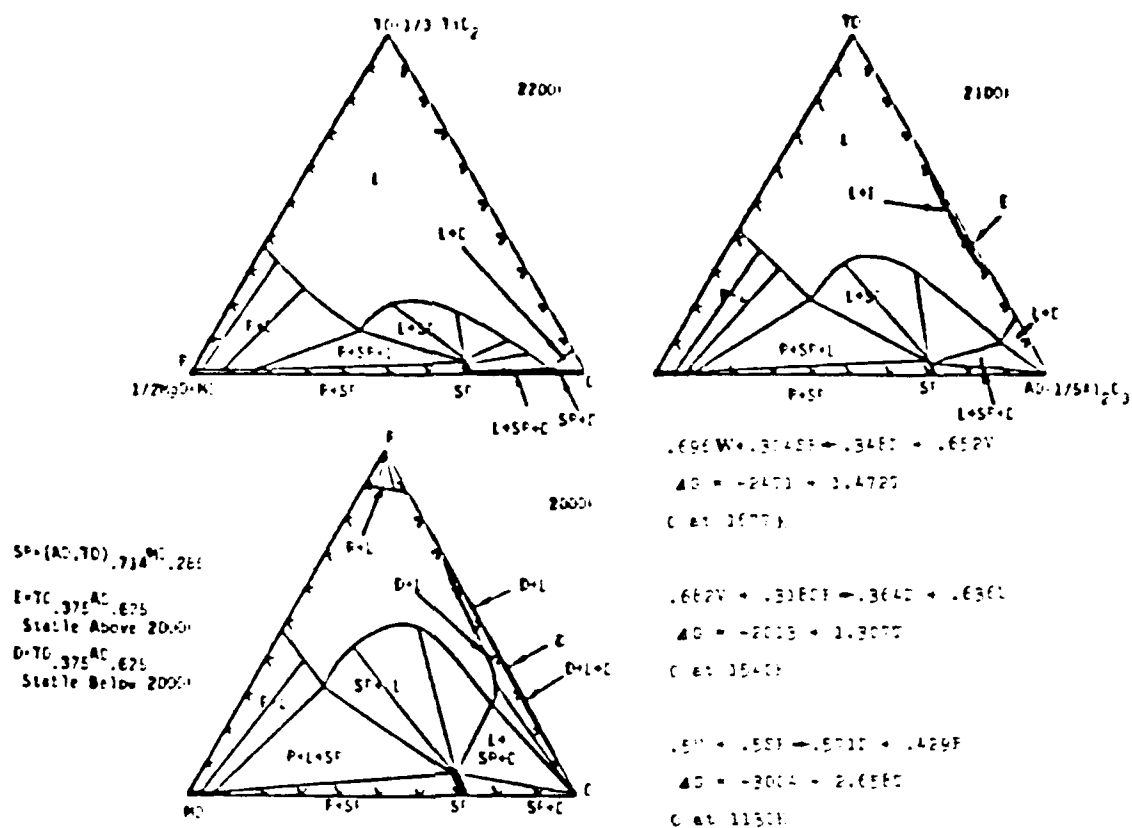


Figure 7. Calculated Isothermal Sections in TO-AC-MO

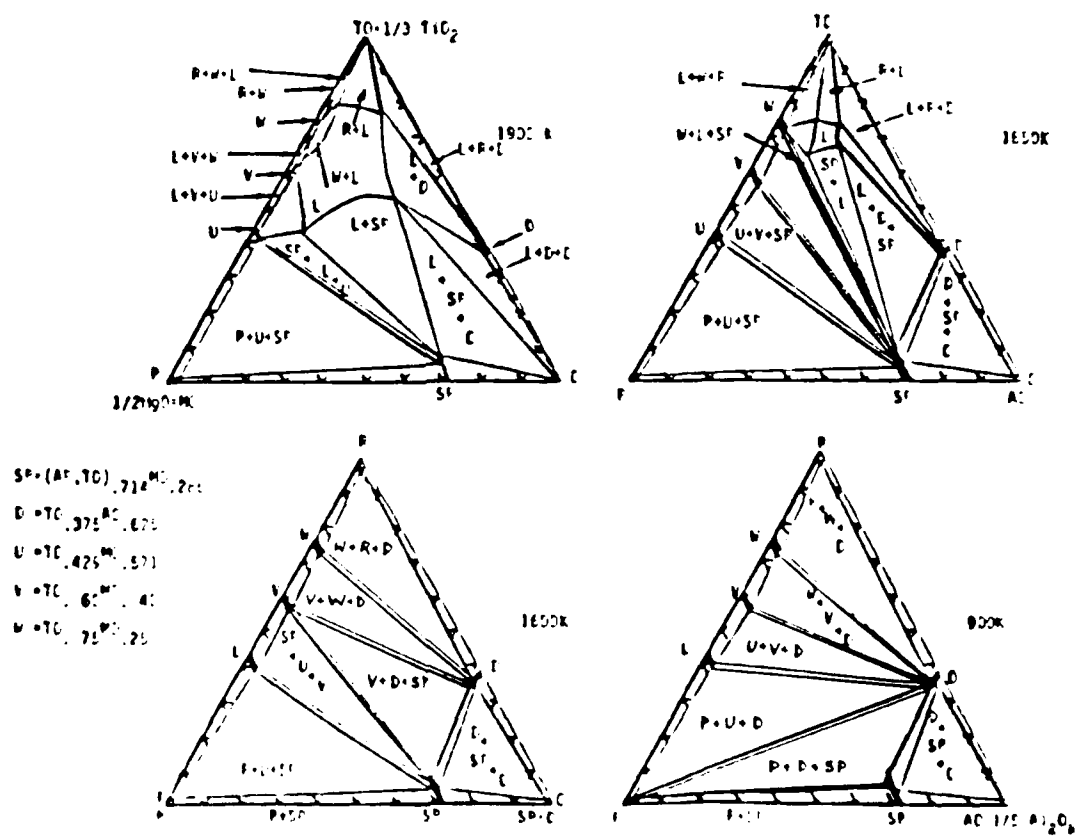


Figure 8. Calculated Isothermal Sections in TO-AC-M

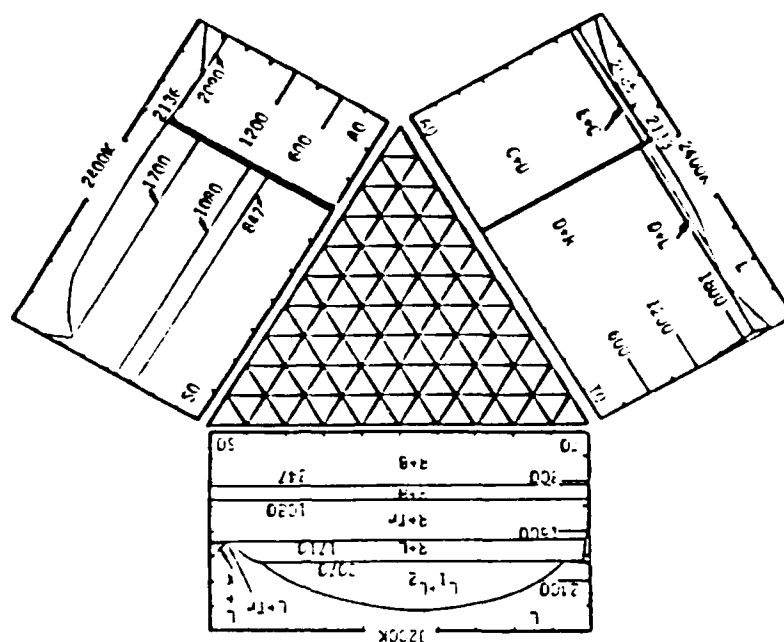


Figure 9. Calculated Isothermal Sections in the $AO(1/5 Al_2O_3) - TO(1/3 TiO_2) - SO(1/3 SiO_2)$ system.

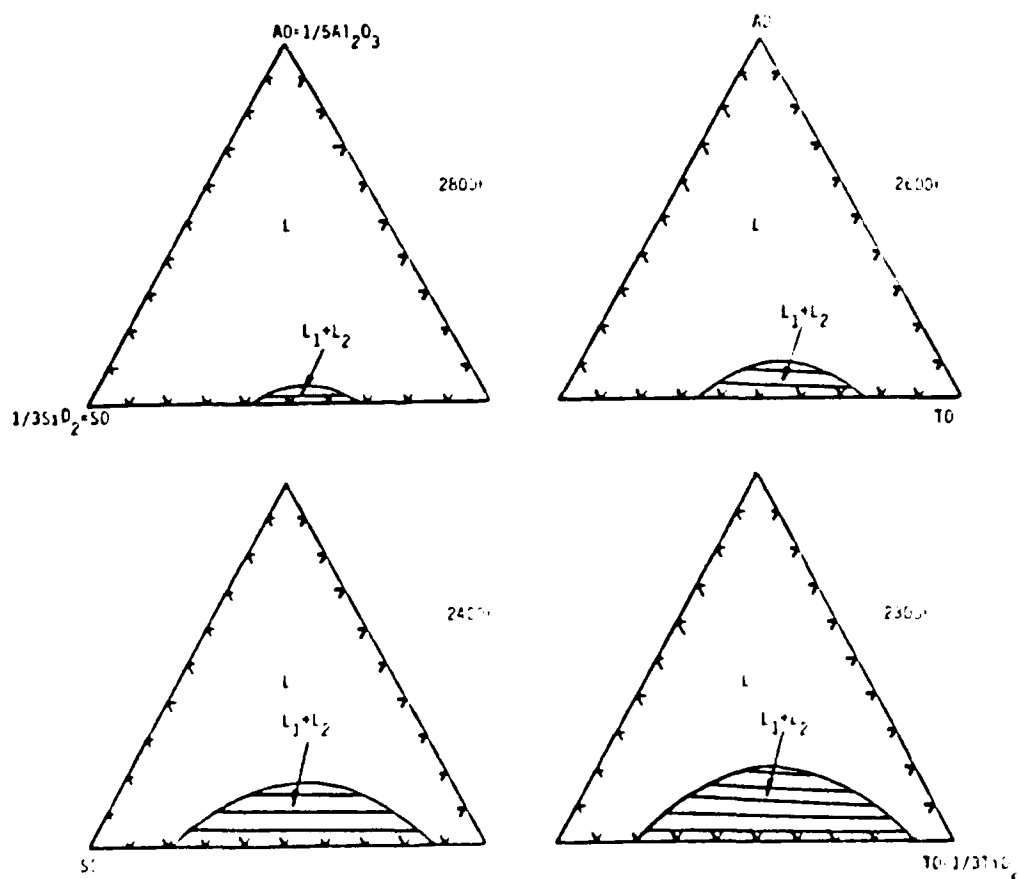


Figure 10. Calculated Isothermal Sections in AO-TO-SO

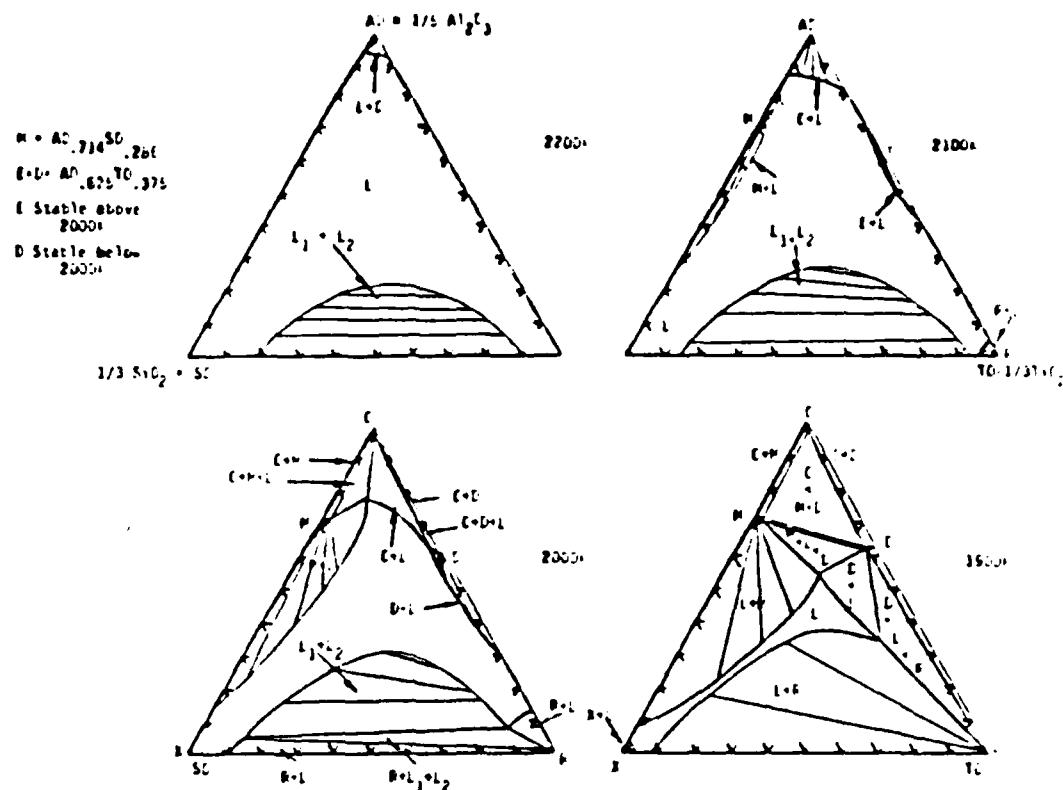


Figure 11. Calculated Isothermal Sections in AD-TO-SC

$.7M + .3E = .8C + .2T$
 $\Delta G = -6585 + 3.857T$
 0 at 1707°

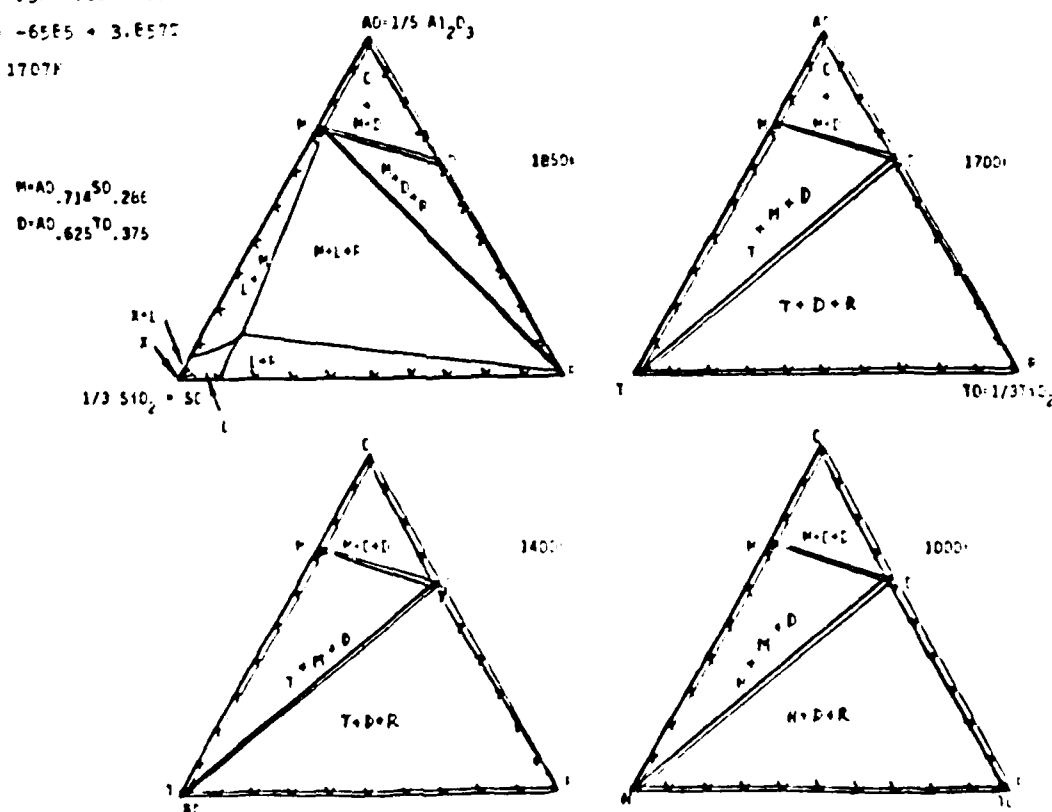


Figure 12. Calculated Isothermal Sections in AD-TO-SC

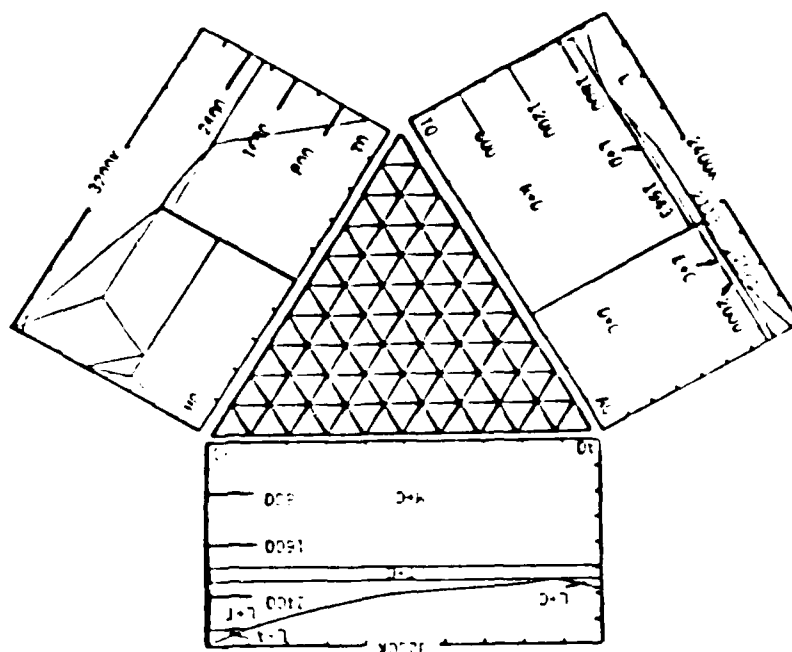


Figure 12. Calculated Isothermal Sections in the $\text{TiO}(1/3 \text{ TiO}_2) - \text{Al}_2\text{O}_3(1/3 \text{ Al}_2\text{O}_3) - \text{H}_2\text{O}(1/3 \text{ H}_2\text{O})$ system.

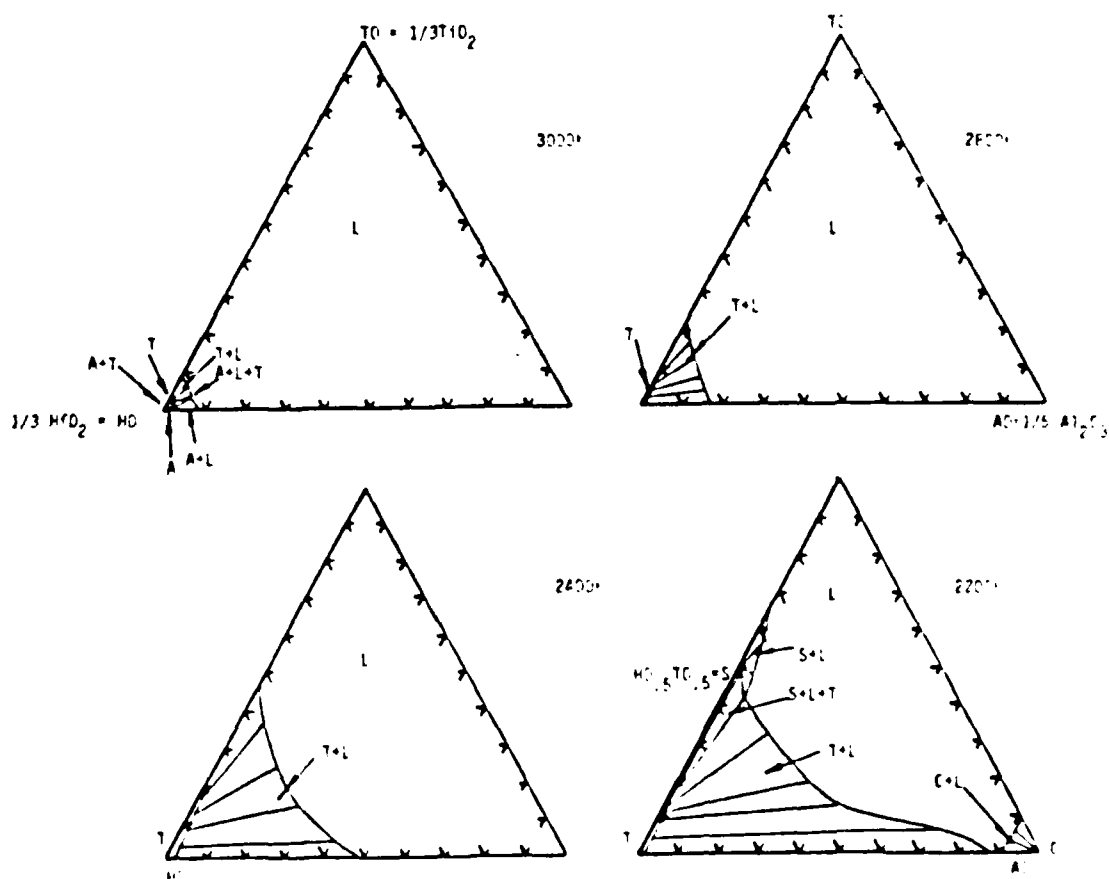
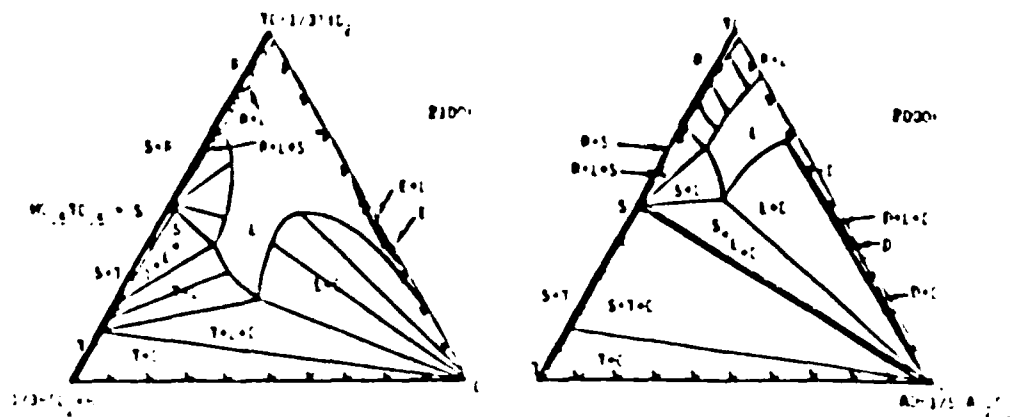


Figure 13. Calculated Isothermal Sections in $\text{TiO}-\text{Al}_2\text{O}_3-\text{H}_2\text{O}$



E = TO, 375 AC, 625 MC
 D = TO, 375 AC, 625 MC
 C Stable below 2000°C
 L Stable above 2000°C

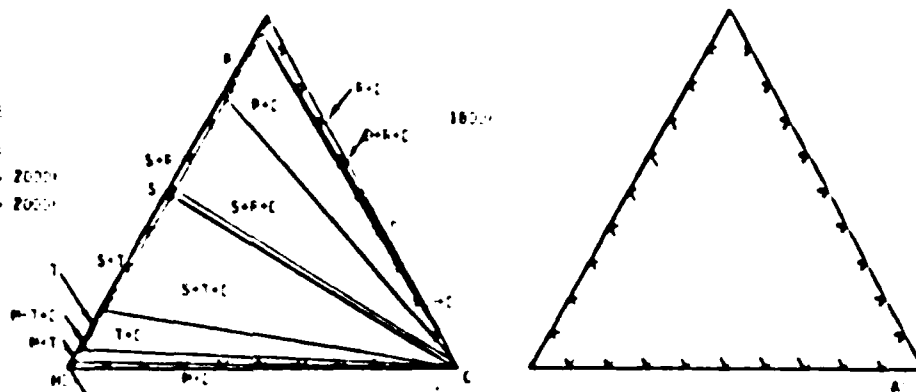
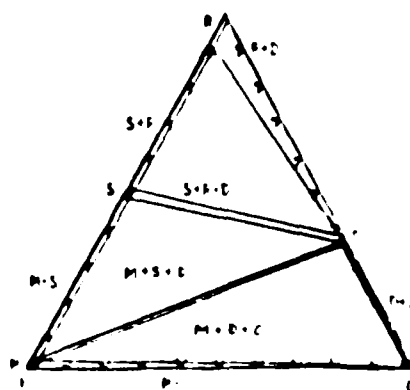
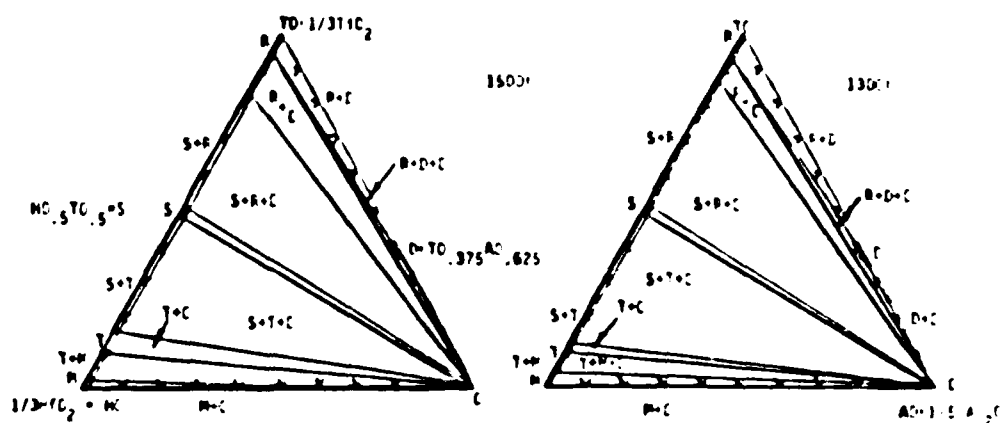


Figure 15. Calculated Isothermal Sections in TO-AC-MC



$$.545S + .455C = .727D + .273T$$

$$\Delta G = -6274 + 5.43T$$

0 at 1155°

Figure 16. Calculated Isothermal Sections in TO-AC-MC

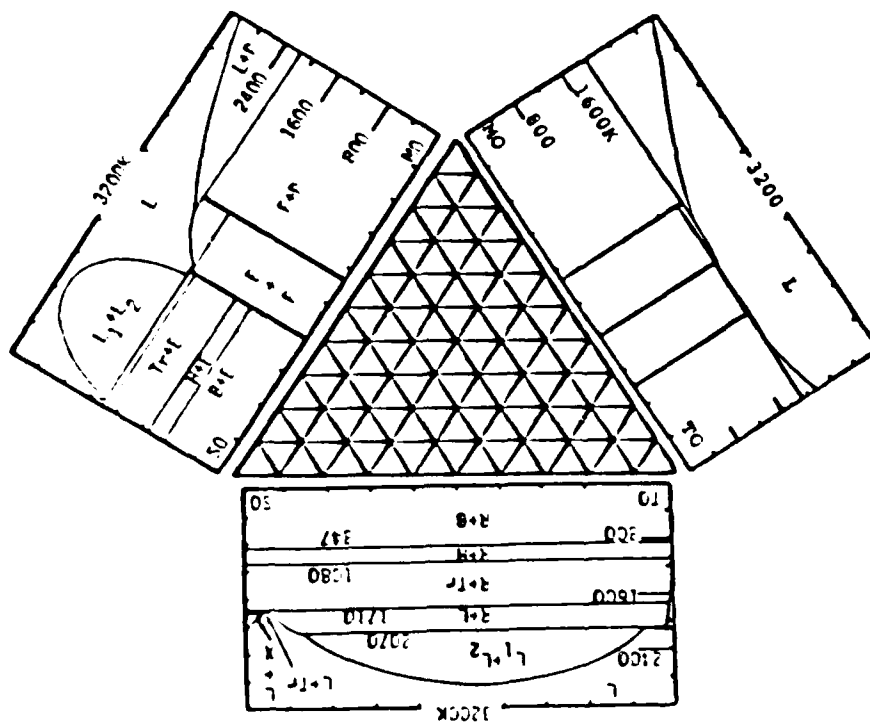


Figure 17 Calculated Isothermal Sections for the $Mn(1/2MnO) - TO(1/3TiO_2) - SC(1/3SnO_2)$ System

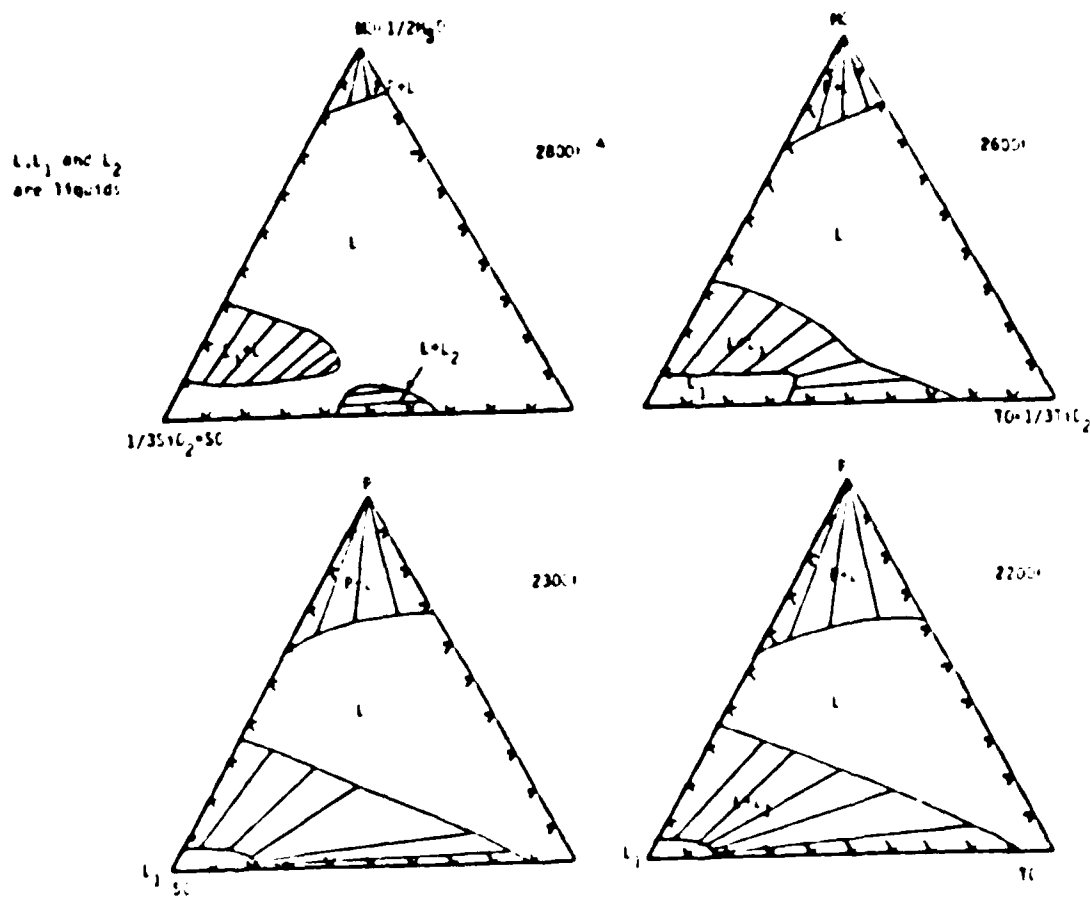
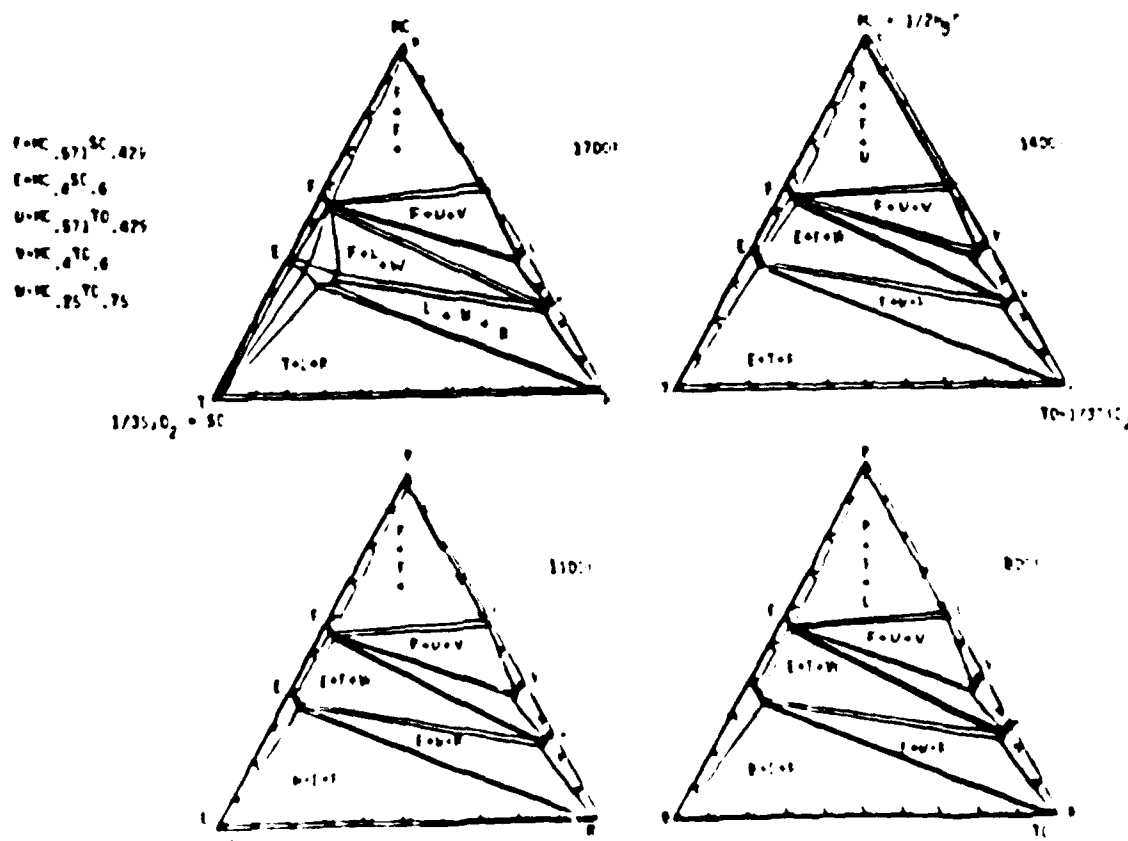
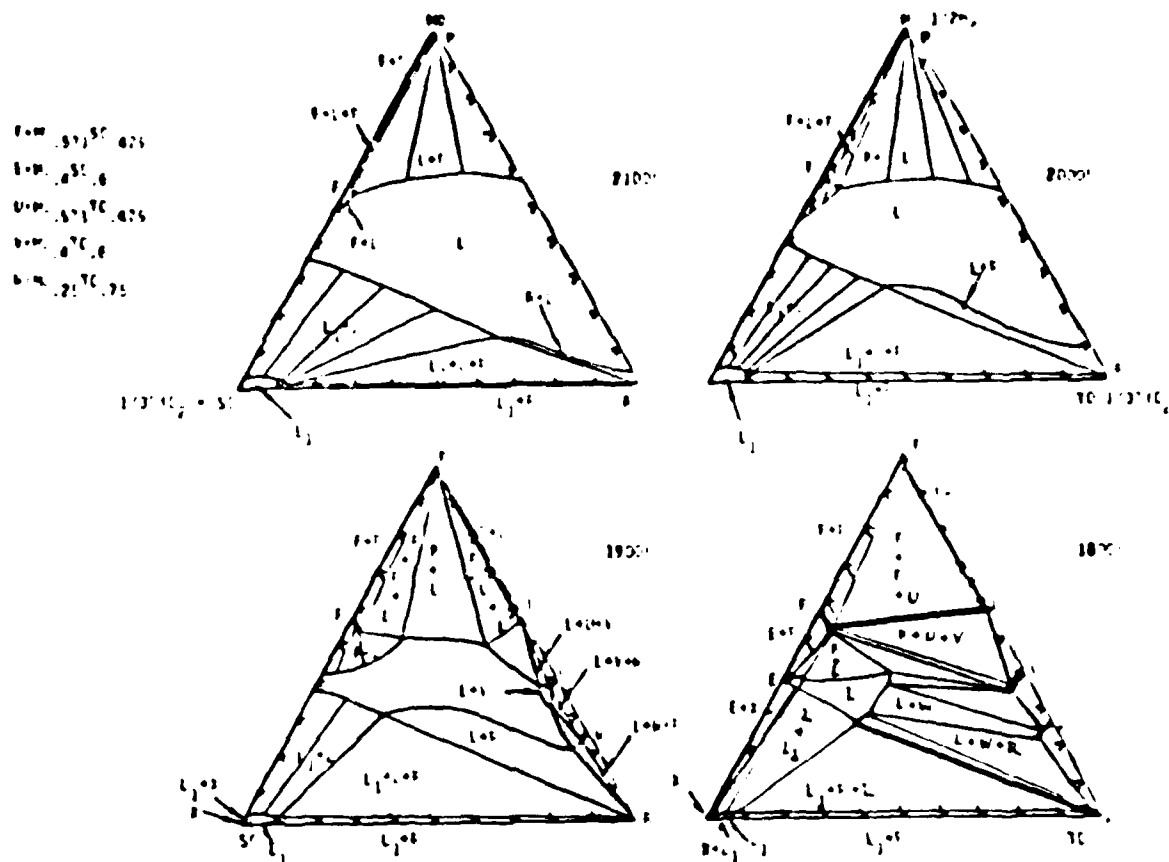


Figure 18 Calculated Isothermal Sections for the $Mn(1/2MnO) - TO(1/3TiO_2) - SC(1/3SnO_2)$ System



723, 2542, 4573 and 4574 of (9). The minimum liquid temperature is reported at 1663°K which compares favorably with the calculated section at 1700°K in Figure 20. At high temperatures the equilibria is dominated by the interaction between the liquid miscibility gaps emanating from the MO-SO and SO-TO edges.

SUMMARY

The forgoing set of binary and ternary examples show how model calculations of ceramic phase diagrams can be performed to guide development of new structural systems and to evaluate existing experimental data. It also provides a means for planning future experimental studies.

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V. CALCULATION OF METAL-OXYGEN, METAL-CARBONITRIDE AND CERAMIC PHASE DIAGRAMS WITH THE THERMOCALC SYSTEM

The THERMOCALC system is a databank for thermochemistry and metallurgy developed at the division of Physical Metallurgy of the Royal Institute of Technology (KTH) in Stockholm. Using the facilities of THERMOCALC one can tabulate thermodynamic data, calculate the heat change of chemical reactions and their driving force, evaluate equilibria for chemical systems and phase transformations and calculate various types of multicomponent phase diagrams by an automatic mapping procedure. The databank is quite general and can be applied to all systems where data assessed by a model implemented in the databank are available [1]. Examples of applications include nine component tool steels, metal carbide and nitride systems as well as binary and multicomponent alloy systems. The present work was directed toward illustrating the application to metal-oxygen carbonitride and oxide systems. Copies of the THERMOCALC system were provided for this exercise by the division of Physical Metallurgy at KTH in Stockholm and stored on VAX computers at the Massachusetts Institute of Technology in Cambridge, Massachusetts and the National Bureau of Standards in Gaithersburg, Maryland. These installations were accessed remotely using IBM-PC and AT models via modem and telephone. Tables 1-4 and Figures 1-14 show the results obtained in the course of this study which covered the Fe-Ni-C, Fe-Cr-C, Ti-C-N and Al_2O_3 - Y_2O_3 - ZrO_2 systems.

Table 1 and Figure 1 display the results for the Fe-Ni-C system [2] at 1873. Table 1 contains the THERMOCALC description modeled after reference 2 while Figure 1 shows the output depicting the phase diagram as a function of oxygen pressure (ordinate) and nickel/(nickel + iron) ratio (abscissa). The THERMOCALC output is shown in the left while the output generated at the FACIT (Facility for Chemical Thermodynamics) at Ecole Polytechnique in Montreal is shown in the right. The lower panels show the tie lines defining the phase boundaries of the three phase fields as well as those covering the two phase fields in the ternary. The abscissa is nickel/(nickel + iron).

TABLE 1

THERMOCALC DATA FOR THE Fe-Ni-O SYSTEM

...FROM GLENN ENERGY SYSTEM ON VMA JPS 17F

1-71 6. - 4. - 1

ALL DATA IN SI UNITS

FUNCTION, CALIB FOR 250.15 T 6000.00 T UNLESS OTHER LIMITS STATED

ELEMENT	STABLE ELEMENT REFERENCE	MASS	MW=MO	SIZE
C-12	Carbon	0.0000E+00	0.0000E+00	0.0000E+00
H-1	H	5.9847E+01	0.0000E+00	0.0000E+00
N-14	N	5.6095E+01	0.0000E+00	0.0000E+00
O-16	O	5.6095E+01	0.0000E+00	0.0000E+00
S-32	Sulfur	5.6095E+01	0.0000E+00	0.0000E+00

SPC 10-10-177	SPC 10-10-177
FI	FI
FI	FI
FI	FI
FI	FI

1. NAME: JAMES EARL RAY
2. DATE: 11-11-68
3. TIME: 10:00 AM
4. PLACE: 1000 1/2 N. 1ST ST. S. ST. LOUIS, MO.
5. TYPE: MURDER
6. REASON: UNKNOWN
7. SUSPECT: JAMES EARL RAY
8. VICTIM: MARTIN LUTHER KING, JR.
9. WEAPON: .38 CALIBER RIFLE
10. OTHER: 1000 1/2 N. 1ST ST. S. ST. LOUIS, MO.
11. COMMENTS: 1000 1/2 N. 1ST ST. S. ST. LOUIS, MO.
12. SIGNATURE: JAMES EARL RAY
13. ADDRESS: 1000 1/2 N. 1ST ST. S. ST. LOUIS, MO.
14. PHONE: 1000 1/2 N. 1ST ST. S. ST. LOUIS, MO.
15. OTHER: 1000 1/2 N. 1ST ST. S. ST. LOUIS, MO.

```

1. F10,F10C -W558 F00,F10C = -05000F1
2. F00,W10C -W558 F00,W10C = C10
3. F10,F1,W10C = -18394.1E-0101603-7*1-41369E-00-7*0.0
4. F00,F1,W10C = -18480.0E-0141177*0.0-01174E-00-7*0.0

```

1. SIGHTINGS, SITES : 1
2. CONTACTS: FILM : 0

```

C FEN1,G,F10,FEN1000+HISE,F10,FEN10+HISE,1,2,OLE_GAS,C10 * -100000
+G5,10,4*7-GFCF1
C FEN1,G,F10,FEN1000+HISE,F10,FEN10+HISE,1,2,OLE_GAS,C10 * -100000
+G5,10,4*7

```

REP-1711
S. PLATTICES, SITE 2 : 3
CONSTITUENTS: FI : 0

NEGATIVE, FICOM - 3 MORE FICOM, FICOM - 3 MORE FICOM, GAS, O, O -
100.00 7 2000.000 - 818010-340.0147-11.880147-100.00 7

[illegible][illegible]

1. DATE
2. LOCATION
3. TIME

* COMMUNICATIONS SECTION - HIGH SCHOOL AND COLLEGE STUDENTS
* COMMUNITY OUTREACH

[illegible]

TABLE 2

THERMOCALC DATA FOR THE Fe-Cr-O SYSTEM

EXTRACT FROM COMPILE ENERGY SYSTEM IN CAL/GRM 120

DATE EC- 5- 7

ALL DATA IN SI UNITS

FUNCTIONS VALUE FOR 258.15 TO 2000.00 K UNLESS OTHER LIMITS SPECIFIED

ELEMENT	STABLE ELEMENT REFERENCE	MASS	WSE-HC	ESSE
0 04	04000	0.0000E+00	0.0000E+00	0.0000E+00
1 01	001	5.5958E+01	0.0000E+00	0.0000E+00
2 01	001	5.5847E+01	0.0000E+00	0.0000E+00
3 01	1 00101_GAS	1.5555E+01	0.0000E+00	0.0000E+00

SPECIES	STOICHIOMETRY
0 04	04
1 01	01
2 01	01
3 01	01

FE
EXCESS MODEL IS FEILICH-HIETER-MUGGIAND
CONSTITUENTS: CR,FE

G(FE,CR;0)-HSEB(FE,CR;0) = +10400-6.276*T
G(FE,FE;0)-HSEB(FE,FE;0) = +GFCFFE
G(FE,CR,FE;0) = +7405-6.276*T

FE-1,0,0,0
2 SUBLATTICES, SITES 1 : 1
CONSTITUENTS: FE : C

G(FE-1,C;0)-HSEB(FE,CR;0)-HSEB(1/2POLE_GAS,C;0) = -30000
+25.0014*T-GFCFFE

FE-ATITE
EXCESS MODEL IS FEILICH-HIETER-MUGGIAND
2 SUBLATTICES, SITES 2 : 3
CONSTITUENTS: CR,FE : C

G(FE-ATITE,CR;0;0) = 2 HSEB(FE,CR;0) - 3 HSEB(1/2POLE_GAS,C;0) =
500.00 TO 2000.00: -1129700+255.24*T-5.445*T*LOG(T)
G(FE-ATITE,FE;0;0) = 2 HSEB(FE,FE;0) - 3 HSEB(1/2POLE_GAS,C;0) =
500.00 TO 2000.00: -823010+240.51*T-11.585*T*LOG(T)

SPINEL
EXCESS MODEL IS FEILICH-HIETER-MUGGIAND
3 SUBLATTICES, SITES 1.66666 : 1.33333 : 1.33333
CONSTITUENTS: CR,FE : FE : C

G(SPINEL,CR;FE;0;0) = 0.66666 HSEB(FE,CR;0) - 0.33333 HSEB(FE,FE;0)
- 1.33333 HSEB(1/2POLE_GAS,C;0) =
258.15 TO 2000.00: +0.33333*GFCFFE
G(SPINEL,FE;FE;0;0) = 0.55555 HSEB(FE,FE;0) - 1.33333 HSEB(1/2POLE_GAS,C
;0) = -358545+20.006*T+5.255*T*LOG(T)-GFCFFE

ALUMINA
2 SUBLATTICES, SITES 1 : 1
CONSTITUENTS: FE : C

G(ALUMINA,FE;0;0)-HSEB(FE,FE;0)-HSEB(1/2POLE_GAS,C;0) = -264805
+25.0014*T-GFCFFE

SPECIES	STATUS	VALUE/FUNCTION
1 04	00000000	0.3144E00E+00
2 0104	00000000	+5*T*LOG(5.80122E-01)*3
3 000000	00000000	
4 000000	00000000	
5 000000	00000000	
6 000000	00000000	
7 000000	00000000	
8 000000	00000000	
9 000000	00000000	
10 000000	00000000	
11 000000	00000000	
12 000000	00000000	
13 000000	00000000	
14 000000	00000000	
15 000000	00000000	
16 000000	00000000	
17 000000	00000000	
18 000000	00000000	
19 000000	00000000	
20 000000	00000000	
21 000000	00000000	
22 000000	00000000	
23 000000	00000000	
24 000000	00000000	
25 000000	00000000	
26 000000	00000000	
27 000000	00000000	
28 000000	00000000	
29 000000	00000000	
30 000000	00000000	
31 000000	00000000	
32 000000	00000000	
33 000000	00000000	
34 000000	00000000	
35 000000	00000000	
36 000000	00000000	
37 000000	00000000	
38 000000	00000000	
39 000000	00000000	
40 000000	00000000	
41 000000	00000000	
42 000000	00000000	
43 000000	00000000	
44 000000	00000000	
45 000000	00000000	
46 000000	00000000	
47 000000	00000000	
48 000000	00000000	
49 000000	00000000	
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51 000000	00000000	
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89 000000	00000000	
90 000000	00000000	
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92 000000	00000000	
93 000000	00000000	
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97 000000	00000000	
98 000000	00000000	
99 000000	00000000	
100 000000	00000000	

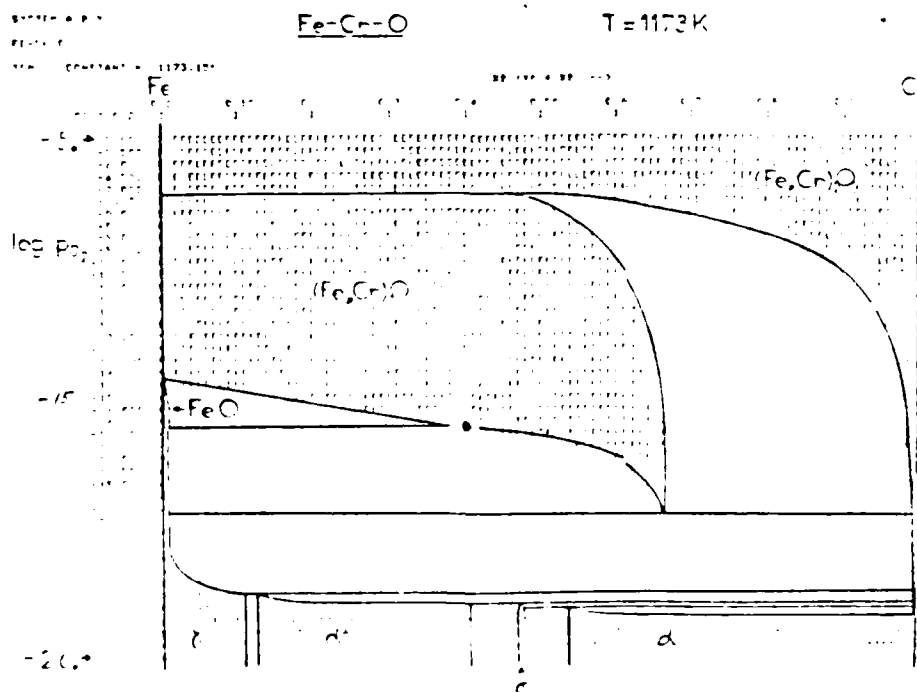
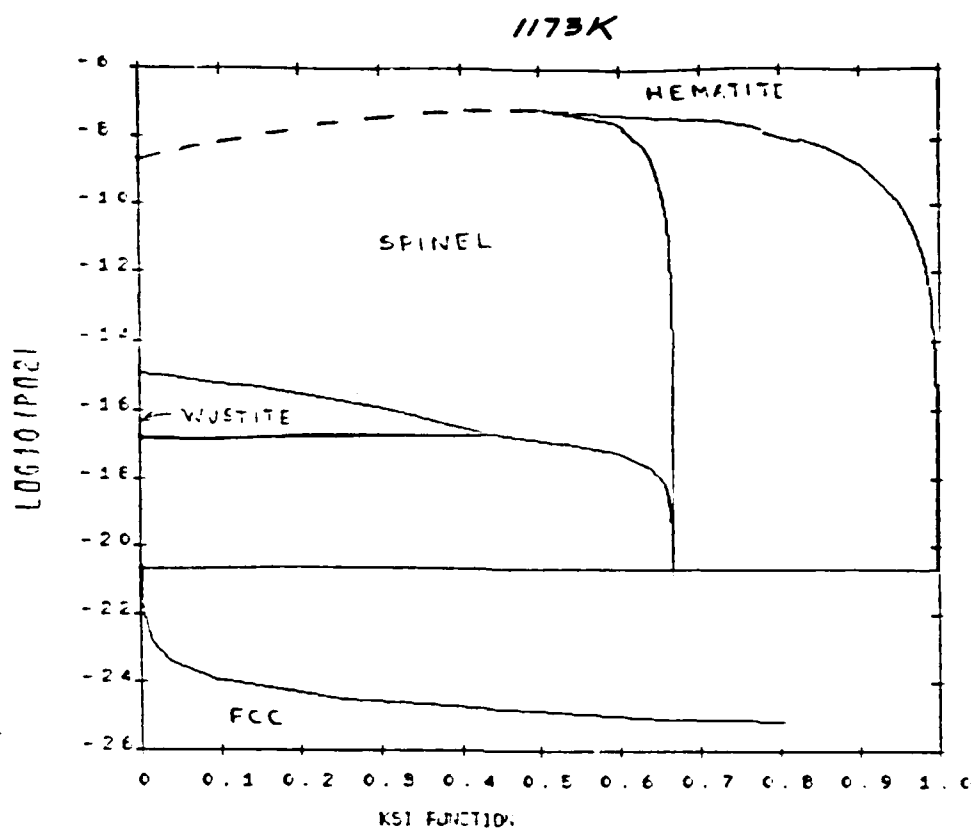


Figure 2. Thermocalc Output for fcc, Hematite, Spinel and Wustite (above) compare with FACT Output (below)

TABLE 3 THERMOCALC DATA FOR THE Ti-C-N SYSTEM

OUTPUT FROM DISEE ENERGY SYSTEM ON VAX/VMS KTH

DATE 87- 4-12

ALL DATA IN SI UNITS

FUNCTIONS VALID FOR 298.15 TO 6000.00 K UNLESS OTHER LIMITS STATED

ELEMENT	STABLE	ELEMENT	REFERENCE	MASS	H298-H0	S298
0 VA	VACUUM			0.0000E+00	0.0000E+00	0.0000E+00
1 C	GRAPHITE			1.2011E+01	0.0000E+00	0.0000E+00
2 N	1/2 MOLE GAS			1.4007E+01	0.0000E+00	0.0000E+00
3 Ti	HCP			4.7880E+01	0.0000E+00	0.0000E+00

SPECIES	STOICHIOMETRY
1 C	C
2 N	N
3 Ti	Ti
4 VA	VA

GRAPHITE
CONSTITUENTS: C

$$G(\text{GRAPHITE}, \text{C}; 0) - H298(\text{GRAPHITE}, \text{C}; 0) = 0.0$$

NITROGEN
CONSTITUENTS: N

$$G(\text{NITROGEN}, \text{N}; 0) - H298(1/2 \text{ MOLE GAS}, \text{N}; 0) = 0.0$$

TiCN
2 SUBLATTICES, SITES 2 + 1
CONSTITUENTS: Ti + N

$$G(\text{TiCN}, \text{Ti}; 0) - H298(1/2 \text{ MOLE GAS}, \text{N}; 0) - 2 H298(\text{HCP}, \text{Ti}; 0) = -504976 + 18.16 \text{ J}$$

Ti FCC
EXCESS MODEL 16 REDLICH-ISTER, MUGGIANU
2 SUBLATTICES, SITES 1 + 1
CONSTITUENTS: Ti + C, N, VA

$$\begin{aligned} G(\text{Ti FCC}, \text{Ti}; 0) - H298(\text{GRAPHITE}, \text{C}; 0) - H298(\text{HCP}, \text{Ti}; 0) &= -163979 + 10.577 \text{ J} \\ G(\text{Ti FCC}, \text{Ti}; 0) - H298(1/2 \text{ MOLE GAS}, \text{N}; 0) - H298(\text{HCP}, \text{Ti}; 0) &= -333323 + 2.6214 \text{ J} \\ G(\text{Ti FCC}, \text{Ti}; 0) - H298(\text{HCP}, \text{Ti}; 0) &= 3347.2 \\ G(\text{Ti FCC}, \text{Ti}; 0, \text{VA}; 0) &= -67358 + 1.5062 \text{ J} \\ G(\text{Ti FCC}, \text{Ti}; 0, \text{VA}; 1) &= -163343 + 46.15 \text{ J} \\ G(\text{Ti FCC}, \text{Ti}; 0, \text{N}; 0) &= -22073 \\ G(\text{Ti FCC}, \text{Ti}; 0, \text{N}; 1) &= -10460 + 18.41 \text{ J} \end{aligned}$$

Ti HCP
EXCESS MODEL 16 REDLICH-ISTER, MUGGIANU
2 SUBLATTICES, SITES 1 + 2
CONSTITUENTS: Ti + C, N, VA

$$\begin{aligned} G(\text{Ti HCP}, \text{Ti}; 0) - H298(\text{GRAPHITE}, \text{C}; 0) - H298(\text{HCP}, \text{Ti}; 0) &= +20936.8 + 0.0507 \text{ J} \\ G(\text{Ti HCP}, \text{Ti}; 0) - H298(1/2 \text{ MOLE GAS}, \text{N}; 0) - H298(\text{HCP}, \text{Ti}; 0) &= -466203 + 0.77 \text{ J} \\ G(\text{Ti HCP}, \text{Ti}; 0, \text{VA}; 0) - H298(\text{HCP}, \text{Ti}; 0) &= +4351 + 3.7656 \text{ J} \\ G(\text{Ti HCP}, \text{Ti}; 0, \text{VA}; 1) &= -4184 \text{ J} \\ G(\text{Ti HCP}, \text{Ti}; 0, \text{N}; 0) &= -8768 \text{ J} \end{aligned}$$

Ti HCP
EXCESS MODEL 16 REDLICH-ISTER, MUGGIANU
2 SUBLATTICES, SITES 2 + 1
CONSTITUENTS: Ti + C, N, VA

$$\begin{aligned} G(\text{Ti HCP}, \text{Ti}; 0) - H298(1/2 \text{ MOLE GAS}, \text{N}; 0) - 2 H298(\text{HCP}, \text{Ti}; 0) &= -393046 + 0.7047 \text{ J} \\ G(\text{Ti HCP}, \text{Ti}; 0, \text{VA}; 0) - H298(\text{HCP}, \text{Ti}; 0) &= 0.0 \\ G(\text{Ti HCP}, \text{Ti}; 0, \text{VA}; 1) &= -41775 + 14.392 \text{ J} \end{aligned}$$

TiTiC
EXCESS MODEL 16 REDLICH-ISTER, MUGGIANU
2 SUBLATTICES, SITES 1 + 1

$$\begin{aligned} G(\text{TiTiC}, \text{Ti}; 0) - H298(\text{GRAPHITE}, \text{C}; 0) &= +114210.2 - 27.196 \text{ J} \\ G(\text{TiTiC}, \text{Ti}; 0) - H298(1/2 \text{ MOLE GAS}, \text{N}; 0) &= -8890.168 + 66.9632 \text{ J} \\ G(\text{TiTiC}, \text{Ti}; 0, \text{VA}; 0) - H298(\text{HCP}, \text{Ti}; 0) &= +10564.91 - 12.1336 \text{ J} \\ G(\text{TiTiC}, \text{Ti}; 0, \text{VA}; 1) &= -57000 + 47.932 \text{ J} \\ G(\text{TiTiC}, \text{Ti}; 0, \text{N}; 0) &= +107679 + 41.971 \text{ J} \\ G(\text{TiTiC}, \text{Ti}; 0, \text{N}; 1) &= +108579 + 71.70 \text{ J} \\ G(\text{TiTiC}, \text{Ti}; 0, \text{N}; 2) &= +107479 + 41.71 \text{ J} \end{aligned}$$

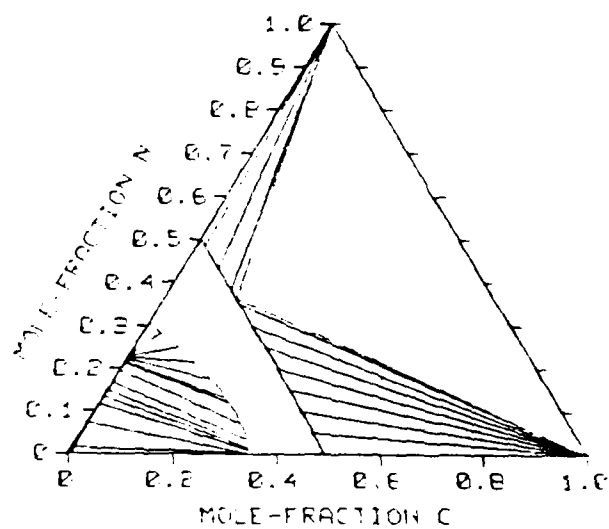
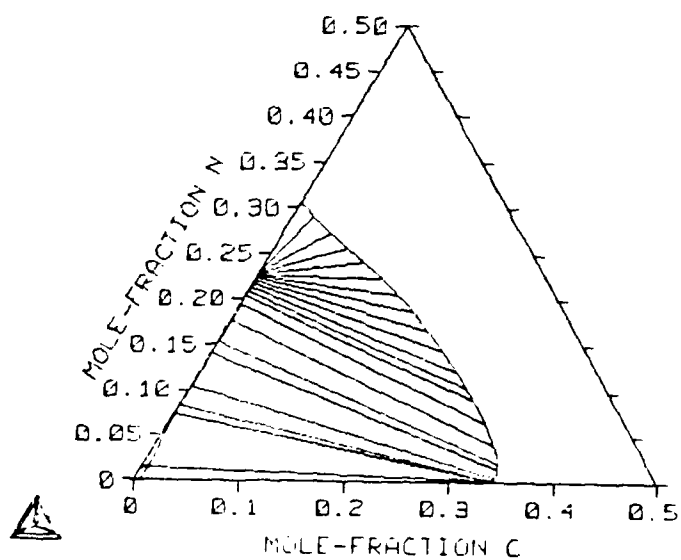
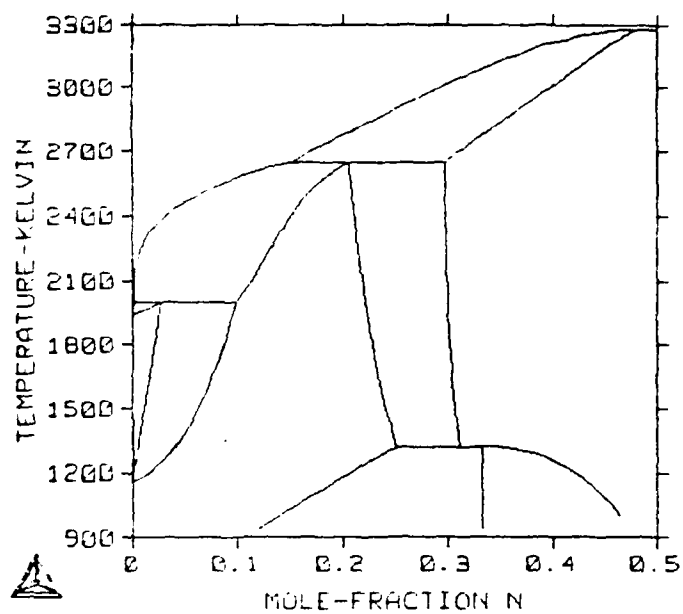


Figure 3. Calculated Ti-N Phase Diagram and
Ti-N-C Isothermal Section at 1600K from
the Thermocalc System.

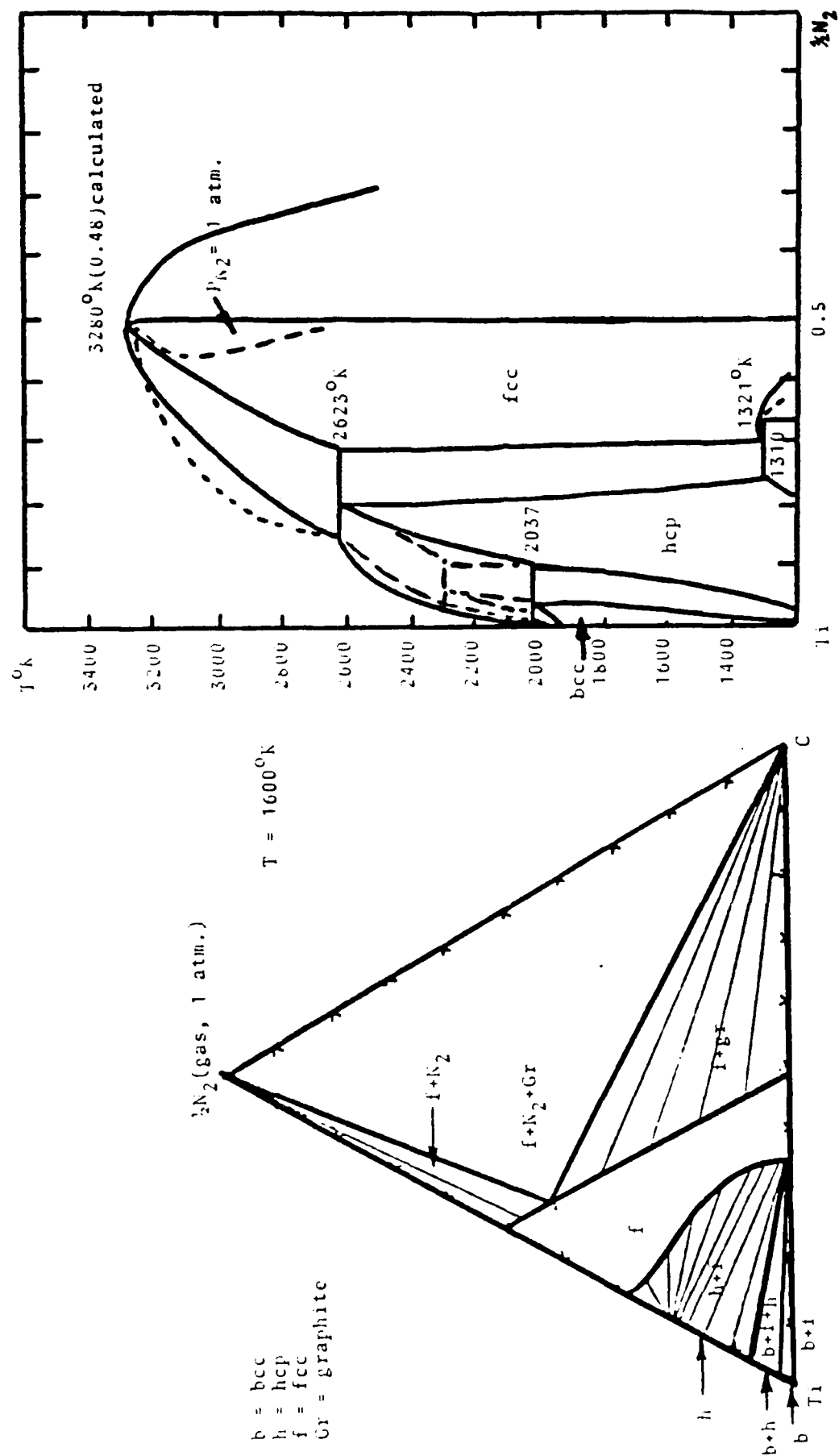


Figure 4. Calculated Ti-N and Ti-N-C Isothermal section at 1600K
 From the ManIabs System.

TABLE 4

SUMMARY OF LATTICE STABILITY, SOLUTION AND COMPOUND
PHASE PARAMETERS FOR THE CALCULATION OF AO-ZO-YO

Summary of Lattice Stability Parameters, T in Kelvins

(All units in Joules per gram atom (mole of atoms), T in Kelvins)

AO=(1/5) Al_2O_3 , ZO=(1/3) ZrO_2 , YO=(1/5) Y_2O_3 , L= Liquid

C= Corundum, Y=high temperature (hexagonal) Y_2O_3 , B=bcc

(Mn_2O_3) Y_2O_3 , A= ZrO_2 cubic, T= ZrO_2 Tetragonal, M= ZrO_2 Monoclinic

ZOZOLA=(1/3) ZrO_2 (liquid)-(1/3) ZrO_2 (CaF_2) Cubic

AOAOLC=23640-10.209T

AOAOLA= - 9.832T

AOAOLT= -10.586T

ZOZOAT=1987-0.753T

ZOZOTM=2008-1.381T

ZOZOLA=29008- 9.832T

ZOZOLC= -10.209T

ZOZOLY=17991- 8.368T

ZOZOLB=24853- 9.205T

YOYOLY=22694- 8.368T

YOYOLB=26878-10.042T

YOYOLA=22615- 9.832T

YOYOAT=-8368- 0.753T

YOYOTM=-8368- 1.381T

Summary of Solution and Compound Phase Parameters

(All units in Joules per gram atom (mole of atoms), T in Kelvins)

LZCAO=17573, LAOZO=39748, TZCAO=TAOZO=AZCAO=AAOZO=62760

CAOZO=CZCAO=62760

Solution Phases

LZOYO= 14016 + 4.184T

LYOZO= 14016 + 4.184T

BZOYO= 2929 + 8.368T

BYOZO= 2929 + 8.368T

AZOYO=-12552 +11.297T

AYOZO= 837 + 4.602T

YZOYO= 2929 + 8.368T

YYOZO= 2929 + 8.368T

TZOYO=-12552 + 11.297T

TYOZO= 837 + 4.602T

MZOYO= 4184 + 11.297T

MYOZO= 837 + 4.602T

Compound Phase = P = (1/19) ($3\text{ZrO}_2 \cdot 2\text{Y}_2\text{O}_3$) = (1/19) $\text{Zr}_9\text{Y}_2\text{O}_{10}$

$\text{Zr}_{0.474}\text{Y}_{0.526}$, Base =A, Compound Parameter C= 31798 +

11.097T, Gibbs Energy of Formation from (1/3) ZrO_2 (L) and

(1/5) Y_2O_3 (L), $\Delta G_f = -286 - 2.86T$



$W = (Y_0, Z_0)_{0.375 \text{ Al}_2\text{O}_3, 0.625 \text{ Y}_2\text{O}_3}$
 $U = (Y_0, Z_0)_{0.667 \text{ Al}_2\text{O}_3, 0.333 \text{ Y}_2\text{O}_3}$

L = Liquid
 C = Corundum
 A = Cubic (CaF_2)
 B = Cubic
 T = Tetragonal

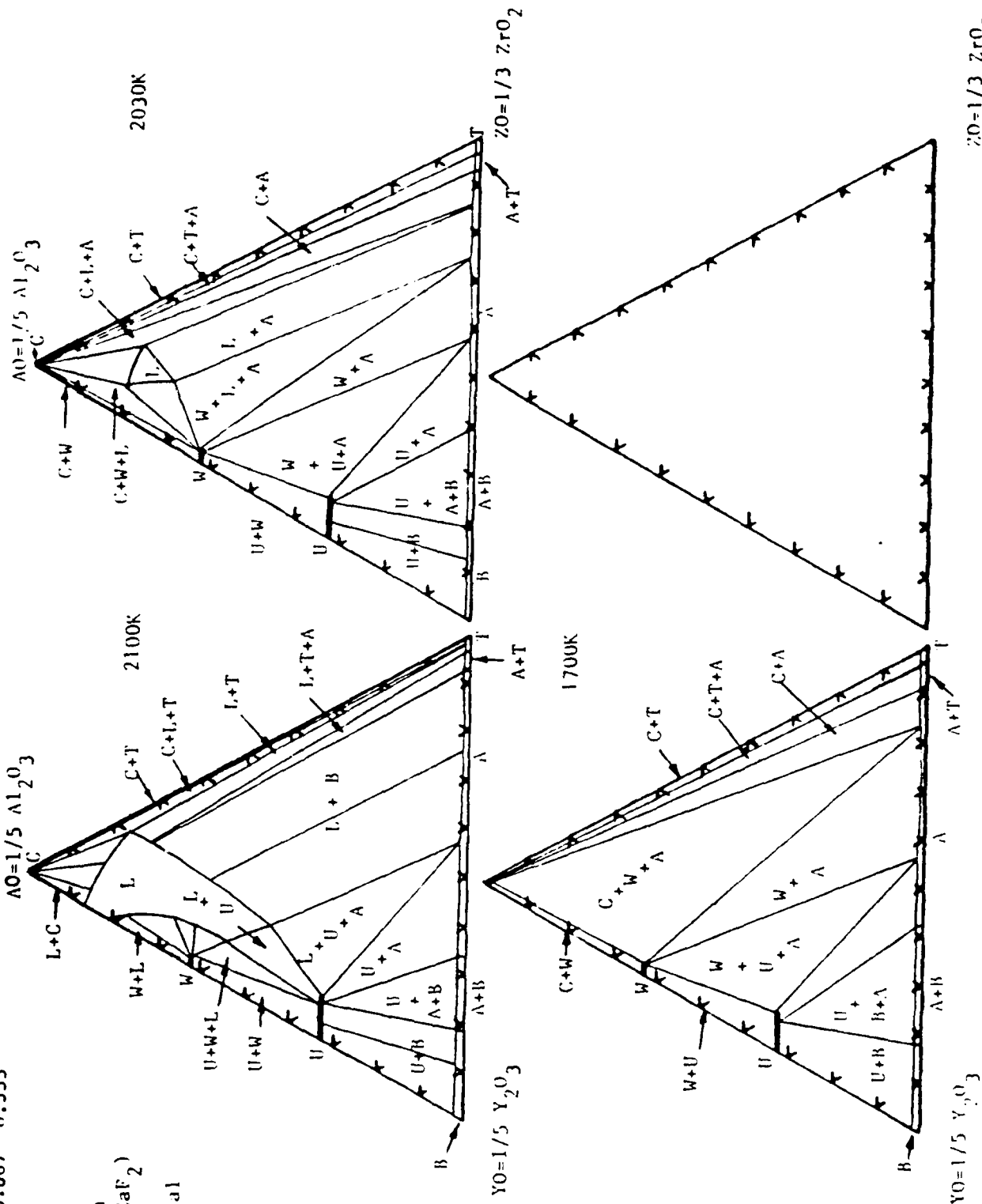


Figure 7. Calculated Isothermal Sections in the Al_2O_3 - Y_2O_3 - ZrO_2 System (ManLabs System).

TABLE 5

SUMMARY OF LATTICE STABILITY, SOLUTION AND COMPOUND PHASE PARAMETERS
FOR THE CALCULATION OF THE AM-ZM-YM SYSTEM (Joules, °K)

AM= $1/2\text{Al}_2\text{O}_3$, ZM= ZrO_2 , YM = $1/2\text{Y}_2\text{O}_3$

C=Corundum, Y=high temperature (hexagonal), B=bcc (Mn_2O_3) Y_2O_3

A= ZrO_2 cubic, T= ZrO_2 Tetragonal, M= ZrO_2 Monoclinic

LATTICE STABILITY PARAMETERS

AMAMLC=59100-25.5T	ZMZMAT=5961-2.26T
AMAMLA= -24.58T	ZMZMTM=6024-4.14T
AOAOLT= -24.47T	
	YMYMLY=56735-20.92T
ZMZMLA=87024-29.50T	YMYMLB=67195-25.11T
ZMZMLC= -30.63T	TMTMLA=56538-24.58T
ZMZMLY=53973-25.10T	YMYMAT=-20920-1.88T
ZMZMLB=74559-27.61T	YMYMTM=-20920-3.45T

SOLUTION PHASE PARAMETERS

LZMYM=-1674-37.66T	LYMZM=63178-37.66T	LZMAM=LAMZM=4184
BZMYM=-18410-25.10T	BYMZM=46442-31.38T	TZMAM=TZMAN=41840
AZMYM=-26778-33.47T	AYMZM=38074-39.75T	AZMAM=AMZM=41840
TZMYM=-26778-33.47T	TYMAM=38074-39.75T	CZMAM=CAMZM=41840
YZMYM=-18410-25.10T	YYMZM=46442-31.38T	
MZMYM=7950-33.47T	MYMZM=38074-39.75T	
LYMAM=-66944	LAMYM=-66944	
YYMAM=-20920	YAMYM=-20920	
BYMAM=-20920	BAMYM=-20920	
CYMAM=-20920	CAMYM=-20920	

COMPOUND PARAMETERS

P=ZM_{.429}YM_{.571}= $1/7(3\text{ZrO}_2 \cdot 2\text{Y}_2\text{O}_3)$; Base phase=A; C=84308-17.99T

U=YM_{.677}AM_{.333}= $1/6(2\text{Y}_2\text{O}_3 \cdot \text{Al}_2\text{O}_3)$; Base phase=B; C=67362+17.49T

V=YM_{.5}AM_{.5}= $1/4(\text{Y}_2\text{O}_3 \cdot \text{Al}_2\text{O}_3)$; Base phase=C; C=24058+60.25T

W=YM_{.375}AM_{.625}= $1/16(5\text{Y}_2\text{O}_3 \cdot 3\text{Al}_2\text{O}_3)$; Base phase=C; C=135980-2.89T

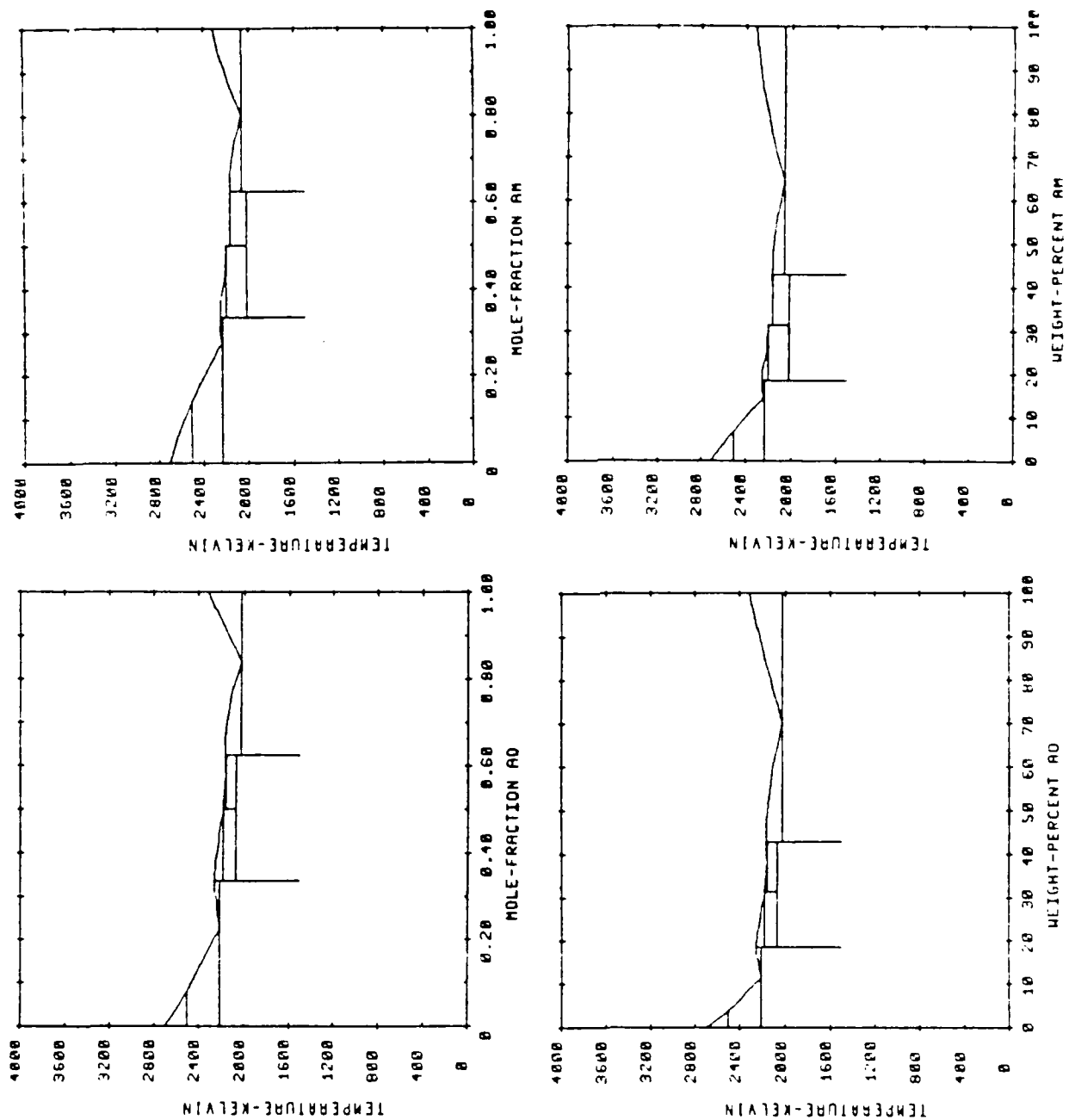


Figure 8. Thermocouple output for calculation of YO-AO and YM-AM Systems as a function of mol fraction and weight percent AO and AM (see Figure 5)

$$ZO = 1/3 \text{ Zr O}_2$$

$$ZM = \text{Zr O}_2$$

$$YO = 1/5 \text{ Y}_2 \text{ O}_3$$

$$YM = 1/2 \text{ Y}_2 \text{ O}_3$$

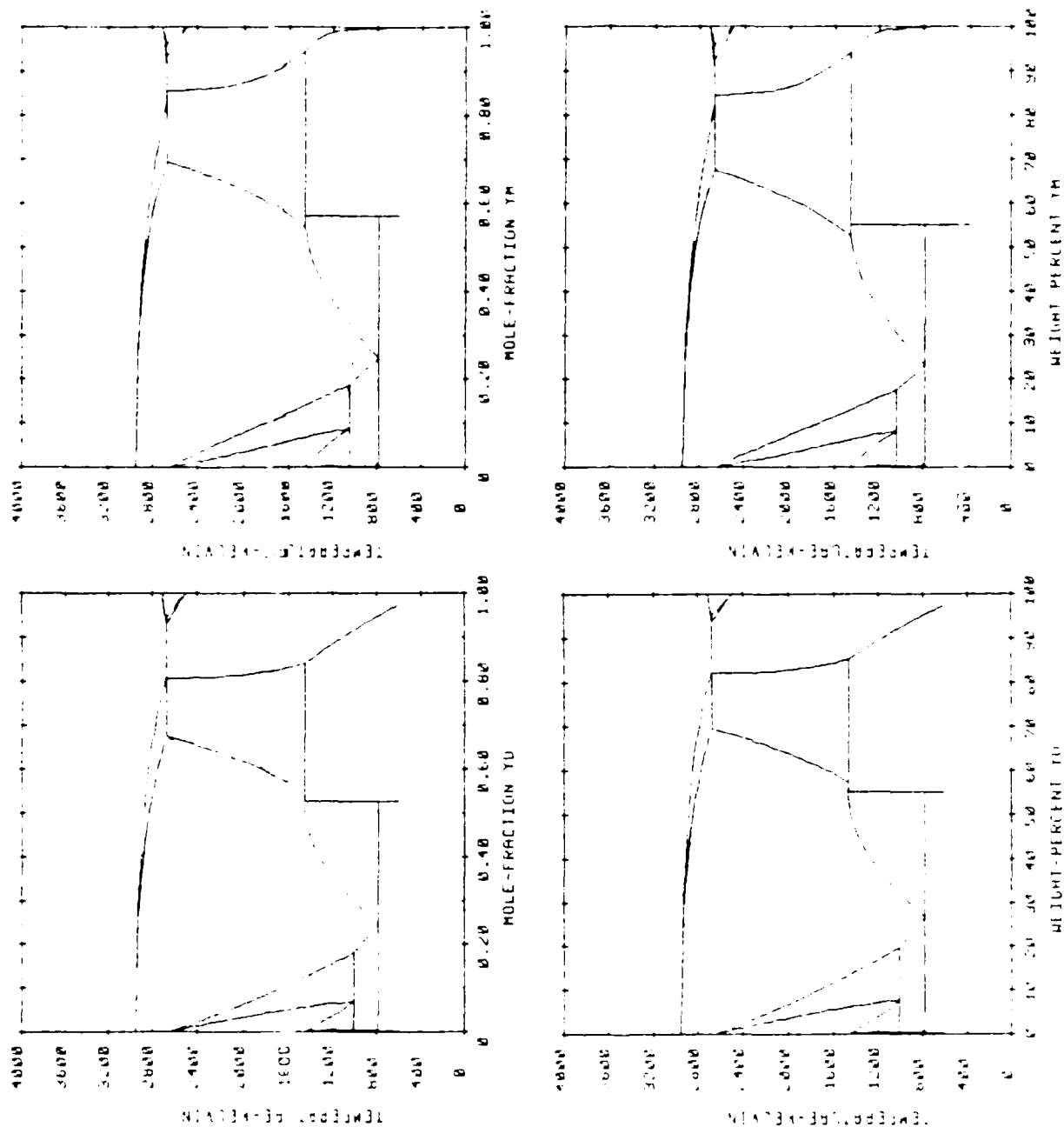


Figure 9. Thermocalc output for calculation of ZO-YO and ZM-YM

as a function of mol fraction and weight percent

YO and YM. (See Figure 5)

$ZO = 1/3 \text{ ZrO}_2$
 $ZM = \text{ZrO}_2$
 $AO = 1/5 \text{ Al}_2\text{O}_3$
 $AM = 1/2 \text{ Al}_2\text{O}_3$

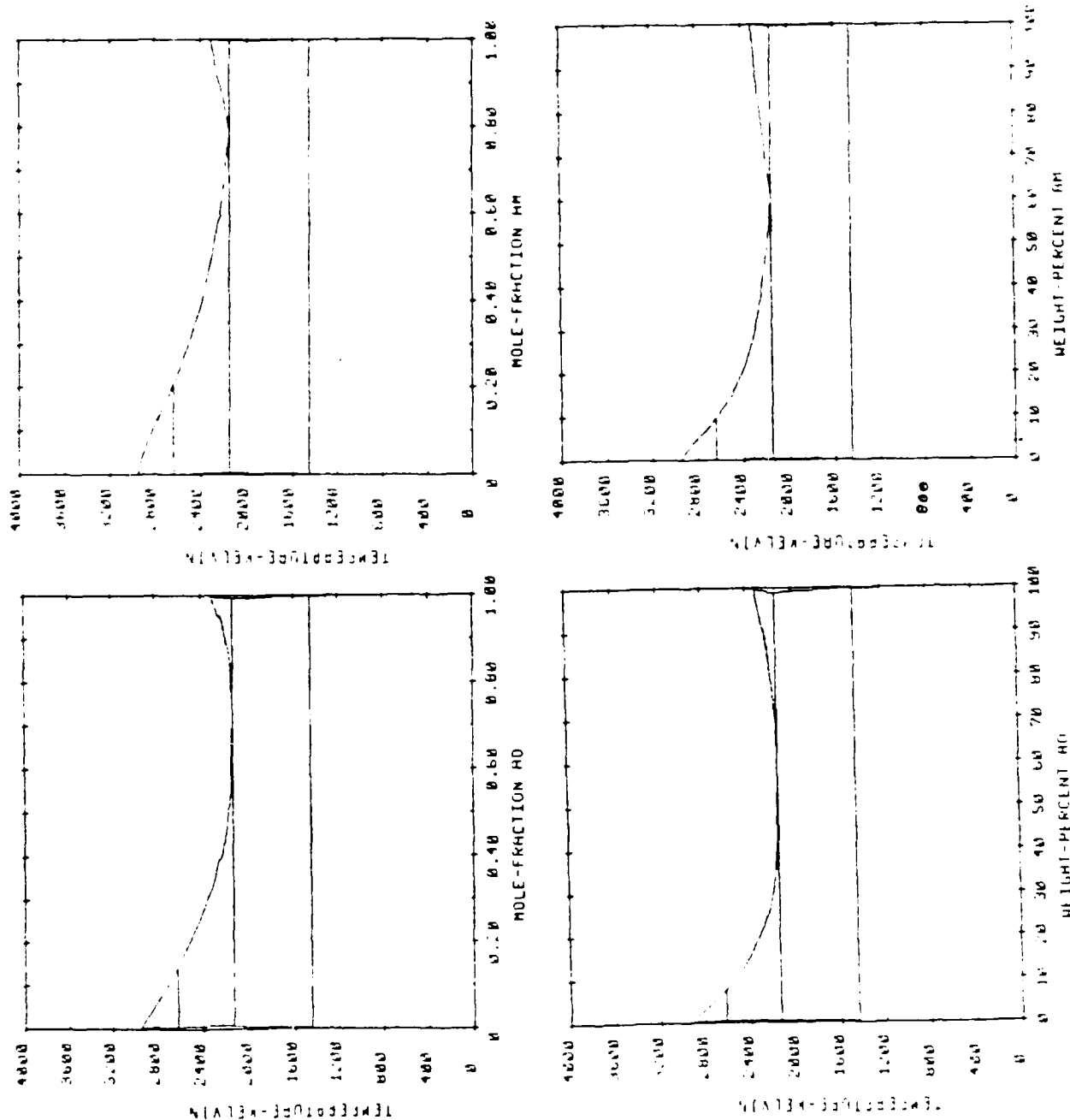


Figure 10 Thermocalc Output for calculation of ZO-AO and ZM-YM
 as a function of mol fraction and weight percent
 AO and AM (See Figure 5).

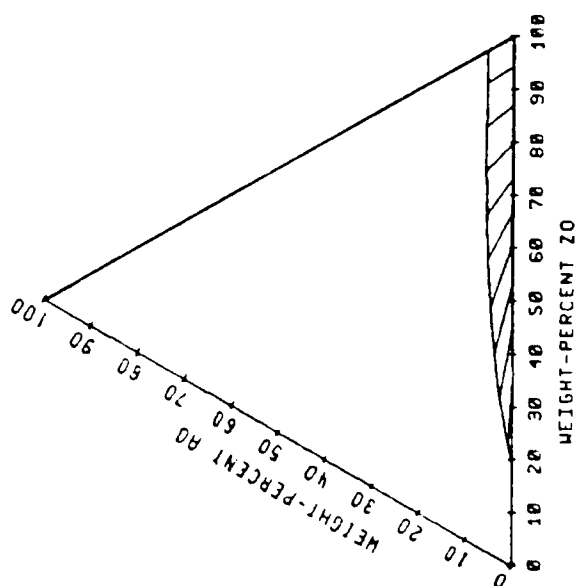
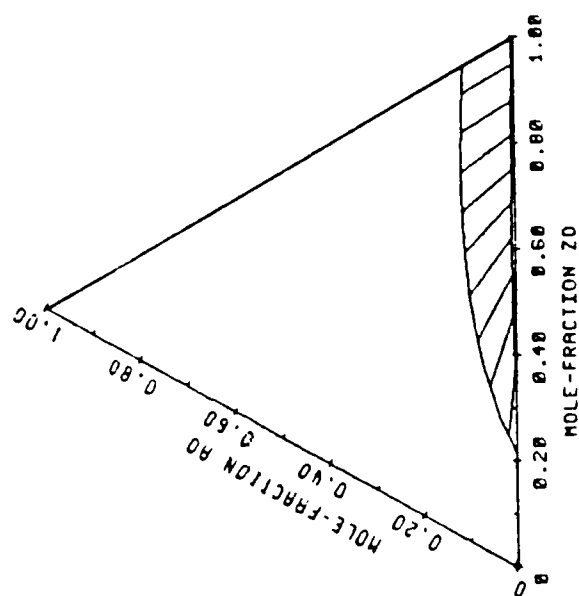
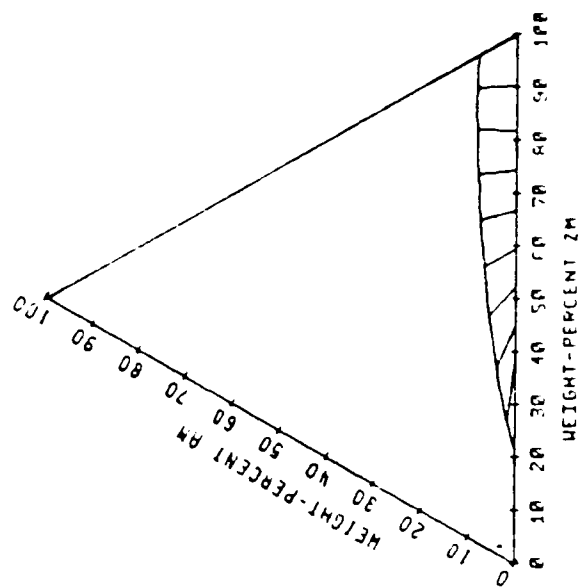
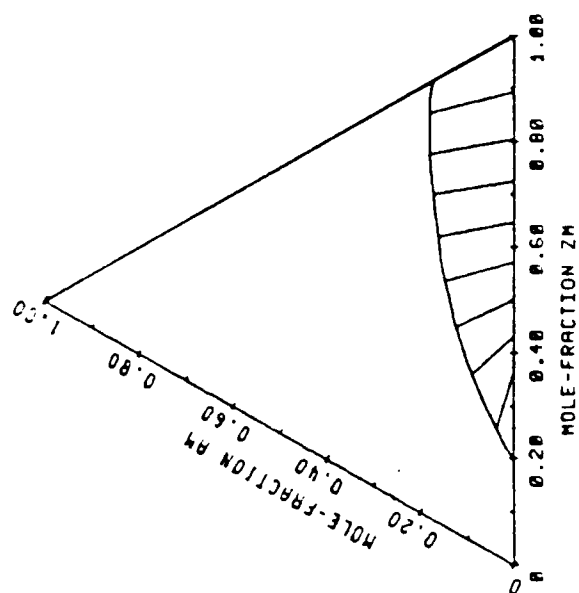


Figure 11 Thermocalc output for AO-ZO-YO and AM-ZM-YM at 2700K
(See Figure 6)

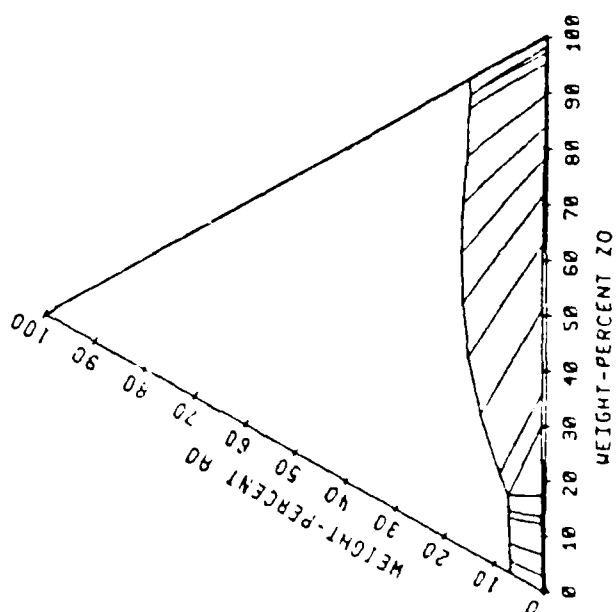
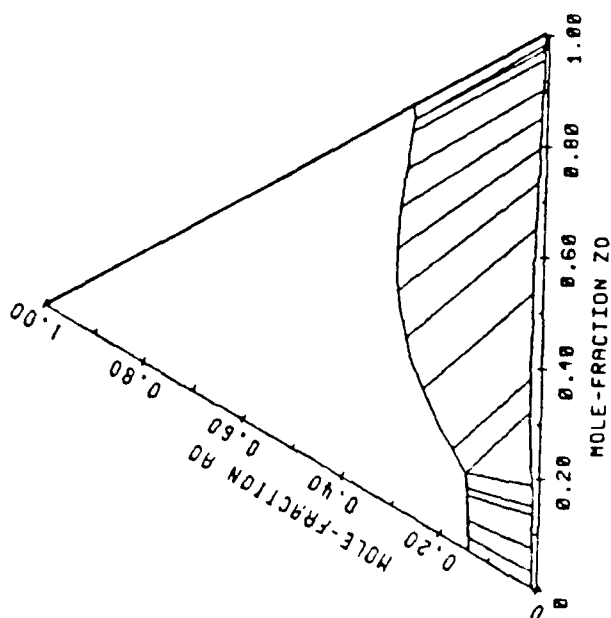
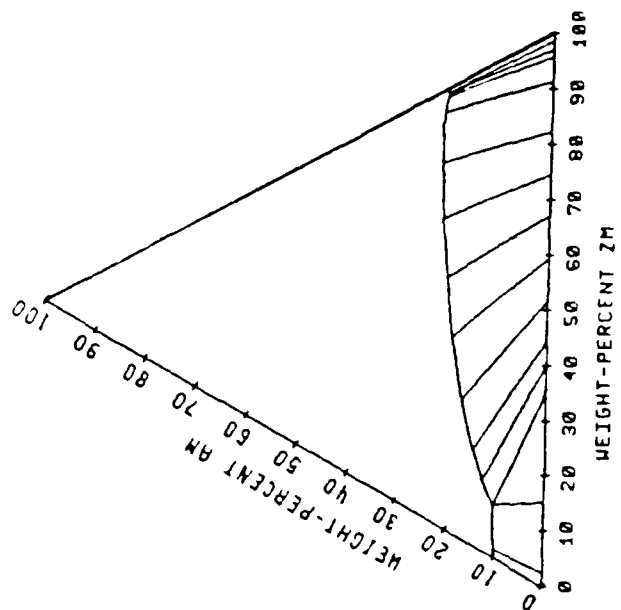
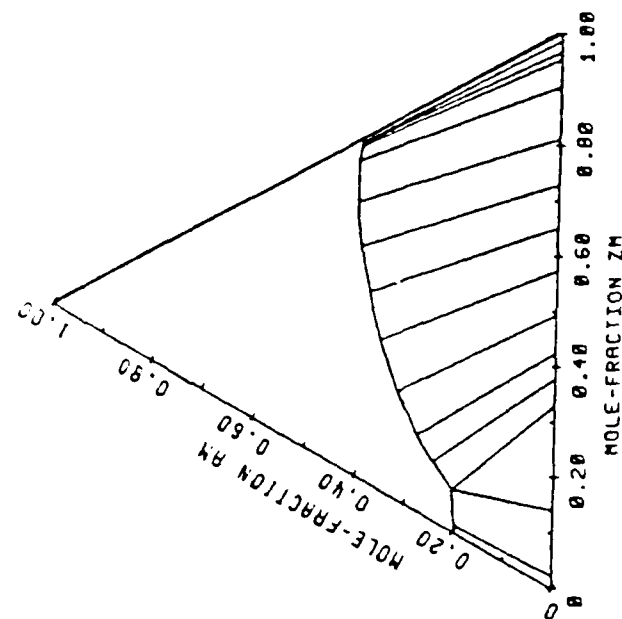


Figure 12 Thermocalc Output for AO-ZO-YO and AM-ZM-YM at 2400K
(See Figure 6)

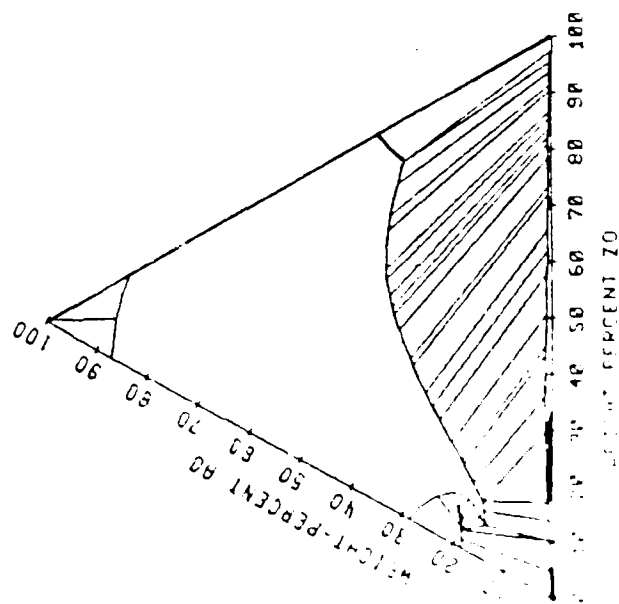
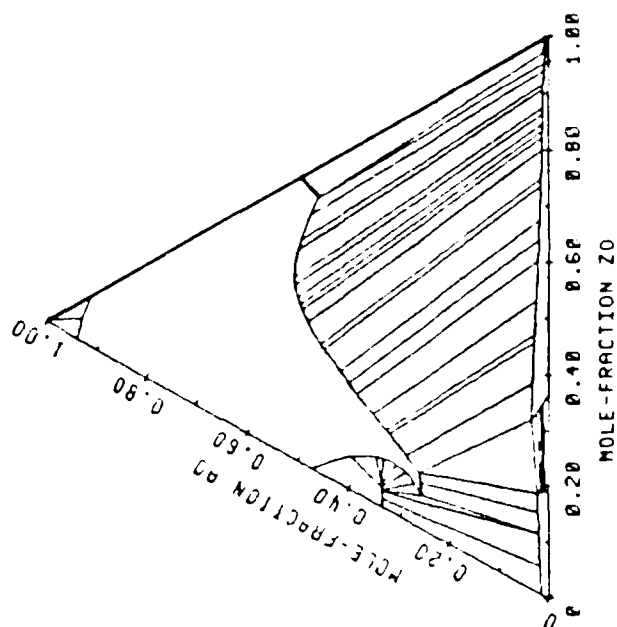
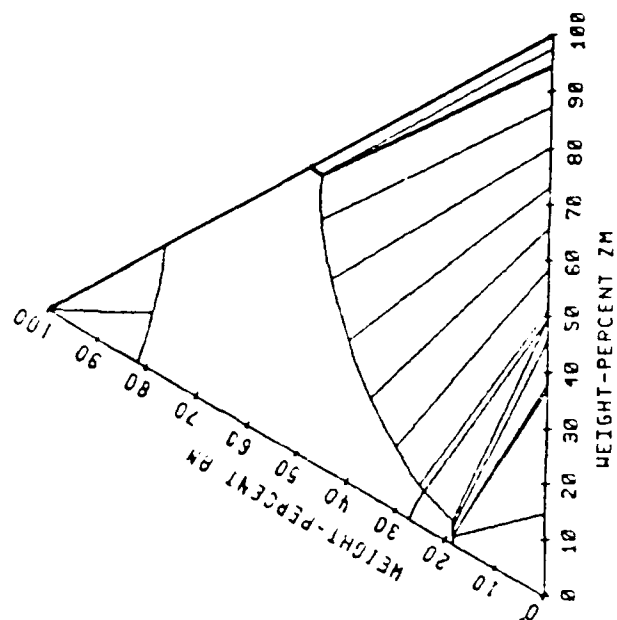
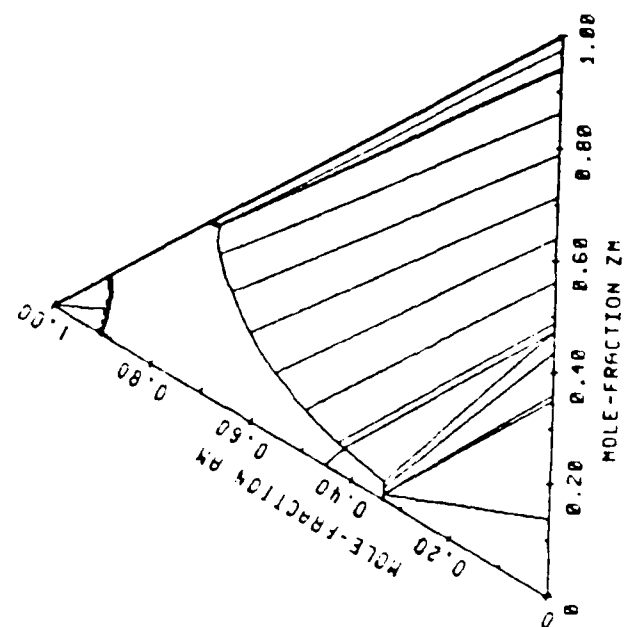


Figure 7) Thermocouple Output for AO-ZO-YO and AM-ZM-YM at 2200K

(See Figure 7)

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APPLICATION OF COMPUTER METHODS FOR CALCULATION OF
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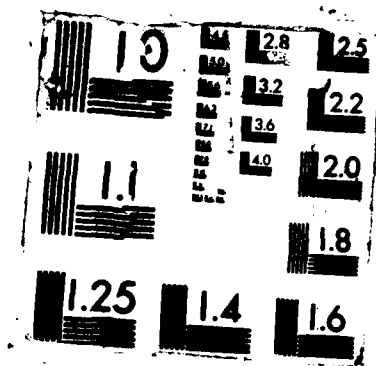
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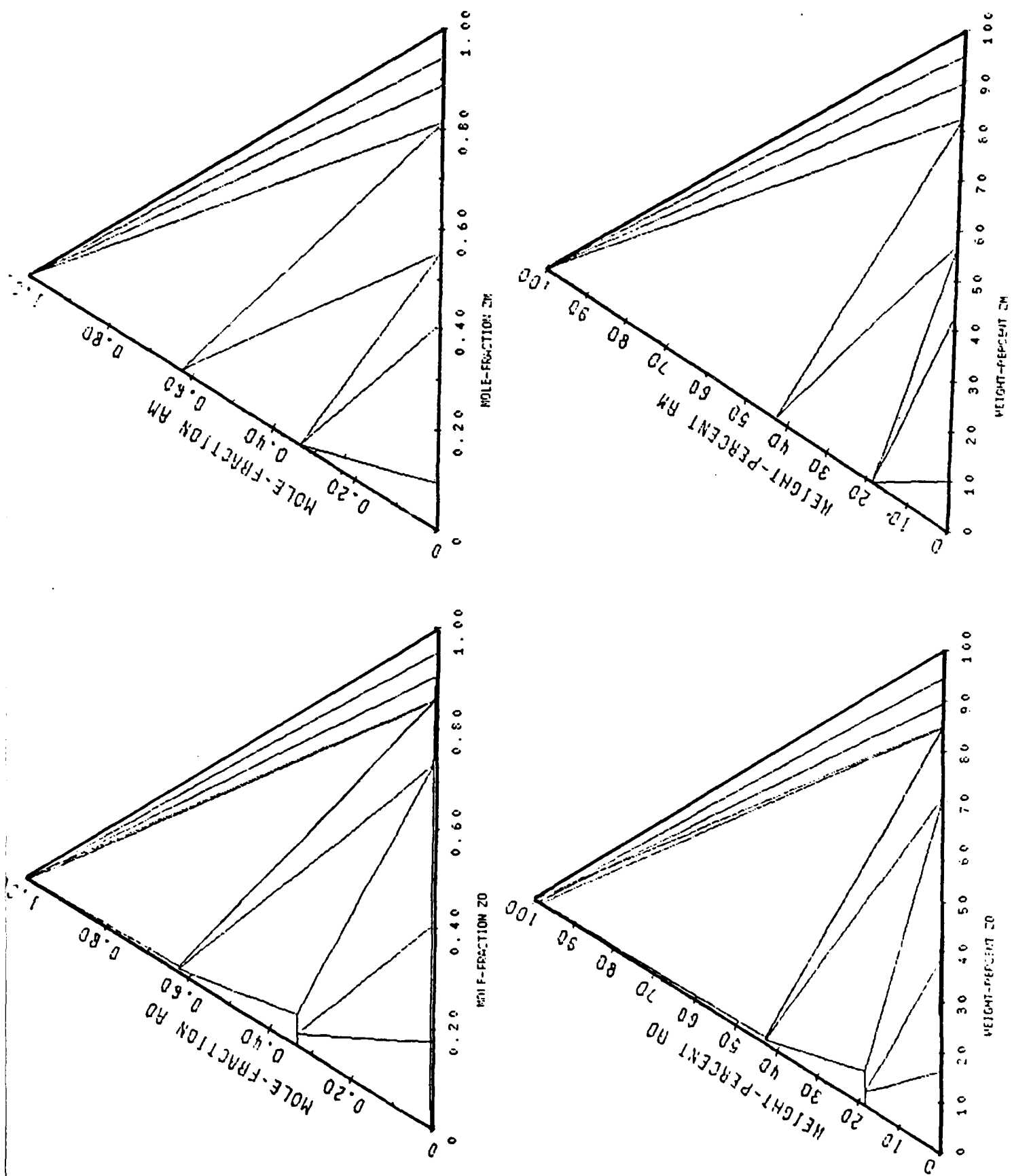


Figure 14 Thermocalc Output for AO-ZO-YO- and AM-ZM-YM at 1700K (See Figure 6)

Table 2 and Figure 2 show a similar exercise for the Fe-Cr-O system. The only difference being that the fcc metallic phase (γ) is the only one included in the THERMOCALC description while the FACT description (2) also contains the sigma (σ) and bcc (α) phases which emanate from the Fe-Cr edge. The Fe-Ni-O and Fe-Cr-O examples shown above demonstrate how this system can be employed to compute metal-metal-oxygen phase diagrams. Table 3 and Figures 3 and 4 show a similar exercise performed in a calculation of the Ti-C-N system at 1600K comparing the ManLabs system (Figure 4) and the THERMOCALC system (Figure 3). The fcc phase is the monocarbide Ti(C,N). The dashed lines in Figure 4 are the experimental phase boundaries in the Ti-1/2N₂ phase diagram. Table 3 and Figures 3 and 4 show how the THERMOCALC system can readily be applied to treating hard metal carbonitride systems.

The final example selected for the study is the Al₂O₃-ZrO₂-Y₂O₃ system which is shown in Table 4 and Figures 5-7. In this case the system is model as AO-ZO-YO ie 1/5 Al₂O₃-1/3 ZrO₂-1/5 Y₂O₃. Table 4 summarizes the lattice stability, solution and compound phase parameters along the lines described earlier in Sections III and IV. The systems in question have been recalculated on a mole metal basis ie as AM-ZM-YM where AM=1/2 /Al₂O₃, ZM=ZrO₂ and YM=1/2 Y₂O₃ on the basis of Table 5 by matching the Gibbs energies of each of the phases in the systems of interest. Table 5 summarizes the AM-ZM-YM description while Figures 8-14 show the results derived with THERMOCALC. Figure 8 shows YO-AO and YM-AM in mol and weight percent. The upper left panel in Figure 8 compares directly with the upper panel in Figure 5. Figures 9 and 10 show comparable results for ZO-YO, ZM-YM, ZO-AO and ZM-YM. Finally Figures 11-14 show isothermal sections at 2700, 2400, 2200 and 1700. Comparison of the weight percent ternary sections shows little difference between the gram atom model and the mole-metal model. These results illustrate the general utility of the THERMOCALC system!

REFERENCES

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